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U. S. NAVY DIVING-GAS MANUAL

October 1, 1969

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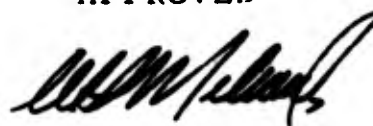
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U. S. NAVY DIVING GAS MANUAL

October 1, 1969

**U. S. NAVY SUPERVISOR OF DIVING
NAVAL SHIPS SYSTEMS COMMAND**

APPROVED



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Prepared Under Contract No. N⁰0014-66-C-0199

by
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INTRODUCTION

The concept of saturation diving has, for the first time, provided promise that practical undersea work can be carried out at depths approaching 1000 feet. The elimination of daily decompression requirements makes it possible for man to spend a significant part of each day at productive work at depth, and makes it economically feasible to undertake undersea tasks not previously practical. However, continued progress in development of saturation-diving techniques requires precise knowledge of physical and engineering properties of the helium-oxygen mixtures used for breathing gas, and no such information has yet been reported in published literature.

The principal objective of this manual is to provide the best available information on gas properties in a form convenient for use in diving research, engineering, and operations. All of the data in this manual are based upon calculation from theoretical relationships, substantiated where experimental information can be found in the literature (as for pure gases), and unsubstantiated where such information does not exist (as for helium-oxygen mixtures). It is felt that the data presented are the best that can be generated today, and that they will be generally satisfactory in the pressure range up to about 500 psi. For the higher pressures used in storing of diving gases, errors are indeterminate and may be significant. Future experimental research is needed to improve the state of knowledge of mixture properties at very high pressures, and to explore the properties of helium-oxygen-nitrogen mixtures.

A second objective of this manual is to summarize the present practice regarding choice of breathing-gas mixtures and some of the calculation procedures used in design and operation of diving equipment. An attempt has been made to present these procedures in such a manner that they can be used by both designers and operating personnel.

ACKNOWLEDGEMENTS

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Herbert R. Hazard directed the work and served as editor and principal author.

Joseph W. Walling developed the methodology and prepared the tables of gas properties found in Part 2.

John Broehl carried out the computer programming and calculations for tables of gas properties in Part 2.

Frederick A. Creswick prepared the humidity charts and examples. These were based upon an approach proposed by Dr. Richard A. Gaggioli, now with Marquette University.

Peter S. Riegel developed the chart of Figure 4 for semi-closed-system breathing gas composition and flow rates, and provided valuable technical review of the manual.

Arthur J. Coyle contributed the section on The Ocean Environment and provided valuable technical review of the work.

Mr. O. R. Hansen, U. S. Navy Office of Salvage, served as project monitor and provided valuable information, guidance, and technical review of the work.

TABLE OF CONTENTS

	<u>Page</u>
INTRODUCTION	
ACKNOWLEDGMENTS	
PART I. DIVING CONSIDERATIONS WITH HELIUM-OXYGEN MIXTURES	1
THE OCEAN ENVIRONMENT	1
Density of Seawater	1
Water Temperature	1
Atmospheric Conditions	2
Dynamic Aspects of the Ocean	2
PHYSIOLOGICAL REQUIREMENTS OF BREATHING-GAS MIXTURES	4
Breathing-Gas Composition	4
Oxygen Concentration	4
Diluent Gases	6
Decompression Considerations	6
Carbon Dioxide Concentration and Its Control	7
Breathing-Gas Consumption	8
OPERATING CHARACTERISTICS AND BREATHING-GAS CONSUMPTION OF DIVING APPARATUS	11
SCUBA Demand-Regulator Apparatus	12
Closed-Circuit Rebreathing Apparatus	12
Semiclosed-Circuit Breathing Apparatus	14
Calculation Procedure	16
Example 1	18
Example 2 - Using Available Gas Mixture	19
Example 3 - Deep Dive From Surface	19
Example 4 - Saturation Diving Mission	20
Surface-Supplied Open-Circuit Diving Rigs	21
Helium-Oxygen Deep-Sea Diving Outfit	24
Calculation of Breathing-Gas Oxygen Content and Flow Rate	24
MIXING OF BREATHING GASES	26
Mixing by Weight - Two or More Gases (Real-Gas Method)	27
Example 1. Oxygen-Helium Mixture	27
Example 2. Oxygen-Helium-Nitrogen Mixture (Real-Gas Method)	27
Mixing by Partial Pressures	29
Example 1. Mixing by Filling an Empty Receiver	31
Example 2. Refilling a Partially Filled Tank With the Same Mixture or a Different Mixture	33
Example 3. Changing Oxygen Content of Mixture by Adding Pure Oxygen	34

TABLE OF CONTENTS
(Continued)

	<u>Page</u>
Example 4. Mixing of Two Mixtures to Make a	
Third Mixture	36
Continuous-Flow Mixing Systems	39
USEFUL CAPACITY OF STORAGE TANKS (REAL-GAS METHOD)	41
VENTILATION OF HYPERBARIC PERSONNEL CHAMBERS	42
CARBON DIOXIDE ABSORPTION IN DIVING OPERATIONS	44
PSYCHROMETRIC CALCULATIONS FOR DIVING APPLICATIONS	46
Description of Psychrometric Charts	46
Definitions of Psychrometric Terms	50
Example Calculations	51
MISCELLANEOUS CALCULATIONS	57
Seawater Pressure at Depth	57
Oxygen Percentage at Specified Partial Pressure and Depth	58
Conversion of Liters to Cubic Feet	58
Conversion Factors for Viscosity	58
REFERENCES	59
PART 2. PROPERTIES OF HELIUM-OXYGEN MIXTURES	
AND OF PURE GASES	61
CONTENT AND FORMAT OF TABLES	61
SUMMARY OF DATA TABULATIONS T-2 THROUGH T-103	T-1
APPENDIX A. DENSITY OF SEAWATER	A-1
APPENDIX B. DATA SOURCES AND CALCULATIONS FOR	
PSYCHROMETRIC CHARTS	B-1
APPENDIX C. DATA SOURCES AND INTERPOLATION TECHNIQUES	
FOR CALCULATING TABLES OF PHYSICAL PROPERTIES OF GASES	C-1

U. S. NAVY DIVING-GAS MANUAL

PART I.

DIVING CONSIDERATIONS WITH HELIUM-OXYGEN MIXTURES

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PART 1. DIVING CONSIDERATIONS WITH HELIUM-OXYGEN MIXTURES

THE OCEAN ENVIRONMENT

The ocean environment can be described in terms of the density, temperature, and salinity of water; atmospheric conditions (barometric pressure, air temperature and humidity, winds and contaminants); tides and currents; and bottom topography, water clarity, and marine organisms. All of these factors interact to make an ever-changing and formidable environment for men and material. For the working diver these conditions become more rigorous with each additional fathom.

These conditions impose high hydrostatic pressure, dynamic forces from waves and currents, limited visibility, and low temperatures, which reduce the working diver's sense of touch and severely limit his effectiveness and time on the bottom. The forces which he can exert are reduced by his buoyancy which, when coupled with low visibility, can lead to disorientation, especially in deep dives and with the onset of narcosis.

The diver's ability to ventilate his lungs with increasing depth is greatly reduced because of increased breathing-gas density and breathing-apparatus resistance. For example, a diver breathing air at 200 feet has only 49 percent of his surface capability for ventilation, (1)* and this reduces his maximum work level to half that at the surface. This limitation occurs with the best diving apparatus available, and thus the importance of low breathing resistance in diving equipment is stressed.

This discussion will be limited to ocean parameters and processes that will be concerned with gas properties, diving equipment, and supporting systems, such as compressed gas systems, for the working diver.

Density of Seawater

The density of seawater varies with latitude, location, seasons, temperature, dissolved salt content (salinity), and depth (compressibility). Fortunately, the dominating factors, temperature and salinity, tend to counterbalance and, for most practical purposes, a density of 64 lb/ft³ or 0.4444 psi for each foot of depth, can be assumed. The error resulting from such an assumption should be less than 0.5 percent, or 5 feet in 1000 feet of water. The effects of compressibility on density is negligible (0.01 percent at 1500 feet). Appendix A discusses in detail the variations of salinity, temperature, and density in the oceans.

Water Temperature

Low-temperature water, in addition to its effects on divers, causes moisture to condense in breathing systems, where it can cause malfunction of regulators, increased

*References are listed on page 54.

flow resistance, and partial blockage of carbon dioxide scrubbers. Low temperature also reduces drastically the useful life of Baralyme carbon dioxide absorbent: in the operational Mark VIII and Mark IX apparatus, the useful life of Baralyme charges is reduced from 4 hours at 70 F to only 1/2 hour at 40 F. The increase in gas density with low temperature also increases the work of breathing, as volumetric flow rates must be maintained for control of carbon dioxide in the body.

Another serious consequence of cold water is the high rate of heat transfer from personnel transfer chambers, especially when helium-oxygen mixtures are used as the breathing gas. The design of such chambers, including the power and communication umbilical cables, should provide for the increased power required for heating and/or insulation of the chambers to maintain comfortable conditions.

Atmospheric Conditions

The interaction of air, water, and land to form winds, storms, and wave systems is well known. The resultant violent motions of surface support ships can impose severe dynamic loads on machinery and men. However, for the purpose of this manual, air temperature, humidity, barometric pressure, and contaminants are of greater interest. These conditions influence satisfactory design and operation of compressed gas systems. For instance, high inlet air temperatures and humidity can cause an appreciable reduction of compressor capacity, thus limiting the capability of supporting a team of divers in the water. Further, the capacity to pressurize and ventilate decompression chambers must be included in the system design. Contaminants in the environment, particularly CO₂ and CO, are critical to adequate design and safe operation of these systems. Since sufficient air must be provided to ventilate metabolic production of CO₂ by the diving team, any CO₂ contamination at the compressor inlet can greatly increase the required capacity of compressors and related systems.

Ambient air temperatures and heat-transfer processes external to decompression chambers contribute to the heating and cooling loads imposed on environmental control systems. Such systems have become increasingly important with deep and saturated diving, which require long periods for decompression. The fact that divers must spend long periods in confined spaces makes environmental control systems a critical necessity rather than just a desirable luxury.

The introduction of compressor lubricants into compressed-gas systems can present serious explosion and physiological hazards to the diver. Where water-lubricated and nonlubricated compressors are not available, the design, selection, and operation of lubricated systems must include consideration of the reduction of the carry-over of lubricants from compressors or alternative supply systems; cylinder temperatures; controlled lubrication; adequate inter-and-after coolers; and filtering systems.

Dynamic Aspects of the Ocean

The interaction of tides, winds, barometric pressure, the rotation of the earth, variable density, bottom topography, river outflow, and other factors lead to water

movement, stratification, and mixing. For example, offshore winds reduce tide levels and push warm surface waters seaward to be replaced by cold and dense waters upwelling from the bottom. These waters can be rich in nutrients and clouded with silt. The nutrients result in growth of marine organisms which, along with suspended matter in the water, reduce underwater visibility. Onshore winds can reverse this process. Further, the transport of warm surface water into a region with cold air can result in dense fogs and other hazards.

Of the many complex processes in the ocean, those having the greatest impact on diving conditions include changes in heat transfer caused by water flow around the divers, the hoses, and the personnel transfer chambers (PTC); loss of gases through solution at open PTC hatches; buildup of humidity; contamination of gases in such chambers by products of marine decomposition; and finally, fluctuation of hydrostatic pressure caused by waves, tide changes, and PTC motion.

PHYSIOLOGICAL REQUIREMENTS OF BREATHING-GAS MIXTURES

The composition of breathing-gas mixtures and the quantity required are determined by physiological characteristics of the human body when exposed to elevated pressures. The breathing mechanisms that serve to regulate partial pressures of oxygen and carbon dioxide in the blood at sea level serve equally well under the sea if the partial pressures of oxygen, nitrogen, and carbon dioxide, as well as the respiratory volume, are similar to those at sea level. Medical research has provided sufficient information on physical tolerance to partial-pressure levels to permit safe selection of both composition and volumes of breathing gas for diving, and this information forms the basis for the discussion that follows.

Breathing-Gas Composition

Oxygen Concentration

The breathing-gas mixtures used in diving must be such as to be nontoxic over the range of working depth, and must include sufficient oxygen for normal activities at depth. Air can be used only for relatively shallow dives because both nitrogen and oxygen become toxic at elevated partial pressures. For deeper dives over longer periods, helium is preferred as the inert gas in the mixture because it is nontoxic, and because its low density reduces breathing effort. The percentage of oxygen in the breathing mixture must be reduced with increasing depth to maintain the oxygen partial pressure within a range of about 0.21 to 1.2 atm (atmospheres absolute pressure)(3 psia to 17.7 psia). For saturated diving over long periods, the preferred mixture would contain oxygen at a partial pressure between 0.21 and 0.30 atm, near that in normal air at sea level. The limits of oxygen concentration in the breathing gas are determined by the levels of solution of oxygen in the blood, which are related directly to oxygen partial pressure. In order to maintain blood oxygen levels within acceptable limits, oxygen partial pressure is held about constant regardless of total pressure, and the remaining pressure is provided by increasing the proportion and partial pressure of diluent gas.

Figure 1 relates the percentage of oxygen in the breathing mixture to depth, and to oxygen partial pressure in both psia and atm. The curves show a large area within which mixtures are physiologically acceptable, bounded on the lower side by curves defining different levels of anoxia (oxygen deficiency) and on the upper side by levels of oxygen toxicity. It will be noted in Figure 1 that the first symptoms of anoxia occur when oxygen partial pressure falls to 0.16 atm, and that the diver becomes helpless at a partial pressure of 0.12 atm.⁽²⁾ In the region of high oxygen concentrations, it may be seen that oxygen toxicity limits the maximum partial pressure for long-term exposure to about 1.2 atm, but that higher concentrations can be tolerated for short periods.⁽²⁾ These toxicity limits are imposed by effects on the central nervous system that result in unconsciousness and convulsions. A different limit is encountered in saturation diving over extended periods, which is imposed by lung irritation and eventual lung damage. This occurs with 10 percent incidence in a 14 day period with an oxygen partial pressure of 0.6 atm.⁽³⁾ In order to avoid pulmonary irritation, present practice in saturation diving for extended periods is to limit oxygen partial pressure to 0.3 atm (5 psi).⁽³⁾

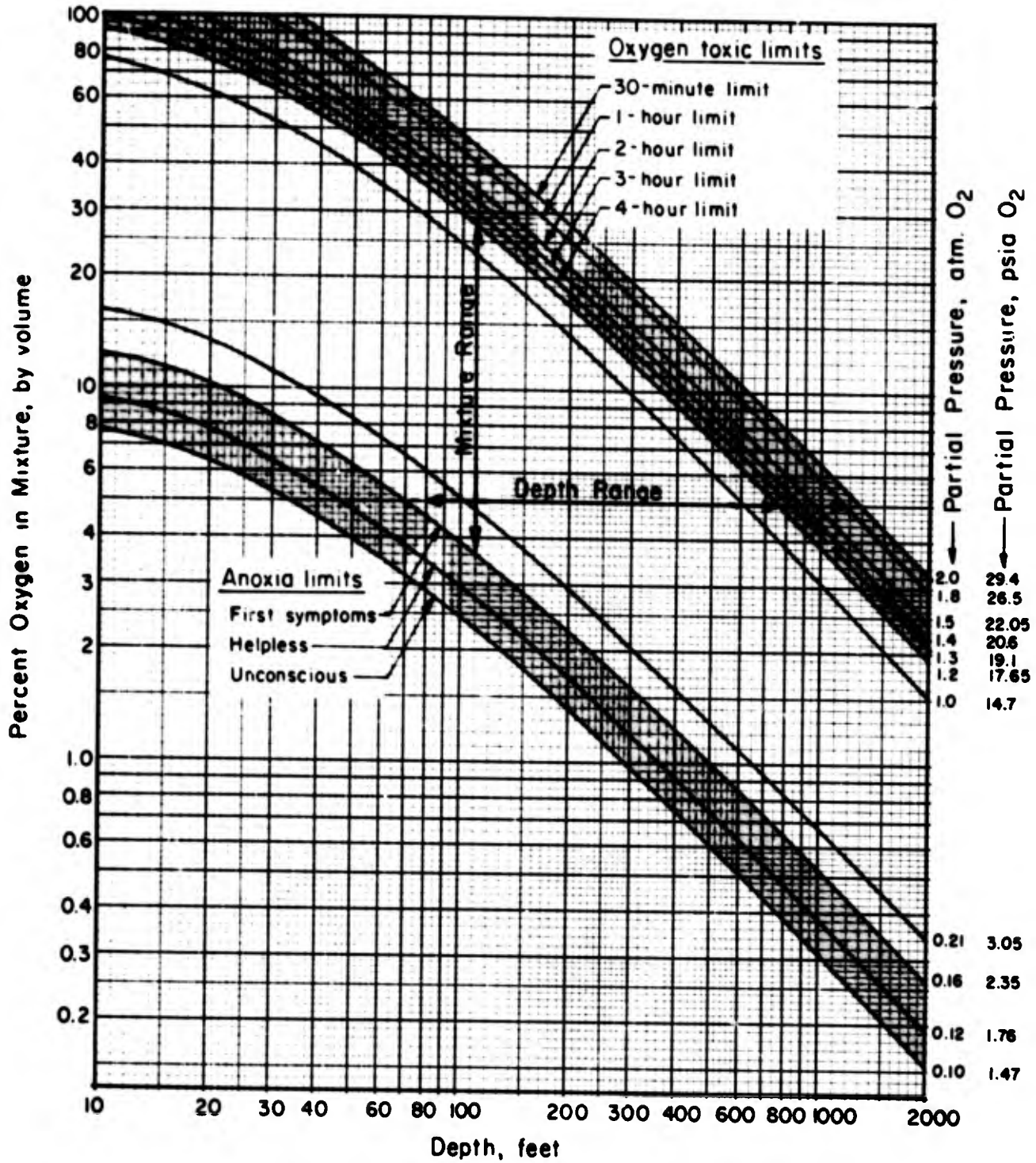


FIGURE 1. PERCENTAGE OF OXYGEN IN BREATHING MIXTURE AS A FUNCTION OF DEPTH AND OXYGEN PARTIAL PRESSURE

$$\text{Percent O}_2 = \frac{\text{Partial pressure O}_2, \text{ atm} \times 100}{\text{Seawater pressure, atm}}$$

From inspection of Figure 1 it is evident that, at any fixed depth, it is feasible to breathe a wide range of mixtures without ill effects. For example, at a depth of 200 feet, the mixture could be as lean as 3 percent oxygen (0.21 atm) or as rich as 17 percent oxygen without encountering any short-term physiological limits. Likewise, with a fixed oxygen concentration in the breathing mixture, the diver can make long excursions in depth. A mixture containing 10 percent oxygen, for example, would permit operations between 36 and 360 feet, with the possibility of short excursions as deep as 600 feet. However, it is important to note that dangerous anoxia can occur if lean mixtures intended for deep operations are breathed during ascent and decompression. It is necessary to provide higher oxygen concentrations at shallower depths to stay within the working range indicated in Figure 1.

Diluent Gases

The oxygen percentages shown in Figure 1 would be applicable with any diluent gas or mixture of diluent gases that might be applicable to diving. Oxygen-nitrogen mixtures, including air, can be used freely if nitrogen partial pressure is limited to 3 atm, the pressure at which the first symptoms of nitrogen narcosis appear, although nitrogen partial pressures to 5.5 atm can be used with care by experienced divers for short periods. (2) (This is equivalent to diving at 200 feet with air.) For greater depths, helium is the diluent gas now generally used because it is nontoxic at any practical diving depth and because its low density reduces breathing effort. The principal disadvantages of helium are its limited availability and high cost, its high sonic velocity which results in high-pitched speech that is hard to understand, and the high leakage rates experienced when minute imperfections are present in joints and seals. Nitrogen at partial pressures up to 3 atm can be mixed with helium as a means of conserving helium, improving speech intelligibility, and reducing cost, but precise information on physical and engineering properties of helium-nitrogen-oxygen mixtures is not available at the present time.

Hydrogen has been proposed as a diluent gas for depths beyond 1000 feet because of its low cost and because its density is half that of helium. However, the effects of hydrogen on body tissues at high pressures have not yet been explored. Furthermore, hydrogen-oxygen mixtures are readily ignited by static-electricity discharges unless oxygen concentrations are below the combustible limit, which is at 6 percent oxygen at 1 atm pressure. (4) Explosive limits of oxygen-hydrogen mixtures have not yet been investigated at high pressures and are not known, but would be well below 6 percent oxygen.

Decompression Considerations

Decompression considerations also influence the selection of breathing-gas composition. The length of the decompression period is a function of the partial pressure of diluent gas (helium or nitrogen) in the breathing mixture. The decompression period can be shortened by using oxygen concentrations close to the toxic limit during a scheduled diving period and during the decompression period. (2, 3) This approach is most effective for dives at shallow to moderate depths, where the proportion of oxygen can be large enough to significantly reduce the partial pressure of diluent gas. Under marginal conditions, use of a high-oxygen mixture may eliminate the decompression period entirely. Oxygen is not normally a limiting factor in decompression because it appears to be metabolized at a rate sufficient to avoid adverse effects.

Carbon Dioxide Concentration and Its Control

In normal breathing, the concentration of oxygen in the breathing gas is reduced and the oxygen is replaced by a nearly equal volume of carbon dioxide. If carbon dioxide is included in the inhaled gas, the partial pressure of carbon dioxide in the blood increases, and the respiratory center in the brain increases breathing rate to restore normal carbon dioxide tension. Excessive amounts of carbon dioxide in the breathing gas result in toxic effects, as summarized in Table 1. ⁽⁵⁾

TABLE 1. RELATION OF PARTIAL PRESSURE OF CARBON DIOXIDE TO TOXIC EFFECTS

CO ₂ , vol % (Sea Level)	Partial Pressure		Physiological Effects
	Atm	Psi	
0-1	0.00-0.01	0.00-0.15	None
1-2	0.01-0.02	0.15-0.30	Increase in breathing rate, slight hearing loss
2-6.5	0.02-0.065	0.30-0.96	Mental depression, headache, dizziness, nausea, visual-acuity loss
6.5+	0.065+	0.96+	Dizziness, stupor, unconsciousness

As with oxygen, the physiological effects of carbon dioxide depend upon its partial pressure in the blood and, therefore, in the breathing gas. The percentage of carbon dioxide in the breathing gas that can be tolerated at diving depths decreases with depth, and the respiratory volume required to ventilate the lungs at depth remains approximately equal to that at sea level.

If a diver's breathing circuit includes dead space from which he rebreathes exhaled air, this dead space must be ventilated to dilute the carbon dioxide partial pressure to a nontoxic level. If all of the diver's carbon dioxide production flows into the dead space and is mixed with all of the ventilating air, the reduction of partial pressure depends only upon the ratio of carbon dioxide flow to dilution-air flow, and this ratio must be about 1:100 to limit carbon dioxide to 0.01 atm partial pressure. This is the usual condition in helmet ventilation.

The quantity of ventilation gas can be reduced if the concentration of carbon dioxide can be reduced by other means, such as absorption by a carbon dioxide absorbent. Thus, the standard deep-diving helmet has an attached canister of absorbent through which the breathing gas in the helmet is recirculated to reduce carbon dioxide level and, thus, reduce the quantity of breathing gas required.

In closed-circuit and semiclosed-circuit breathing apparatus, all exhaled gas can be rebreathed. This is made possible by passing all exhaled gas through a canister of carbon dioxide absorbent before it can again be inhaled. Excess exhaled gas is vented between the mouthpiece and the absorbent canister in semiclosed-circuit apparatus, and incoming gas can be introduced in an ejector to assist in circulation through the canister.

Unventilated dead space has a significant effect on the respiratory volume required for control of carbon dioxide tension. Examples of unventilated dead space are the natural volume of the mouth and throat, the volume of a mouthpiece between non-return valves, or the volume of a full face mask. The effect of dead space is to increase the tidal volume, or volume for each inhalation, by the amount of volume of the dead space. Thus, if normal tidal volume is 1 liter and dead space in a full face mask of 1/2 liter is added, the tidal volume needed for equivalent carbon dioxide tension in the blood increases to 1.5 liters. The extra tidal volume needed would reduce maximum work level, increase the quantity of breathing gas needed by 50 percent, and increase the effort of breathing. If tidal volume of 1.0 liter is maintained, only 0.5 liter of fresh gas would be obtained with each inhalation, and this would double the respiratory frequency and gas consumption. It is always desirable to minimize dead space in any diving equipment to the extent possible while meeting other requirements.

When divers are supplied with air from compressors at the surface there is always danger that some carbon dioxide may find its way into the compressor inlet. Exhaust gas from shipboard engines, for example, or from motor traffic in harbor areas can result in significant air contamination. If the compressed air contains 0.2 percent carbon dioxide, for example, it contains half the toxic limit of 0.4 percent at 5 atm pressure (130 feet depth). Thus, only half of the air supplied is useful as diluent for the carbon dioxide generated by the diver's respiration, and the quantity of air supplied must be doubled to avoid toxic reactions. Thus, it is extremely important to assure that clean, fresh air is piped to the compressor inlet, and that it is not downwind from any nearby sources of carbon dioxide.

Breathing-Gas Consumption

Breathing gas contains oxygen that is consumed in respiration and a diluent gas such as nitrogen or helium that is not consumed or altered. The oxygen consumed in the body reacts with carbon and hydrogen to produce CO₂ and water, and the CO₂ thus formed is exhaled in the breathing gas. In addition, the exhaled gas is saturated with water vapor. The volume fraction of CO₂ released varies from 0.7 to 0.9 relative to the volume of oxygen consumed, depending upon the carbon-hydrogen ratio of food being metabolized. (2)

Figure 2 relates oxygen consumption and respiratory volume to rate of exertion. The data for Figure 2 are based on experimental measurements, (2, 3) and provide a basis for selection of breathing-gas quantities for diving equipment. (6)

The mass rate of oxygen consumption and corresponding carbon dioxide production varies with the rate of exertion, and is independent of depth. As indicated in Figure 2, values of oxygen consumption vary from about 0.5 standard liters per minute (slm) when at rest to about 4.0 slm with heavy exertion. These values are equivalent to 0.0895 and 0.716 pound of oxygen per hour, or 1.14 to 9.13 standard cubic feet (scf) per hour*. The actual volume of oxygen consumed at depth would decrease in inverse proportion to absolute pressure in accordance with gas laws.

The values for respiratory minute volume (RMV) shown in Figure 2 are determined by carbon dioxide ventilation requirements within the body, and these volumes are

*The standard liter is defined at 0°C (32°F) and the standard cubic foot is defined at 60°F. Thus, although 1 ft³ = 28.3 liters, one standard ft³ = 26.3 standard liters.

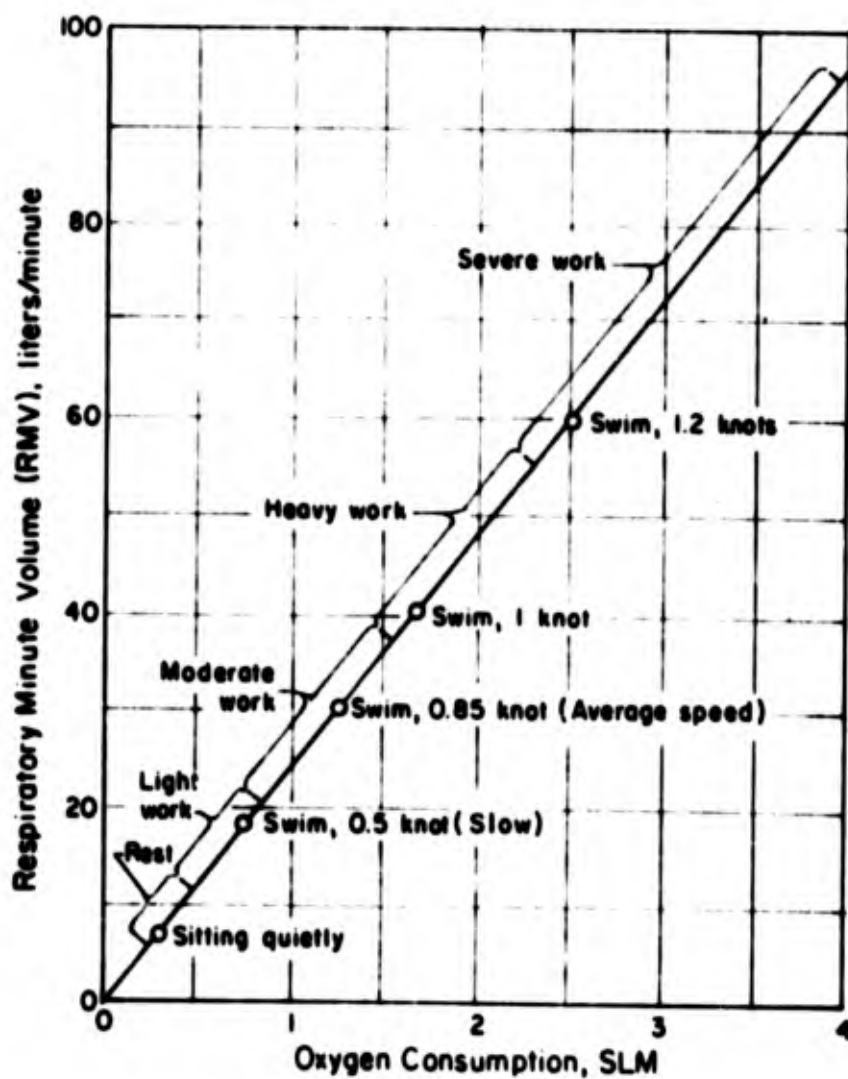


FIGURE 2. RELATION OF RESPIRATORY VOLUME AND OXYGEN CONSUMPTION TO TYPE AND LEVEL OF EXERTION⁽³⁾

the same at diving depths. Thus, the mass flow of gas to meet respiratory requirements increases in direct proportion with depth and with oxygen consumption. The ratio of respiratory volume to the volume of oxygen consumed is such that only 20 percent of the oxygen inhaled is consumed, and the other 80 percent is exhaled. In sea-level respiration the oxygen content of inhaled air is reduced from 20.8 percent to 16.6 percent, equivalent to a reduction from 0.208 to 0.166 atm partial pressure.

In any demand type breathing apparatus, the rate of supply of breathing gas must satisfy the RMV requirement. Thus, with oxygen partial pressure of 0.208 atm in the breathing gas, only 20 percent of the oxygen supplied will be consumed. The actual gas consumption will be 10 to 100 liters per minute (lpm) at depth, depending upon level of exertion.

In closed-circuit rebreathing apparatus supplied with pure oxygen, the rate of oxygen consumption would be equal to the rate of supply, as all carbon dioxide is absorbed in recirculation and no gas is vented. The oxygen consumption rate would depend only upon rate of exertion and would be independent of depth. Oxygen consumption can be 0.5 to 3 slm, and about 2 slm is average. (2)

In semiclosed-circuit breathing apparatus, the problems of oxygen supply and usage are more complex. Oxygen is supplied to the inhalation bag at 1.2 atm partial pressure, and with a preset mixture flow rate sufficient to provide oxygen for continuous exertion at 3 slm oxygen demand. The oxygen concentration in the inhalation bag will vary from about 1.0 atm at rest, to 0.21 atm during severe exertion. Excess gas is vented between the exhalation bag and the carbon dioxide absorber, at oxygen partial pressures varying from 0.8 atm at rest to 0.16 atm during heavy exertion. The total flow rate of breathing gas depends upon the depth, and can vary greatly because of the wide range of depths and mixtures for which the semiclosed-system is suitable.

The volume of breathing gas needed for surface-supplied diving apparatus having ventilated dead space, such as helmet or face mask, is determined by the flow rate needed to dilute carbon dioxide in the dead space to nontoxic levels of 0.01 to 0.02 atm partial pressure. The flow rate needed is similar to the RMV at maximum exertion. It is usual to supply 4.5 cfm (127 lpm) of gas, measured at depth. This would be sufficient to permit exertion levels corresponding to oxygen consumption of 1.27 slm if carbon dioxide in the dead space is held to a level of 0.01 atm, or 2.55 slm for 0.02 atm of carbon dioxide. More air would be required should levels of exertion exceed those corresponding to 2.5 slm of oxygen consumption.

The helium-oxygen deep-sea diving outfit operates much like the semiclosed-circuit apparatus in that breathing gas in the helmet is recirculated through a carbon dioxide absorber to minimize the flow of breathing gas needed. However, flow rate is determined by ventilation requirements. Breathing gas is supplied from the surface through a hose at a rate of 0.5 cfm, measured at depth, and enters the helmet through an ejector which drives the recirculation flow. With a recirculation ratio of about 10, the ventilation effect is equivalent to 5 cfm or 140 lpm, which is adequate for heavy work. The oxygen content of the mixture is maintained at a high level and can be varied during the dive to suit depth and decompression requirements. If this rig is used at shallow depths with air, anoxia could result unless flow rate is increased to provide needed oxygen. The required flow rate will reach 2.1 scfm at the surface for oxygen consumption of 3 slm.

OPERATING CHARACTERISTICS AND BREATHING-GAS
CONSUMPTION OF DIVING APPARATUS

All breathing apparatus used in diving must, in some way, satisfy both respiratory-volume requirements and oxygen-supply requirements at all depths. Various types of apparatus having different characteristics have been developed, each having applications of greatest suitability. The general characteristics of each type of system, and means of calculating or estimating the composition and flow rates of breathing gas required for a diving mission are discussed below.

Table 2 compares the relative air-flow rates needed by four basic types of underwater breathing equipment under identical conditions: the demand-regulator SCUBA, a semiclosed-circuit system, a surface-supplied deep-sea diving outfit, and a surface-supplied deep-sea diving outfit with carbon dioxide absorption provisions. For these comparisons, the breathing gas is air containing 21 percent oxygen. Data are shown for three levels of oxygen consumption: 3.0 slm, corresponding to heavy exertion, 2.0 slm, corresponding to moderate exertion, and 1 slm, corresponding to light exertion or rest. From the table it may be seen that for heavy exertion, the air flow required for the semiclosed-circuit system decreases markedly with increasing depth, and that air flow required for the other systems increases markedly with depth. Thus, the semiclosed-circuit system appears especially attractive for deep-diving missions, the SCUBA for shallow dives, and the deep-sea rigs for all heavy-duty applications where surface supply of air is advantageous. It should be noted that, at shallow depths, 0.5 cfm of air does not supply enough oxygen for exertion in the deep-sea diving outfit with carbon dioxide absorption. Flow rate then becomes equal to that for the semi-closed circuit apparatus.

TABLE 2. COMPARISON OF AIR-FLOW RATES FOR DIFFERENT BREATHING SYSTEMS

Depth, ft	Pressure, atm	O ₂ Demand, slm	Air Supplied to Breathing Apparatus, slm			
			Semiclosed	SCUBA	Deep Sea Rig (a)	Deep Sea (CO ₂ abs)(b)
150	5.54	3.0	16.3	388	707	78.6
50	2.51	3.0	19.7	176	320	35.6
30	1.91	3.0	22.7	134	244	27.2
20	1.60	3.0	25.6	112	204	25.6(c)
10	1.30	3.0	32.9	91	166	32.9(c)
0	1.00	3.0	60.0	70	127	60.0(c)
30	1.91	2.0	15.2	89	244	27.2
10	1.3	2.0	22.0	61	166	22.0(c)
0	1.0	2.0	40.0	47	127	40.0(c)
10	1.3	1.0	11.0	0	166	18.5
0	1.0	1.0	20.0	23	127	20.0(c)

(a) 4.5 cfm air as measured at working depth.

(b) 0.5 cfm air as measured at working depth.

(c) Flow rate needed to provide needed oxygen.

For all of the conditions listed in Table 2, the air required to supply the semiclosed-circuit system is less than for the demand regulator system. However, this is not necessarily true of all missions. For example, the demand regulator system operates on actual demand for breathing gas, which may be low if little exertion is required, while the semiclosed circuit system requires continuous flow sufficient for the highest exertion level.

SCUBA Demand-Regulator Apparatus

Open-circuit SCUBA equipment is based upon use of a demand regulator, with which gas is supplied with each inhalation in the exact quantity needed, thus conserving stored compressed gas. However, all exhaled gas is vented and lost, so that the volume rate of usage is equal to the respiratory volume. The actual rate of usage of stored gas, and the duration of the stored supply will, thus, depend upon the rate of effort and the depth of usage. Volume-flow rates required for breathing will vary between 7 and 70 slm, or 15 and 150 ft³/hr at the density corresponding to depth⁽²⁾, depending upon level of effort.

Figure 3 shows the variation of endurance with depth and level of effort for a demand-regulator SCUBA system with tank capacity of 1450 in.³ (twin 90 ft³ tanks) charged to 3000 psi. It may be seen that mission duration can range from 4 hours in shallow water with little exertion, to only 15 minutes at 130 feet with heavy exertion. These curves provide for reserve air at 500 psi in one tank only, and do not include air for decompression if needed.⁽⁵⁾ The endurance of other SCUBA systems would depend upon their tank capacity and pressure, and similar curves could be drawn.

Closed-Circuit Rebreathing Apparatus

In principle, it is possible to breath pure oxygen in a closed system if exhaled oxygen and carbon dioxide are passed through an absorbent that removes the carbon dioxide, and fresh oxygen is added at the rate of consumption. Such apparatus has been developed and used within the depth limitations imposed by oxygen toxicity. Although suitable only for shallow depths, such apparatus leaves no trail of bubbles, making it advantageous for clandestine military operations. It is also extremely efficient in use of oxygen, as all of the oxygen stored is actually used. Finally, because of the efficient oxygen usage and the elimination of diluent gas, sufficient oxygen can be carried for rather long diving periods, and the diving time is independent of depth. Diving periods are limited by both the quantity of oxygen carried and the amount of CO₂ absorbent provided, and these should be approximately balanced in design of equipment.

Several experimental closed-circuit rigs now under development use diluent gas in breathing bags to permit extension of closed-circuit advantages to greater depths. The only diluent gas needed is that required to fill the breathing bags and to adjust their volume with changes in depth; pure oxygen is added to the breathing mixture at the rate of consumption, so that all oxygen is utilized, and the duration of oxygen supply is independent of depth. These systems rely upon an automatic control system to sense and maintain oxygen partial pressure within tolerances of 0.2 to 1.2 atm. If it proves possible to meet reliability requirements, such systems should be extremely advantageous for deep operations as they consume virtually no diluent gas and only minimum quantities of oxygen.

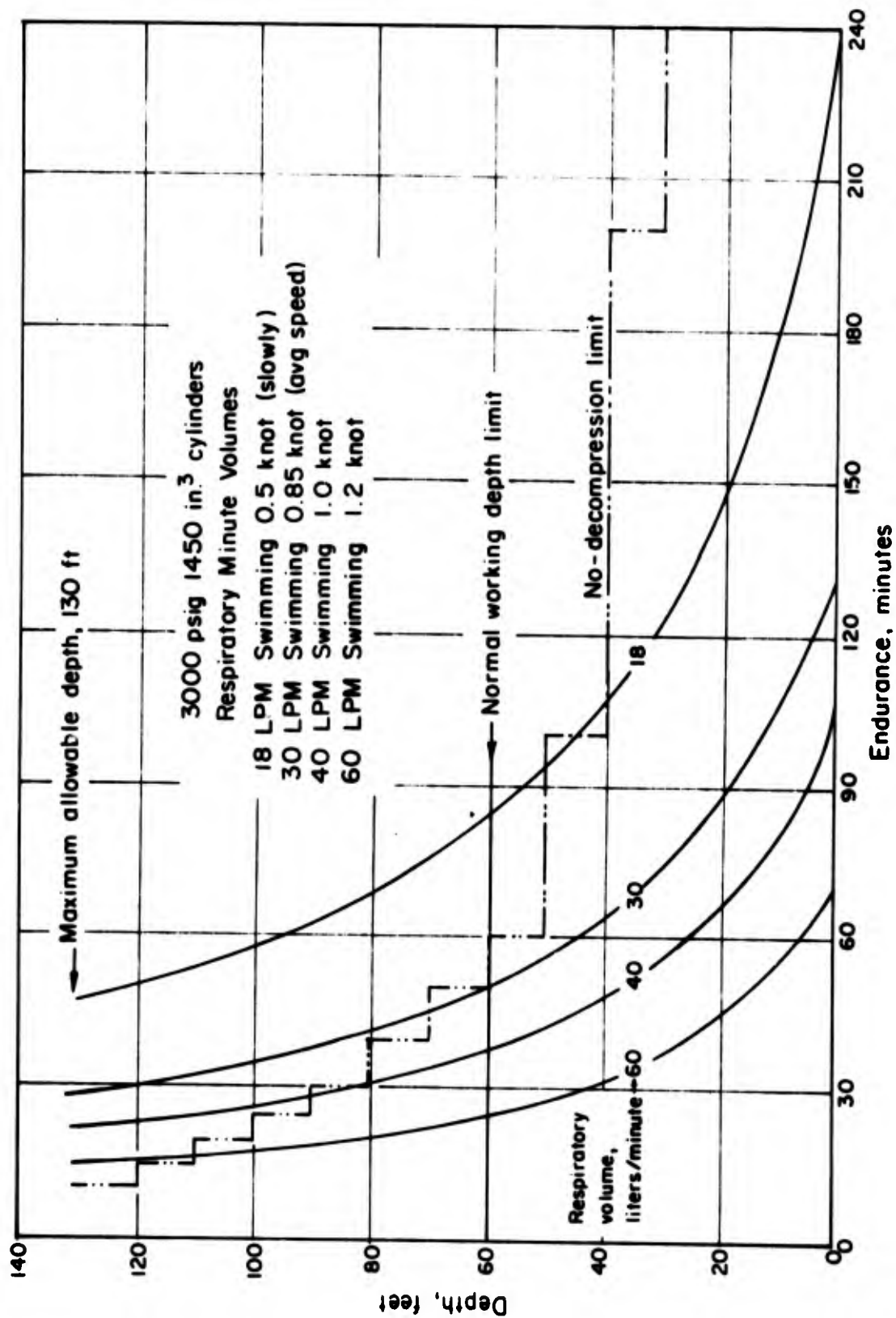


FIGURE 3. RELATION OF ENDURANCE OF SCUBA TO DEPTH AND LEVEL OF EXERTION

Twin 90 ft³ tanks at 3000 psi. Reserve of 500 psi in one tank. (5)

In any pure-oxygen rebreathing apparatus, the rate of consumption is the oxygen usage rate. For normal underwater activities, this often averages 2 slm, and increases to 3 slm with heavy exertion. Thus, oxygen storage of about 120 to 180 sl, or 4.57 to 6.85 scf, per hour of endurance is required.

Semiclosed-Circuit Breathing Apparatus

Semiclosed-circuit breathing apparatus is of special interest because it is adaptable to use at very great depths and is very efficient in use of breathing gas. In theory, oxygen can be supplied to the apparatus at a partial pressure of 1.2 atm, just under the toxic limit, and rebreathed through a CO₂ absorber until oxygen partial pressure is reduced to 0.16 atm on exhalation. In this way, as much as 86 percent of the oxygen supplied could be utilized, compared with only 20 percent for an open-circuit system. In addition, the high partial pressure of oxygen in the gas mixture minimizes the amount of mixed gas used for the required oxygen consumption.

In a typical helium-oxygen semiclosed-circuit apparatus suitable for saturated diving, breathing gas is supplied continuously at a mass flow rate (referred to as the "liter-flow" rate) sufficient to provide for oxygen usage of about 3 slm. The partial pressure of oxygen in the mixture supplied to the apparatus is at 1.2 atm, just under the toxic limit when breathing incoming gas. The incoming gas enters an inhalation bag where it mixes with and is diluted by partially depleted gas. After inhalation it is exhaled into an exhalation bag and then circulated through a CO₂ absorber back to the inhalation bag. Surplus exhaled gas is vented from the exhalation bag at a rate that is somewhat less in volume than the incoming stream because of CO₂ absorbed. The level of oxygen partial pressure in the inhalation bag will vary with the rate of oxygen usage, from about 1.0 atm when at rest to 0.21 atm at maximum exertion, and the partial pressure of oxygen in vented gas will vary from 0.96 to 0.16 atm. With constant mass flow of breathing gas, most efficient use of the supply occurs only at maximum exertion; at other conditions, surplus oxygen is vented.

The quantity of oxygen that must be supplied is independent of depth if the range of oxygen partial pressures from 1.2 to 0.21 atm is maintained in the inhalation bag at all depths. This is done by selection of oxygen concentration and liter flow rate. However, to avoid toxicity, the oxygen percentage must decrease as depth increases, so that the quantity of helium supplied with the oxygen will increase with depth. Thus, the total consumption of breathing gas will increase in proportion to total pressure and depth.

By its nature the semiclosed-circuit apparatus is most economical of breathing gas if used within a moderate depth range. In setting up the equipment for a mission, the oxygen percentage is limited by the maximum depth planned, and the liter-flow rate is established to provide sufficient oxygen for exertion at the minimum depth. As the maximum and minimum depths become farther apart, the difference in oxygen partial pressure between breathing gas and vented gas is reduced, which reduces the efficiency of oxygen utilization and increases the required liter-flow rates.

Figure 4 is a chart for selecting oxygen percentage and liter flow rate for semiclosed-circuit breathing apparatus. It shows the relation of oxygen content of the breathing gas, partial pressure of oxygen in the inhalation bag, and liter flow, to the

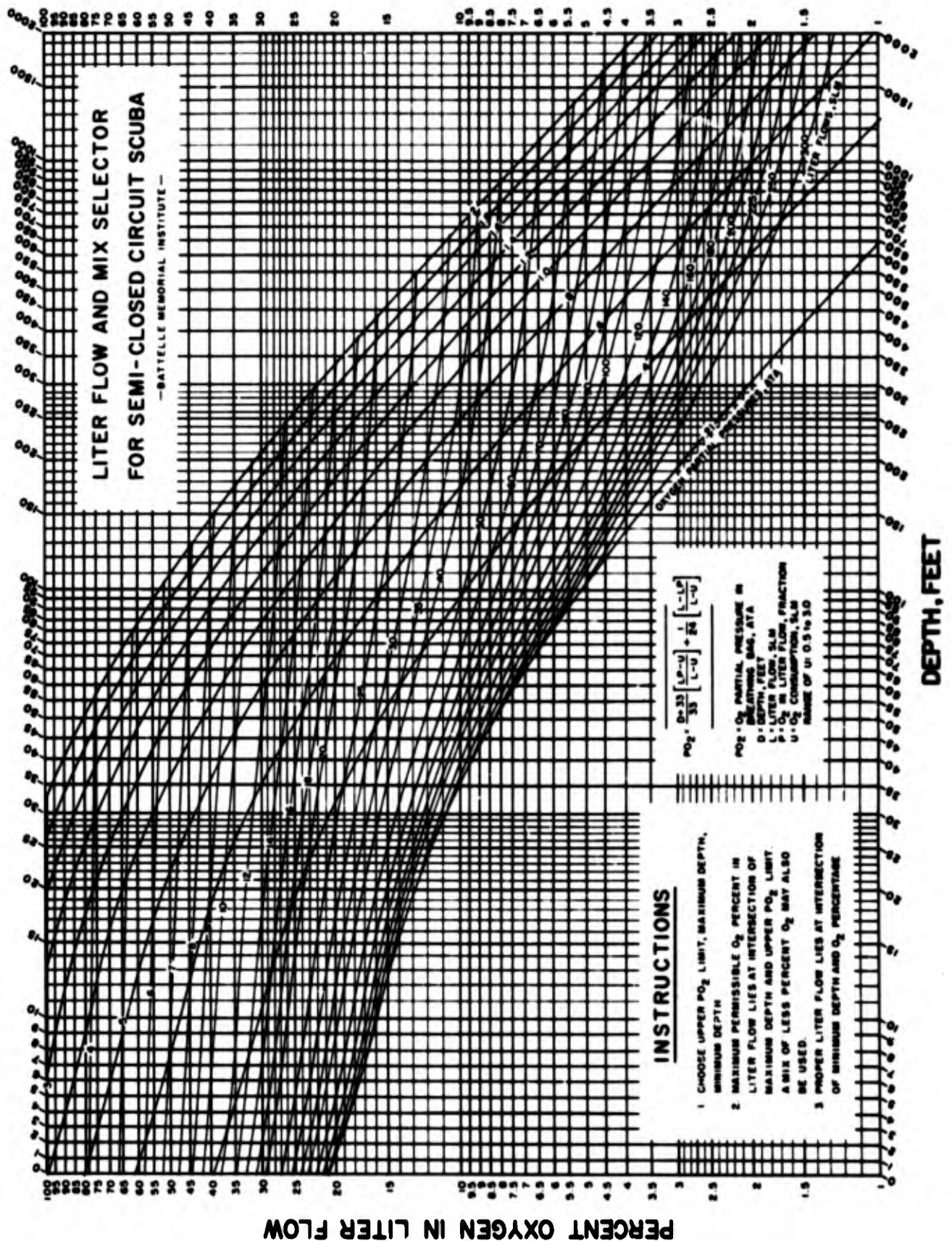


FIGURE 4. CHART FOR SELECTING LITER FLOW AND OXYGEN PERCENTAGE FOR SEMICLOSED-CIRCUIT BREATHING APPARATUS

Partial pressure lines refer to oxygen partial pressure in inhalation bag with 0.5 slm oxygen usage.

diving depth. In using the chart the oxygen percentage in the liter flow is found at the intersection of the vertical line of maximum depth and the sloping curve of 1.2 atm oxygen partial pressure; the liter flow rate is found to the left at the same oxygen percentage, on the vertical line representing the minimum diving depth. When used in this way the conditions selected will provide 1.2 atm oxygen partial pressure in the inhalation bag when resting and consuming 0.5 slm oxygen at the greatest depth, and sufficient flow for severe exertion with consumption of 3 slm oxygen at the least depth. It should be noted that the oxygen partial pressure lines indicate higher oxygen percentages than in Figure 1 because they represent remaining oxygen partial pressure after some oxygen has been used. For short dives it is possible to utilize higher oxygen partial pressures if the time limits indicated in Figure 1 are observed.

The semiclosed-circuit breathing apparatus is currently the most versatile breathing apparatus available. It is suitable for use at any depth to 1000 feet, as demonstrated in experimental test-chamber dives, Sealab programs, and other saturation-diving missions. Accordingly, the rates of consumption of oxygen and of helium can vary widely with depth, and with range of depth required for specified missions. The procedures for determining oxygen concentration and gas flow rates are shown by example below.

Calculation Procedure

The procedure for calculating oxygen concentration in breathing gas and the liter-flow rate for semiclosed-circuit apparatus is as follows:

- (1) Select the maximum diving depth, the minimum depth, and the maximum oxygen demand rate.
- (2) Determine percentage of oxygen in the breathing gas that corresponds to 1.2 atm at the maximum diving depth.
- (3) Determine percentage of oxygen in exhalation bags and vented that corresponds to 0.166 atm at the minimum diving depth or at the surface.
- (4) Calculate liter-flow rate for breathing gas using above oxygen percentages.
- (5) If desired, check efficiency of oxygen utilization and oxygen supply rate.
- (6) If pertinent, check oxygen available for surfacing if this is not provided for in Step (3).

The diving depths will be defined by mission objectives. The maximum oxygen demand could be estimated on the basis of rate of effort planned, using Figure 2 as a guide. However, it is recommended that, as a standard procedure, an oxygen demand of 3.0 slm be provided in order to permit extreme exertion should an emergency arise.

The percentage of oxygen in the breathing gas corresponding to 1.2 atm oxygen partial pressure at the maximum diving depth can be read directly from Figure 1 or Figure 4. The value of 1.2 atm is low enough for saturation diving, but higher partial pressures can be used for shorter periods within time limits shown in Figure 1 should this be necessary.

The partial pressure of oxygen in exhalation bags should be taken as 0.166 atm at the minimum diving depth; for dives from the surface, it would be usual to provide enough oxygen for surface swimming.

The liter flow can be calculated using the following equation:

$$L = \frac{U}{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right)}$$

where

L = liter flow, slm

U = oxygen used (normally 3.0 slm)

O₁ = percentage (decimal) of oxygen in liter flow (1.2 atm at greatest depth)

O₂ = percentage (decimal) of oxygen vented from exhalation bag. (0.166 atm at least depth).

O₃ = percentage (decimal) of oxygen inhaled from inhalation bag. (0.208 atm at least depth).

The efficiency of oxygen utilization, E, can be determined as:

$$E = \frac{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right)}{O_1}$$

The oxygen flow in the breathing gas, O slm, is

$$O \text{ slm} = L \times O_1 .$$

The slm of oxygen available in the breathing gas at the surface, or any other depth, can be calculated from the equation:

$$U = L \left[(O_1 - O_2) \left(\frac{1 - O_1}{1 - O_3} \right) \right]$$

where

U = oxygen slm that the diver can use

L = liter-flow rate, slm

O₁ = percentage (decimal) of oxygen in breathing gas

O₂ = percentage (decimal) of oxygen at 0.166 atm at depth of interest

O₃ = percentage (decimal) oxygen at 0.208 atm at depth of interest.

Example 1. A semiclosed-circuit breathing apparatus is to be used for work at a depth of 300 feet under saturated diving conditions. No changes in depth are anticipated and no decompression is required, as the diver will utilize a PTC (Personal Transfer Chamber) and a surface compression chamber between dives. Determine the optimum oxygen percentage in the breathing mixture and the liter flow required for heavy exertion. Also check the efficiency of oxygen utilization, the rate of oxygen flow as part of the breathing gas, and the rate of helium flow.

- (1) Find pressure at 300 ft - Take from Column 2, Table T-2, or calculate as follows:

$$P_{atm} = \frac{\text{Depth}}{33.1} + 1 = \frac{300}{33.1} + 1 = 10.07 \text{ atm}$$

- (2) Find oxygen percentages for oxygen partial pressures of 1.2, 0.166, and 0.208 atm at depth of 300 ft:

Take from curves of Figure 1, or calculate as follows:

$$O_1 = \frac{ppO_1}{P_{atm}} = \frac{1.2}{10.07} = 0.1182 \text{ (11.82 percent)}$$

$$O_2 = \frac{ppO_2}{P_{atm}} = \frac{0.166}{10.07} = 0.0165 \text{ (1.65 percent)}$$

$$O_3 = \frac{ppO_3}{P_{atm}} = \frac{0.208}{10.07} = 0.0206 \text{ (2.06 percent)}$$

- (3) Calculate liter-flow rate

$$\begin{aligned} L &= \frac{U}{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right)} = \frac{3}{0.1182 - 0.0165 \left(\frac{1 - 0.1182}{1 - 0.0206} \right)} \\ &= \frac{3}{0.1186 - 0.0165 \left(\frac{.882}{.989} \right)} = \frac{3}{.1182 - 0.0147} \\ &= \frac{3}{.1035} = 29.0 \text{ slm} \end{aligned}$$

- (4) Find oxygen flow in breathing gas, slm

$$O_{slm} = L \times O_1 = 29.0 \times (0.1182) = 3.43 \text{ slm}$$

- (5) Find helium flow in breathing gas

$$He = L - O_{slm}$$

$$= 29.0 - 3.43 = 25.57 \text{ slm}$$

Example 2 - Using Available Gas Mixture. Under the conditions of Example 1, a supply of gas containing 10 percent oxygen is already mixed and available. Determine the liter-flow rate for this mixture.

(1) O_1 will be taken as 0.10 (10 percent) to match available gas.

O_2 will be 0.0165 (1.65 percent), from Example 1

O_3 will be 0.0206 (2.06 percent) from Example 1

$$L = \frac{U}{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right)} = \frac{3}{0.10 - 0.016 \left(\frac{1 - 0.10}{1 - 0.0206} \right)} = \frac{3}{0.10 - 0.016 \left(\frac{0.90}{0.989} \right)}$$

$$= \frac{3}{0.10 - 0.0146} = \frac{3}{0.0854} = 35.1 \text{ slm}$$

Thus, with a 10-percent-oxygen mixture, liter flow is 35.1 slm.

Example 3 - Deep Dive From Surface. A diver must descend from the surface to carry out a mission that requires heavy exertion at depths ranging from 50 to 150 feet. What breathing-gas mixture and liter-flow rate are required?

(1) Find oxygen percentage at greatest depth (150 ft), 1.2 atm, (from Figure 1):

$$O_1 = 0.213 \text{ (21.3 percent)}$$

(2) Find oxygen percentages at least depth (at surface)

Although the mission involves work at 50 ft, the diver should be provided with sufficient oxygen for surface swimming, where oxygen vented will be 16.6 percent and oxygen inhaled will be 20.8 percent.

$$O_2 = 0.0166 \text{ (16.0 percent)}$$

$$O_3 = 0.208 \text{ (20.8 percent)}$$

(3) Liter-flow rate:

$$L = \frac{U}{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right)} = \frac{3}{0.213 - 0.166 \left(\frac{1 - .213}{1 - .208} \right)} = \frac{3}{0.213 - 0.166(0.993)} = 62.6 \text{ slm}$$

This liter-flow rate is relatively high and may limit duration of the dive to an inconveniently short period. It is required because of the small difference between oxygen percentage of gas supplied and gas vented, and it can be reduced somewhat by increasing this difference. If the time scheduled for work at the 150-foot level is less than 3 hours, the 3-hour toxicity limit of 1.3 atm oxygen partial pressure can be used to advantage. This will raise permissible oxygen concentration in the breathing mixture to 23 percent (0.23)(from Figure 1). Then,

$$L = \frac{U}{O_1 - O_2(0.993)} = \frac{3}{0.23 - 0.165} = \frac{3.0}{0.065} = 46.2 \text{ slm}$$

A second large reduction in liter flow can be obtained by simply reducing the amount of oxygen supplied for surface swimming. However, in any case, sufficient oxygen for severe exertion at 50 feet must be provided. This can be calculated as follows:

$$U_{50} = 3.0 \text{ slm}$$

$$O_2 = 0.0662 \text{ (6.62 percent)}$$

$$O_3 = 0.0829 \text{ (8.29 percent)}$$

$$\begin{aligned} L_{50ft} &= \frac{U}{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_2} \right)} = \frac{0.3}{0.23 - .0662 \left(\frac{1 - .23}{1 - .0829} \right)} \\ &= \frac{0.3}{0.23 - .0662 \left(\frac{.770}{.917} \right)} = \frac{.3}{.1744} = 17.2 \text{ slm} \end{aligned}$$

If 20 slm is provided, then the oxygen actually available for surface swimming would be:

At surface,

$$O_1 = 0.23 \text{ (23 percent)}$$

$$O_2 = 0.16 \text{ (16.6 percent)}$$

$$O_3 = 0.208 \text{ (20.8 percent)}$$

$$L = 20 \text{ slm}$$

$$U = L \left[O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right) \right] = 20 \left[0.23 - 0.166 \left(\frac{1 - .23}{1 - .208} \right) \right]$$

$$U = 20 \left[0.23 - 0.166 \left(\frac{.770}{.792} \right) \right]$$

$$= 20 [0.23 - 0.161]$$

$$= 20 [0.069] = 1.38 \text{ slm.}$$

Referring to Figure 2, 1.38 slm is sufficient oxygen for moderate exertion, and might be an acceptable condition if diving conditions are such that severe exertion should not be required. However, if oxygen for surface swimming is a limiting factor, the diver should be made fully aware of this limitation.

Example 4 - Saturation Diving Mission. In a saturation-diving mission a habitat is located at a depth of 600 feet and work is planned at levels as deep as 900 feet. Select gas consumption and flow rate appropriate for heavy exertion from 600 to 900 feet.

(1) Oxygen percentage at 900 ft, 1.2 atm, from Figure 1:

$$O_1 = 4.26 \text{ percent (0.0426)}$$

(2) Oxygen percentage at 600 ft, 0.166 atm and 0.208 atm:

$$O_2 = 0.866 \text{ percent (0.00866)}$$

$$O_3 = 1.08 \text{ percent (0.0108)}$$

(3) Liter-flow rate:

$$L = \frac{U}{O_1 - O_2 \left(\frac{1 - O_1}{1 - O_3} \right)} = \frac{3.0}{0.0426 - 0.00866 \left(\frac{1 - 0.0426}{1 - 0.0108} \right)} = \frac{3.0}{0.0426 - 0.0083} = 87.4$$

The relatively high liter-flow rate results from the low oxygen content of the breathing-gas mixture. The efficiency of oxygen utilization is 80 percent, which is high, because the depth range of 300 feet is not large in terms of pressure change.

Surface-Supplied Open-Circuit Diving Rigs

Simple open-circuit surface-supplied diving rigs are widely used for light activity and moderate depths using air as the breathing gas. The breathing air is supplied to a helmet or mask within which the diver breathes normally. The helmet acts as a ventilated dead space within which breathing air is mixed with carbon dioxide exhaled by the diver. The flow of air must be great enough to dilute the exhaled carbon dioxide to non-toxic levels for rebreathing, in the range of 0.01 to 0.02 atm partial pressure. The rate of ventilation required is defined by the following equation:

$$\text{Equation 1: } V = \frac{P_{\text{atm}} \times O_{\text{slm}} \times R \times F}{26.2 (C_2 - C_1 \times P_{\text{atm}})}$$

where

V = Volume of air required, in scfm

O_{slm} = Oxygen requirement, in slm

R = Respiratory quotient, or ratio: Vol CO_2 produced/Vol O_2 consumed

26.2 = Conversion factor, slm to scfm

C_2 = Desired partial pressure CO_2 in inhaled air (atm)

C_1 = Partial pressure CO_2 in breathing air from compressor (atm)

P_{atm} = Pressure at depth, expressed in atm, = $\frac{D + 33.1}{33.1}$

F = Mixing Effectiveness Factor.

The following values can be used in almost all operations:

O_{slm} = 2.6 slm (enough for heavy to severe exertion)

R = 0.9 (the highest value likely to occur)

C_2 = 0.02 atm, toxic limit of carbon dioxide.

$C_2 = 0.02$ atm, toxic limit of carbon dioxide.

$F = 1.0$ assumed because no actual values are now known.

If these values are substituted in Equation 1, Equation 2, below results:

$$\text{Equation 2: } V = \frac{0.0893 P_{\text{atm}}}{0.02 - C_1 P_{\text{atm}}}$$

If the breathing air contains no carbon dioxide this equation reduces to:

$$\text{Equation 3: } V = 4.5 P_{\text{atm}}$$

F , the Mixing Effectiveness Factor, is a factor to account for the mixing of exhaled carbon dioxide with breathing gas within the helmet. It is the ratio (% CO_2 in vented air)/(% CO_2 in inhaled air). The value of F would vary under the following conditions:

- $F = 1$ When mixing is perfect, and the vented mixture contains the same percentage of carbon dioxide as the inhaled mixture.
- $F < 1$ (F less than 1.) When mixing is intentionally imperfect and flow patterns are directed so that the inhaled mixture contains less CO_2 than the vented mixture.
- $F > 1$ (F more than 1.) When mixing is imperfect and, through poor flow patterns, the inhaled mixture contains more CO_2 than the vented mixture.

The only way of evaluating F for a specific helmet design is to measure the carbon dioxide concentration in the exhaust and that at the point of inhalation while a diver wears the helmet and performs hard work. In the absence of such experimental data, it appears reasonable to assume that $F = 1$.

The Diving Manual⁽²⁾ recommends that flow rate be set at $4.5 P_{\text{atm}}$. This rate will supply only enough ventilation for 2.6 slm oxygen consumption. Thus, there may be occasions when flow rate should be higher than $4.5 P_{\text{atm}}$.

Example 1: How much air should be supplied to a diver working at a depth of 70 feet? Assume that air contains no carbon dioxide.

Procedure	Example
1. Find P_{atm} (From Table T-2, Col 3)	3.12 atm
2. Use Equation 2: $V = 4.5 P_{\text{atm}}$	$V = 4.5 (3.12) = 14.05$ scfm

Example 2: A diver is to work at a depth of 150 ft. Calculate the amount of air he will need under three different conditions of air purity, as follows:

- (a) Pure air containing no carbon dioxide
- (b) Air cleaned to just meet Bu. Med. air purity standards of 0.05 percent (0.0005 atm) carbon dioxide

- (c) Unpurified air from a compressor in a machinery space, containing 0.2 percent (0.002 atm) carbon dioxide

Procedure	Example
1. Find P_{atm} (From Table T-2, Col 3)	5.53 for all conditions
2. Use Equation 2 to find ventilation rate:	
$V = \frac{0.0893 P_{atm}}{0.02 - C_1 P_{atm}}$	
Condition (a)	$V = \frac{0.0893(5.53)}{0.02} = 24.7 \text{ scfm}$
Condition (b)	$V = \frac{0.0893(5.53)}{0.02 - 0.0005(5.53)} = 28.8 \text{ scfm}$
Condition (c)	$V = \frac{0.0893(5.53)}{0.02 - 0.002(5.53)} = 53.3 \text{ scfm}$

Example 2 shows the relative effects of carbon dioxide in breathing air. When carbon dioxide level is at the maximum partial pressure permitted by Bu. Med, Air Purity Standards, Condition (b) air flow must be increased 25 percent above that for air containing no carbon dioxide. At the much higher level of 0.2 percent carbon dioxide, the required air flow is more than doubled.

The U. S. Navy Bureau of Medicine and Surgery has established purity standards for diver's breathing air which should be met by all apparatus used to supply breathing air. Requirements are as follows:

Oxygen	20 to 22 percent by volume
Carbon dioxide	300 to 500 ppm (0.03 to 0.05 percent) by vol.
Carbon monoxide	20 ppm maximum
Oil, mist, and vapor	5 mg/m ³ maximum
Water, maximum content	Saturated at maximum pressure of use
Solid and liquid particles	Not detectable except as noted above under oil, mist, and vapor.
Odor	Not objectionable

In order to meet the above specifications it is necessary to use special non-lubricated compressors, or to pass air from standard compressors through a highly efficient cleanup system. The compressed air provided for general shipboard services is not adequate for use as breathing air, as it usually contains excessive amounts of oil and may contain carbon dioxide if taken from machinery spaces.

Helium-Oxygen Deep-Sea Diving Outfit

The helium-oxygen deep-sea diving outfit is a surface-supplied diving outfit that operates as a semiclosed system with recirculation of breathing gas and carbon dioxide absorption. In principle, the helmet is a ventilated dead space, requiring the usual high volume flow rate for dilution of carbon dioxide. However, this high flow rate is obtained by introducing breathing gas at high velocity in an injector that recirculates helmet gas through a Baralyme canister attached to the helmet. The volume of gas recirculated is about ten times the volume of breathing gas admitted, which reduces the flow of breathing gas needed for helmet ventilation by a factor of ten below that for a system without recirculation.

The oxygen content of the breathing gas is adjusted and controlled at the surface and can be varied to suit depth and decompression requirements. The gas pressure at the inlet to the diver's ejector is also set at the surface and determines the flow rate of breathing gas, which enters the ejector through a 0.025-inch diameter nozzle.

In setting oxygen percentages and flow rates for breathing gas it is necessary to assure that three conditions are met:

- (1) The oxygen percentage should be below the toxic limit for the scheduled diving period (1.2 to 1.8 atm under usual conditions), and high enough to avoid anoxia at high levels of exertion (above 0.37 atm for 3 slm oxygen consumption)
- (2) The flow rate should be high enough for adequate ventilation of carbon dioxide within the helmet; with ejector recirculation ratio of 10:1 a flow rate of $(0.5 \times P_{atm})$ scfm, or $(14.2 \times P_{atm})$ slm is required.
- (3) Pressure drop across the helmet ejector should be high enough to assure critical flow. This requires that gas pressure at the helmet be at least twice the bottom pressure. (For depths greater than 190 ft, supply pressure will be more than that recommended in the Diving Manual of 100 psi above bottom pressure.)

The supply pressure should include an allowance for hose pressure drop, which may be significant. However, if the correct flow rate is metered into the hose the pressures should automatically adjust to the required level.

The duration of a dive when using this outfit is the effective life of the Baralyme carbon dioxide absorbent. The design life of a full canister of 6 pounds is 3 hours⁽²⁾, but in cold water this can be drastically reduced and may be as little as 1/2 hour at 40 F⁽⁸⁾.

Calculation of Breathing-Gas Oxygen Content and Flow Rate

With the breathing-gas flow rate of 0.5 cfm at depth established, the oxygen content of the breathing gas and the required flow rate, in scfm or in lb/hr, can be calculated as follows:

<u>Procedure</u>	<u>Example</u>
(1) Conditions	
Working depth	300 ft
Water temperature	60 F
Length of dive - time at depth	30 min
(2) Select oxygen percentage in mixture	
Use oxygen partial pressure safe for time limit	1 hr
From Figure 1, select oxygen percentage at partial pressure of 1.5 atm	15 percent
(3) Select gas-flow rate:	
$V = 0.5 \times P_{atm}$	$V = 0.5 \times 10.07 = 50$ scfm
where	
V = volume flow rate, scfm	
P_{atm} = pressure at depth, atm (from any table at 300 ft);	
or, for gas flow in lb/hr:	
$W = 0.5 \times 60 \times \rho$,	$W = 0.5 \times 60 \times 0.1279$ $= 3.84$ lb/hr
where	
W = lb/hr breathing gas	
60 = minutes per hour	
ρ = gas density, from Table 56 at 300 ft, 60 F	
(4) Select supply pressure to diver	
The minimum supply pressure will be:	
$p_s = 2(P) - 15 + p_h$	$p_s = 2(148) - 15 + 3 = 284$ psi
where	
p_s = supply pressure, psi	
P = bottom pressure, psia (from Table T-2, Col. 2 at 300 ft)	
p_h = hose pressure drop, psi (estimated)	
(5) Similarly, mixtures and flow rates for reduced depths should be determined and used during the decompression period to assure that adequate oxygen percentages are used at the shallow depths during decompression.	

MIXING OF BREATHING GASES

Mixtures of oxygen and helium, or oxygen and nitrogen, in various proportions are used for different diving missions. These mixtures may be prepared by one of several methods.

In shore-based facilities, it is feasible to mix gases at ambient pressure by adding suitable volumes of each gas to a constant-pressure gas holder, measuring the volume added for each gas. The mixture can then be analyzed and the composition adjusted as needed, following which the gas can be compressed into high-pressure cylinders for storage or shipment. With this approach, precise proportioning of gases in the mixture is readily obtained.

Where small, portable cylinders are to be filled, it is sometimes feasible to mix gases by weight. It is necessary to calculate the weight of gas mixture needed to fill the cylinder to the desired pressure, the percentage by weight of each constituent gas, and the weight of each gas to be added. The cylinder is then placed on a precise scale and the required weights of gas added. Where applicable, this method is free of temperature effects and can provide good mixture accuracy. In filling the cylinder, gas is taken from larger, high-pressure cylinders. If necessary, in order to obtain the desired high cylinder pressure, the final increment of gas may be pumped into the cylinder with a high-pressure compressor. Manual compressors are available for this purpose. This method requires a scale of high accuracy because gas weights are quite small. It is probably not suitable for shipboard use because of ship motion.

Still another approach to gas mixing is use of calibrated flow meters in both oxygen and diluent-gas lines entering a mixing device. Continuous flow and mixing in proper proportions can be provided by selection and adjustment of the metering equipment. The mixed gas may be supplied directly to the diver if pressure is adequate, or it may be compressed for storage in high-pressure cylinders or for supply to a diver.

The fourth approach to mixing of diving gases consists of adding each of the constituent gases to a storage cylinder in turn, the proportion by volume of each gas being in direct proportion to the partial pressure of that gas in the mixture. For example, a mixture containing 10 percent oxygen and 90 percent helium at a total pressure of 1000 psi could be prepared by filling the cylinder by adding oxygen until pressure was 100 psi, then adding helium until the pressure reached 1000 psi. Because of the practice of carrying helium and oxygen in high-pressure gas cylinders aboard ships, this method is used frequently.

The procedures for carrying out each of these mixing methods are discussed in detail in the U. S. Navy Diving Manual. (2) The calculation procedures required to implement the methods are outlined below. Calculation procedures for helium-oxygen mixtures are based on use of the tables presented in this manual; alternative calculation methods based upon perfect-gas laws (PVT relations) are presented for comparison. Calculations based on densities and specific volumes of oxygen, helium, and mixtures taken from the tables include real-gas compressibility effects, and would be expected to provide results more accurate than those obtained by calculations based on perfect-gas laws. However, they are considerably more complex than perfect-gas calculations.

Mixing by Weight - Two or More Gases
(Real-Gas Method)

In mixing by weight it is necessary to know the receiver volume and weight, the final pressure, the temperature at which the receiver is to be filled, and the gaseous constituents of the mixture and their proportions. From these it is possible to calculate the weight of each gas to be added to the receiver. The procedure is outlined in detail below:

<u>Condition</u>	<u>Example 1</u>	<u>Example 2</u>
Volume of Receiver, ft ³	0.5	0.5
Weight of Receiver (Empty), lb	30.0	30.0
Final Gas Pressure, psi	2200	2200
Final Gas Temperature, F	70	70
Gas Composition, volume percent		
Oxygen, Mol Wt 32	10	10
Helium, Mol Wt 4	90	88.5
	0	1.5
	<u>100</u>	<u>100</u>

Example 1. Oxygen-Helium Mixture

Oxygen (from Table T-1), wt %	47
Helium (by Difference), wt %	53
Gas Density After Filling, lb/ft ³ (from Table T-63, 2200 psi, 70 F)	2.45
Total Gas Weight, lb (Gas Density) x (Receiver Vol)	2.45 x 0.5 = 1.225
Weight of Oxygen to be Added, (Weight % Oxygen) x (Total Gas Weight)	47/100 x 1.225 = 0.524 lb or 8.4 oz
Weight of Helium to be Added (Weight % Helium) x (Total Gas Weight)	53/100 x 1.225 = 0.649 lb or 10.4 oz
Initial Tank Weight, lb	30.0
Weight After Adding Oxygen, lb	30.524, or 30 lb 8.4 oz
Weight After Adding Helium, lb	31.225, or 31 lb 3.6 oz

Example 2. Oxygen-Helium-Nitrogen Mixture
(Real-Gas Method)

For this mixture of three gases it is necessary to calculate the percentages by weight of the gases and the density of the mixture from perfect-gas laws. With these known, the weights to be added are calculated as in Example 1.

Mean molecular weight of mixture:

$$100 \times \text{mol wt mix} = (\text{mol wt O}_2 \times V\% \text{ O}_2) + (\text{mol wt He} \times V\% \text{ He}) + (\text{mol wt N}_2 \times V\% \text{ N}_2)$$

$$= (32 \times 10) + (4 \times 88.5) + (28 \times 1.5)$$

$$\text{Mol wt} = \frac{320 + 354 + 42}{100} = 7.16$$

Calculate specific volume and density:

Specific volume is calculated from the perfect-gas relation $PV = RT$, where

$$P = \text{absolute pressure, lb/ft}^2 = (\text{gage pressure, psi} + 14.7) \times 144$$

$$T = \text{absolute temperature, R} = (\text{degrees F} + 460)$$

$$R = 1545 / (\text{mean molecular weight of gas})$$

$$V = \text{specific volume, ft}^3/\text{lb.}$$

Then, if $PV = RT$, $V = RT/P$, where

$$R = 1545 / \text{mol wt} = 1545 / 7.16 = 217$$

$$V = RT/P = (217)(70 + 460) / (2200 + 15) \times 144 = 0.386 \text{ ft}^3/\text{lb.}$$

$$\text{Density is } 1/V = 1/(0.386) = 2.59 \text{ lb/ft}^3.$$

Total weight of gas added to receiver:

$$(\text{receiver volume}) \times (\text{gas density}) = (0.5 \text{ ft}^3) \times (2.59 \text{ lb/ft}^3) = 1.295 \text{ lb}$$

From this point the calculation is the same as that for Example 1; the weight percentage of each gas is found and is multiplied by the total weight to obtain the weight of each gas.

Although the total weight in the tank is calculated by perfect-gas laws, the final mixture proportions are not affected by compressibility effects; only the final pressure may be somewhat low. Accordingly, the mixture can be assumed to be correct under real-gas conditions.

The weight of gases to be used for refilling partially filled tanks can also be calculated by the weight method. This is done by calculating the weight of the gas to fill the tank at full pressure, and subtracting the weight of gas in the tank at the existing pressure. Both these calculations are done as described above, and the difference is the weight to be added. The added weight is proportioned in the same way as when filling an empty tank.

If it is desired to change the mixture ratio in a filled tank, or in a partially filled tank, the above method can be used to calculate the weight of each gas in the tank and the weight of each gas desired in the tank. The desired difference in weight of one gas is then added to the tank. If the tank is full, it will be necessary to bleed off some

mixture in order to adjust composition. These methods (filling a partially filled tank, or adjusting mixture) depend upon knowing very accurately the weight of the tank when empty. Without this knowledge, the partial pressure method described below can be used to perform the calculations.

It is evident that the scale used for proportioning gases by the weight method must be suitable for measuring small weight changes with high accuracy. The error in gas composition involved in a 1/2-oz. error in oxygen weight in Example 1, for example, would be about 6 percent. That is, if the oxygen weight added was 8.9 oz instead of 8.4 oz, the volume percentage of oxygen in the mixture would be 10.6 instead of the target value of 10.0. This method is probably unsuitable for shipboard use because ship motion complicates an already-difficult weighing problem.

Mixing by Partial Pressures

When two or more ideal gases are mixed at constant temperature, the partial pressure of each gas in the mixture is proportional to the percentage by volume of that gas in the mixture, and the sums of the partial pressures of all gases must add to the total pressure. Thus, it is possible to prepare a gas mixture in a high-pressure storage cylinder by adding, in series, several gases. As each gas is added, the cylinder pressure will rise and the pressure, as a percentage of the final pressure, will be proportional to the percentage by volume of the gas.

When real gases such as oxygen and helium are mixed, the mixture compresses slightly more or less than predicted by perfect-gas laws. The differences in partial pressures resulting from real-gas effects are quite small at the pressures usual for diving and can be neglected for most purposes. However, they can be significant at the high pressures used in mixing and storing of breathing gases, and can cause significant error in the percentage of oxygen in a mixture. Accordingly, the calculation methods recommended in this manual are based on real-gas properties as taken from the tables in Part 2. However, perfect-gas calculations are also shown for use when applicable, and for comparison of results.

All mixing calculations assume isothermal conditions throughout the mixing period, with initial and final temperatures equal. In actual practice, if mixing is done quickly the adiabatic compression in the receiving vessel will cause a temperature rise that can result in mixing error. This effect can be minimized by adding the gas to the receiver at low flow rates so that the gas temperature can come to equilibrium with vessel metal temperatures, or by waiting for temperature to return to ambient and then topping off the receiver to proper pressure.

A highly accurate pressure gage having many scale divisions, and of suitable scale range, is required for gas mixing. For example, if a gas mixture containing 10 percent oxygen is to be mixed within an accuracy of ± 0.1 percent oxygen (9.9 to 10.1), and the partial pressure of the oxygen is 200 psi, then the gage accuracy must be ± 2 psi. If, in this case, a 3000-psi gage having 20 psi increments were used, and the actual oxygen pressure was in error by +10 psi, the oxygen content of the mixture would be higher than calculated, at 10.5 percent.

Although the receiving vessel will contain the proper amounts of oxygen and helium when filled by the partial-pressure method, these gases may not be well mixed.

At the high pressures used for gas storage molecular diffusion is very slow, and gas drawn from the tank immediately after mixing could be quite different in composition than the average for the entire receiver. Accordingly, it is desirable to provide for a storage period of many hours before use, if possible, or provide for mechanical agitation to improve the rate of mixing. (2)

When mixing low-oxygen, helium-oxygen mixtures, best accuracy can be obtained by adding oxygen to the receiver first, using a low-pressure gage to measure its partial pressure, and then adding helium to the final mixture pressure. Where oxygen percentage is high, or where helium is already in the receiver, oxygen can be added after the helium, and it is necessary to know the partial pressure of helium and the final mixture pressure for the desired mixture.

When adding oxygen first, the density of oxygen at its partial pressure in the mixture must be known. This can be found by solving the following equation:

$$\rho_{O_2} = \frac{32 a \rho_m}{28 a + 4}$$

where

ρ_{O_2} = density of oxygen, lb/ft³, at its partial pressure in the mixture

ρ_m = density of the mixture at final pressure (from tables)

a = decimal percentage oxygen in mixture

32 = molecular weight of oxygen

4 = molecular weight of helium

28 = difference in molecular weight of oxygen and helium.

The partial pressure of oxygen corresponding to this density is found in oxygen density tables T-7 and T-8, of Part 2 by locating the computed density in the proper temperature column and finding the corresponding tabulated pressure in Column 2. In preparing the mixture, oxygen is added to the receiver until this pressure is reached, and then helium is added until the final mixture pressure used in the calculation above is reached. This must all be done at constant gas temperature, and the procedures recommended in the Navy Diving Manual should be followed in detail. Essentially, the receiver is filled slowly, allowed to cool until pressure is stabilized, then topped off to the desired partial pressure.

If helium is to be added to the mixture first, the helium density at its partial pressure in the mixture can be found using the equation:

$$\rho_{He} = \frac{4b \rho_m}{32 - 28b}$$

where

ρ_{He} = density of helium

b = decimal percentage of helium in mixture and other terms are the same as in the oxygen equation.

Helium partial pressure would be taken from the helium density Table T-14 or T-15.

This real-gas mixture method cannot be used where a third gas such as nitrogen is present, as the nitrogen would affect the compressibility of the mixture and cause error.

It is sometimes advisable to check the pressure drop that occurs in gas-storage cylinders used to fill the receiver, to assure that the desired partial pressure level can be obtained with the gas on hand. The pressure drop in the supply vessel can be calculated using the following equation:

$$\text{Supply } \Delta P = \frac{(\text{receiver volume}) \times (\text{receiver } \Delta P)}{(\text{supply-vessel volume})}$$

where

Supply ΔP = pressure drop in supply vessel, psi

Receiver ΔP = pressure rise in receiver, psi

Volumes = ft³

If gas is taken from the supply vessel at a high rate of pressure change, the adiabatic cooling with expansion of gas in the vessel will lower temperature, and reduce pressure. Accordingly, a margin of excess pressure should be available in the supply vessel. On standing, the gas temperature and pressure will rise, so that some additional pressure should be available for topping off the receiver after it has cooled.

Example 1. Mixing by Filling an Empty Receiver

A mixture containing 10 percent oxygen and 90 percent helium is to be prepared by filling a receiver containing the same mixture at atm pressure. Assume that ambient temperature is 70 F. Find the partial pressure of oxygen to be used in preparing a mixture having a pressure of 3000 psi.

Example 1

Terms

P_1 = Initial mixture pressure, psia
 P_2 = Final mixture pressure, psia
 t = Temperature, F
 a = Decimal fraction (percent) of oxygen in mixture
 ρ_m = Density of mixture
 ρ_{O_2} = Density of oxygen in mixture at final pressure
 ΔppO = Change in oxygen partial pressure
 ppO_2 = Partial pressure of oxygen at final pressure

Example

14.7
 3000 (2985 psi)
 70
 0.10

Procedure - Real-Gas MethodExample(1) Find ρ_m - in Table T-63 for 3000 psi, 70 F3.258 lb/ft³(2) Find ρ_{O_2} :

$$\rho_{O_2} = \frac{32a(\rho_m)}{28a + 4}$$

$$= \frac{32(0.10)(3.258)}{28(0.10) + 4} = 1.534$$

(3) Find ppO₂ in Table T-8 at $\rho_O = 1.534$, t - 70 F
(psia, in Column 2)

Interpolation: (use straight-line interpolation between points in tables)

ppO falls between 272.5 psia at $\rho = 1.5534$ lb/ft³and 268.0 psia at $\rho = 1.5279$

$$\frac{4.5 \text{ psi}}{.0256}$$

$$\rho_O = 1.534$$

$$\rho \text{ at } 268 \text{ psia} = \frac{1.528}{.006}$$

$$4.5 \text{ psi} \times \frac{.006}{.0256} = 1.75 \text{ psia}$$

$$\text{ppO}_2 = \frac{+268.00}{269.75} \text{ psia}$$

(4) ppO₁ = P₁ a₁

$$(14.7)(0.10) = 1.47 \text{ psi}$$

(5) ppO₂ - ppO₁ = Δ ppO

$$(269.75) - (1.47) = 266.3 \text{ psi}$$

(6) PO₂ = P₁ + Δ ppO

$$(14.7) + (266.3) = 281.0 \text{ psi}$$

Thus, the gage pressure after adding oxygen should be 281 psi

If the receiver had contained a previous mixture of different oxygen content, or air, the value of a₁ in Step (4) would be that for the old mixture. Step 4 is a perfect-gas method, but introduces no error at low pressures.

Procedure - Perfect-Gas Method

$$\Delta \text{ppO} = (P_2 - P_1)(a)$$

$$(3000 - 15)(0.10) = 298.5 \text{ psia}$$

$$P_O = P_1 + \Delta \text{ppO}$$

$$(14.7) + (298.5) = 313.2 \text{ psi}$$

With final oxygen pressure of 313 psi, the mixture would have contained 11.2 percent oxygen instead of the specified value of 10.0 percent. This illustrates the desirability of using real-gas properties.

**Example 2. Refilling A Partially Filled Tank
With the Same Mixture or a Different Mixture**

A tank contains a mixture of 10 percent oxygen and 90 percent helium at a pressure of 1200 psi. It is to be refilled with the same mixture at 3000 psi, using pure oxygen and pure helium. Assume a temperature of 70 F. What should the tank pressure be after addition of new oxygen?

Example 2

<u>Terms</u>	<u>Example</u>
t = Mixing temp	70 F
P ₁ = Initial mixture pressure, psia	1200
P ₂ = Final mixture pressure, psia	3000
ρ _{m1} = Initial mixture density	
ρ _{m2} = Final mixture density	
PPO ₁ = Oxygen partial press, initial mixture	
PPO ₂ = Oxygen partial press, final mixture	
ρ _{O₁} = Oxygen density, initial mixture	
ρ _{O₂} = Oxygen density, final mixture	
a ₁ = Decimal fraction of oxygen in initial mixture	0.10
a ₂ = Decimal fraction of oxygen in final mixture	0.10

Procedure - Real-Gas Method

- | | |
|--|--|
| (1) Find ρ _{m1} in Table T-63 at 1200 psi, 70 F | 1.379 |
| (2) Find ρ _{m2} in Table T-63 at 3000 psi, 70 F | 3.258 |
| (3) Find ρ _{O₁} | |
| $\rho_{O_1} = \frac{32a \rho_{m1}}{28a + 4} =$ | $\frac{32(0.10)(1.379)}{28(0.10) + 4} = \frac{5.725}{6.8} = 0.832$ |
| (4) Find ppO ₁ in Table T-8 at ρ _{O₁} = 0.832, t = 70 F | 147.9 psia |
| (5) Find ρ _{O₂} | |
| $\rho_{O_2} = \frac{32a \rho_{m2}}{28a + 4}$ | $\frac{32(0.10)(3.258)}{28(0.10) + 4} = \frac{10.44}{6.8} = 1.536$ |
| (6) Find ppO ₂ in Table T-8 at ρ _{O₂} = 1.536, T = 70 F | 272.5 psia |
| (7) Find change in oxygen partial pressure | |
| $\Delta P_O = ppO_2 - ppO_1$ | 272.5 - 147.9 = 124.6 psia |

- (8) Find pressure after adding new oxygen

$$P_{m1} + \Delta P_O$$

Change to gage pressure.

$$\begin{array}{r} 1200 + 124.6 = 1324.6 \text{ psia} \\ -14.7 \\ \hline 1310 \text{ psi} \end{array}$$

- (9) Find pressure after adding new helium

Change to gage pressure.

$$\begin{array}{r} 3000 \text{ psia} \\ -15 \\ \hline 2985 \text{ psi} \end{array}$$

Procedure - Perfect-Gas MethodExample

$$(1) \Delta pp_O = (P - P_1)a = (3000 - 1200)(-10) = 180 \text{ psi}$$

(Δpp_O = change in oxygen
partial pressure)

$$(2) P_1 + \Delta pp_O = 1200 + 180 = 1380 \text{ psia} = \text{pressure after adding new oxygen}$$

Change to gage pressure

$$\begin{array}{r} -15 \\ \hline 1365 \text{ psi} \end{array}$$

In this example the calculated partial pressure of oxygen to be added was 124 psi using the real-gas method, and 180 psi using the perfect-gas method. The actual oxygen percentage of the gas mixture added would be 10.0 for the real-gas method, and 14.4 for the perfect-gas method.

If the oxygen content of the initial mixture had been different from that of the final mixture the problem would follow the same steps but a_1 and a_2 would have different values, and ρ_{O_1} and ρ_{O_2} would have been found in different tables.

Example 3. Changing Oxygen Content of Mixture
by Adding Pure Oxygen

It is possible to use the partial-pressure method to change the oxygen content of a mixture by adding oxygen. The increase in pressure from adding oxygen can be calculated using the same real-gas method used in Examples 1 and 2, but it is also necessary to know weight percentages to calculate the oxygen pressure rise. Accordingly, the calculation method is carried out by converting to a percentage-by-weight basis first, then proceeding with calculations on a weight basis.

Condition and TerminologyExample

a_1 - Decimal volume percentage of oxygen in
Mixture 1

0.10 (10%)

a_2 - Decimal volume percentage of oxygen in
Mixture 2

0.32 (32%)

b_1, b_2 - Decimal volume percentage helium
in Mixtures 1 and 2

P_1 - Pressure of Mixture 1

900 psia

P_2 - Pressure of Mixture 2

t - Temperature

70 F

- w_{O_1} - Decimal weight fraction of oxygen in Mixture 1
 w_{O_2} - Decimal weight fraction of oxygen in Mixture 2
 w_{He_1} and w_{He_2} - Decimal weight fraction of helium
 ρ_{m_1} , ρ_{m_2} - Mixture densities
 ρ_{O_1} , ρ_{O_2} , ρ_{He} - Densities of oxygen and helium in mixtures
 P_{PO_1} , P_{PO_2} , P_{PHe} - Partial pressures of oxygen, helium

Procedure - Real-Gas Method

Example

- (1) Convert volume mixture percentage to weight percentages (Use Table T-1 for mixtures tabulated there)

	<u>By Vol</u>	<u>By Weight</u>
Initial mix: O ₂ =	10%	47.1
He =	90%	52.9
Final mix: O ₂ =	32%	79.0
He =	68%	21.0

- (2) Find ρ_{m_1} in Table T-63 at 900 psi, 70 F

1.045

- (3) Find ρ_{O_1} , ρ_{He_1}

$$\rho_{O_1} = \rho_{m_1} \times w_{O_1} = (1.045)(47.1) = 0.492$$

$$\rho_{He} = \rho_{m_1} \times w_{He_1} = (1.045)(52.9) = \frac{0.553}{1.045}$$

- (4) Find ρ_{O_2} , ρ_{m_2}

$$\rho_{He_2} = \rho_{He_1}, \text{ - because no helium is added}$$

$$\rho_{O_2} = \frac{w_{O_2} \times \rho_{He_2}}{w_{He_2}} = \frac{(79)}{(21)} (0.553) = 2.085$$

- (5) Find ρ_{m_2}

2.640

$$\rho_{m_2} = \rho_{He_2} + \rho_{O_2} = (0.553) + (2.085) =$$

- (6) Find P_{m_2}

From Table T-45 for 32 percent oxygen,
at $\rho_{m_2} = 2.640$, $t = 70$ F

1201.4 psia

- (7) Change to gage pressure.

-14.7

Thus, final pressure after adding oxygen = 1186.7 psi

Procedure - Perfect-Gas Method

(1) $pp_{O_1} = P_1 \times a_1$

(2) $pp_{He} = P_1 \times b_1$

(3) $pp_{He_2} = pp_{He_1}$

(4) $P_2 = \frac{pp_{He_1}}{b_2} =$

(5) Change from psia to psi

(6) Check

$pp_{O_2} =$

$\Delta pp_O =$

$P_2 = P_1 \times \Delta pp_O =$

Example

$900 \times 0.10 = 90 \text{ psi}$

$900 \times 0.90 = 810 \text{ psi}$

810 psi

$\frac{810}{0.68} = 1190 \text{ psia}$

$\begin{array}{r} -15 \\ \hline 1175 \text{ psi} \end{array}$

$1190 \times 0.32 = 381 \text{ psi}$

$381 - 90 = 291 \text{ psi}$

$900 + 291 = 1191 \text{ psi}$

The difference in final pressure calculated by the two methods is 10 psi, and this would result in a final gas mixture containing 31.2 percent oxygen when using the perfect-gas method. The difference is fairly small because of the moderate pressure of 1200 psi.

Example 4. Mixing of Two Mixtures to Make a Third Mixture.

It is possible to mix an oxygen-rich mixture with an oxygen-lean mixture to obtain any desired oxygen content between the concentrations in the two starting mixtures, using the partial-pressure method. Although the mixing procedure is exactly the same as for mixing two pure gases directly, the calculations for determining the proportions of each mixture that will result in the correct final oxygen concentration are more complicated than for mixing two pure gases. Two simultaneous linear equations are used to determine the volume proportions of each gas mixture to use. With these proportions known, the partial pressure for each gas mixture is carried out by converting to weight percentages, then carrying out the usual mixing calculation.

Example 4Conditions and Terminology $a_1 =$ decimal fraction of oxygen in Gas 1 $a_2 =$ decimal fraction of oxygen in Gas 2 $a_3 =$ decimal fraction of oxygen in Gas 3 $P_3 =$ pressure of Gas 3, psia $X =$ decimal fraction of Gas 1 in Gas 3 $Y =$ decimal fraction of Gas 2 in Gas 3 $Z =$ Gas 3 = 1.0 decimal fractionExample

0.10 (10 percent)

0.40

0.15

2200

Mol Wt = Molecular Weight

ρ = density

pp = partial pressure

w_1, w_2 = weight fractions of Gas 1 and Gas 2 in Gas 3

Procedure

- (1) Solve for percentages of Gas 1 and Gas 2 in Gas 3

Equation (1): $X + Y = 1$, and $Y = 1 - X$

Equation (2): $a_1X + a_2Y = a_3$ (1) .

Substitute Equation (1) in Equation (2) and solve for X:

$$a_1X + a_2(1 - X) = a_3 .$$

This reduces to:

$$X = \frac{a_3 - a_2}{a_1 - a_2} .$$

Substituting values:

$$X = \frac{0.15 - 0.40}{0.10 - 0.40} = \frac{-0.25}{-0.30} = 0.833 \text{ (83.3 percent)}$$

$$\text{Equation 1: } Y = 1 - X = 1.00 - 0.833 = 0.167 \text{ (16.7 percent)}$$

Therefore, Gas 3 will contain 83.3 percent of Gas 1 and 16.7 percent of Gas 2.

Example

Procedure - Real-Gas Method

- (2) Convert volume fraction to weight fraction. Find mol wt. in Table T-1

<u>(Vol Fract), x (Mol Wt)</u>	<u>Weight Fraction, w</u>
(Gas 1) $0.833 \times 6.80 = 5.67 + 8.21 =$	0.6906
(Gas 2) $0.167 \times 15.20 = 2.54 + 8.21 =$	0.3094
<u>8.21</u>	<u>1.0000</u>

- (3) Find ρ_3 (density of Gas 3 at 2200 psia, 70 F) in Table T-57

$$\rho_3 = 2.951 \text{ lb/ft}^3$$

- (4) Find ρ_1 and ρ_2

$$\rho_1 = \rho_3 w_1 = (2.951) \times (0.6906) = 2.038 \text{ lb/ft}^3$$

$$\rho_2 = \rho_3 w_2 = (2.951) \times (0.3094) = .913 \text{ lb/ft}^3$$

$$\underline{\quad\quad\quad} 2.951 \text{ lb/ft}^3$$

(5) Find partial pressures pp_1 and pp_2 from Tables

pp_1 - (Table T-63 - at $\rho = 2.038$): 1808 psia (0 to 1808)

or

pp_2 - (Table T-39 - at $\rho = 0.913$): $\frac{344}{2152}$ psia (0 to 344)
(2152 psia)

The tank should be pressurized to 344 psi with Gas 2, then brought to 2200 psi with Gas 1, so that a low-pressure gage can be used for good accuracy with Gas 2. Alternatively, the tank can first be pressurized to 1808 psi with Gas 1, and Gas 2 can then be added to bring tank to 2200 psi. It will be noted that the sum of 1808 and 344 is not 2200. This is because of gaseous interactions (compressibility effect) and does not indicate an error in the computation.

This calculation neglected any existing gas in the tank before filling. If at 1 atm pressure, the effect would be negligible, and all absolute pressures simply become differential pressure for gas additions without changing absolute pressures to gage pressures.

Procedure - Perfect-Gas Method

Example

$$pp_1 + pp_2 = P_3 = 2200 \text{ psia}$$

$$pp_1 (X) + pp_2 (Y) = 2200$$

$$pp_1 = P_3 \times X$$

$$2200 \times 0.833 = 1832.6 \text{ psi}$$

$$pp_2 = P_3 \times Y$$

$$2200 \times 0.167 = \underline{367.4} \text{ psi}$$

$$pp_1 + pp_2 = P_3$$

$$2200.0 \text{ psi}$$

Comparison with the real-gas method shows pp_2 at 367 and 344 psi, a difference sufficient to change the percentage of Gas 2 in the mixture from the desired value of 16.7 percent to about 17.8 percent, and the oxygen content of the mixture from 15 percent to about 15.3 percent.

Continuous-Flow Mixing Systems

The third basic method of preparing diving-gas mixtures is by metering and controlling separate streams of oxygen, helium, and air or nitrogen, then mixing them at high turbulence downstream from the metering system. The mixture can be supplied directly to a diver, or recompressed and stored for future use. In theory, any method of flow measurement could be used in such a system if properly applied and used. In practice, however, it is highly desirable to provide a mixing system that is simple to operate and that does not require extensive calculation and skill to obtain correct mixtures. Finally, since the mixture may flow to the diver continuously as prepared, a continuous-flow mixture analyzer appears essential to detect any malfunction in the system.

Figure 4A is a schematic diagram showing principles of the system used in an Airco Mixmaker, a continuous-flow diving-gas mixing system in service at EDU*. Several features of this system are of interest, and indicate typical means of simplifying and automating the measurement function. Stored pure gases, at the left of the diagram, are supplied to pressure regulators (1), at a pressure of 1100 psi or more. Gas at constant regulated pressure is then heated to a specified and controlled temperature in the heaters (2), so that the flow meters (3) always measure gas at the same pressure and temperature. This eliminates the need for computing temperature and pressure corrections. After passing through the flowmeters the three gas streams are mixed in a turbulent mixing chamber (4), then passed through a regulator (5) that controls the flowmeter-outlet pressure at 750 psi. A sample of the mixed gas is withdrawn through line (6) and passed through a recording gas analyzer. The main gas stream passes through regulator (7), where its pressure is regulated at the value needed to supply the diver. An accumulator could be added at (9) for mixed-gas storage. With the above pressure levels, the rate of flow of the mixture is 6 cfm at depth for depths of 0-850 ft. This system is the subject of U. S. Patents No. 3062017 and 3369558.

Under the closely-controlled metering conditions maintained in the above system, with pressure drop in the transonic range but not quite great enough for critical flow, two methods of flow metering appear especially appropriate. One is the use of a variable-area orifice, with the flow rate proportional to effective orifice area. For this purpose one or more multi-turn micrometer valves, using 10 or more turns for flow modulation, would provide reproducible flow-area settings; the actual flow for any reading of the micrometer-valve handles could be calibrated and used in curves or tables of settings for desired flow rates and mixtures. A second method would be use of several parallel small orifices which could be valved in and out of the flow circuit with solenoid valves in various combinations, to provide various flow rates. Both of these methods are well suited to precise flow measurement over wide flow ranges.

The flexibility of such a system depends a great deal upon the flexibility of flow measurement capability provided by the combination of metering sensors and the selection system. It would be possible to provide a range of flow rates at any desired pressure and mixture ratio with a sophisticated selection procedure that would provide for variations in metering pressure and area. Alternatively, operation could be limited to a few pre-scheduled mixtures and flow rates with a simpler selection procedure. The entire system could be cycled on and off by a pressure switch on an accumulator (9), actuating solenoid valves on inlet gas lines.

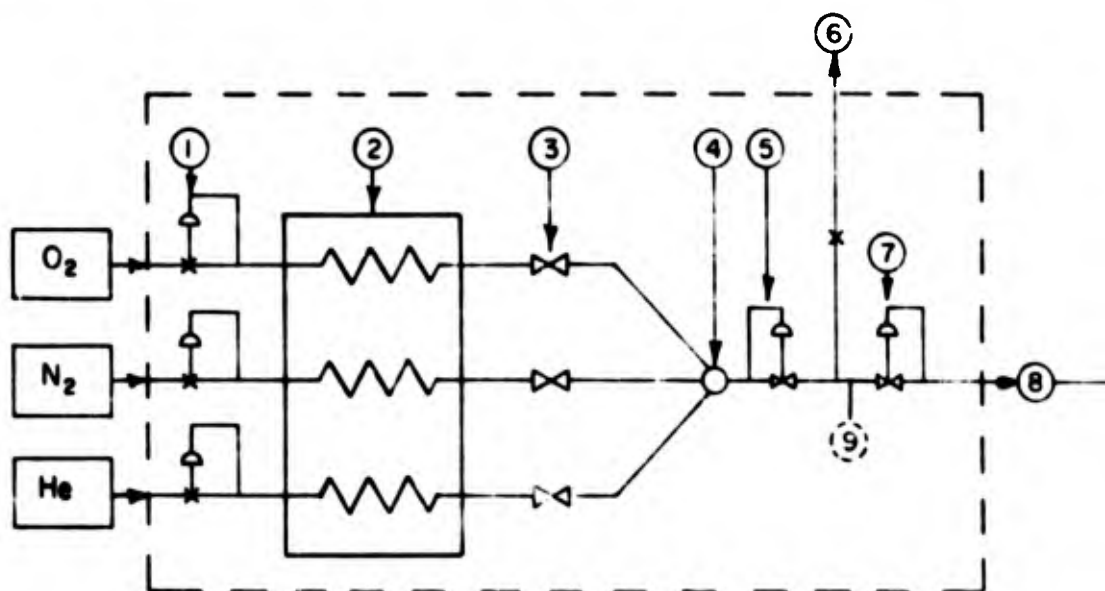
*U. S. Navy Experimental Diving Unit.

The accuracy of mixing claimed for the Airco Mixmaker is as follows:

<u>Percentage oxygen in mixture</u>	<u>Precision of oxygen content</u>
0 - 10	±0.15 percent oxygen
10 - 25	±0.25 percent oxygen
25 - 100	±0.50 percent oxygen

For example, the oxygen content of a mixture prepared with the 5-percent-oxygen setting could vary between 4.85 and 5.15 percent.

Continuous-flow breathing-gas mixture systems can be used to supply divers continuously from the mixing system, or the mixture can be fed to a high-pressure compressor for filling of SCUBA tanks or compressed-gas storage tanks. The oxygen content of the mixture to be prepared must be determined on the basis of the type of breathing apparatus to be used and the range of diving depths.



- | | |
|------------------------------|--|
| 1. Pressure Regulation | 6. Mixed Gas Sample to Analysis Console |
| 2. Temperature Equilibration | 7. Pressure Regulation for Direct Diver Supply or Storage Compressor |
| 3. Metering Valves | 8. Gas Mixing Console Interface |
| 4. Mixing Chamber | 9. (Possible accumulator, if used) |
| 5. Back Pressure Regulator | |

FIGURE 4A. SCHEMATIC ARRANGEMENT OF AIRCO MIXMAKER[®] MIXING SYSTEM

Courtesy Airco Central Research Laboratories

USEFUL CAPACITY OF STORAGE TANKS (REAL-GAS METHOD)

The useful capacity of a tank used for storage of high-pressure breathing gas is the quantity of gas that can be supplied from it for a specific mission. It depends upon the tank volume, the initial gas pressure, and the final gas pressure below which gas cannot be removed. The capacity will be different for different missions because the final pressure may differ with diving depth and type of equipment used.

The initial weight of gas in the tank is the product of initial gas density and tank volume. The residual weight is the product of final gas density and tank volume. The weight removed, thus, can be calculated as the product of the tank volume and the change in gas density between initial pressure and final pressure, or as the difference between initial and final weights of gas.

As an example, the capacity of a tank is determined as follows:

<u>Conditions</u>	<u>Example</u>
Tank Volume	30 ft ³
Tank Temperature	50 F
Initial Pressure	3,000 psia (2985 psi)
Residual Pressure	400 psia (385 psi)
Gas Mixture in Tank	10% oxygen, 90% helium
(1) Initial Density in Tank (from Table T-63 for 3000 psia, 50 F)	3.374 lb/ft ³
(2) Residual Density in Tank (from Table T-63 for 400 psia, 50 F)(see Column 2 at 870 ft)	<u>0.492 lb/ft³</u>
(3) Density Change (Initial - Residual)	2.882 lb/ft ³
(4) Weight of Gas Used (Tank Vol) x (Density Change)	30 ft ³ x 2.882 lb/ft ³ = <u>86.2 lb</u>
(5) Volume of Gas Available to Diver	
Water Temperature	50 F
Working Depth of Diver	300 ft
Specific Volume of Gas at 300 Ft, 50 F (from Table T-62)	5.46 ft ³ /lb
Total Volume = (Weight) x (Sp. Vol)	86.2 lb x 5.46 = 466 ft ³

VENTILATION OF HYPERBARIC PERSONNEL CHAMBERS

In modern diving operations it is usual to utilize a variety of pressurized working spaces - decompression chambers, personnel transfer chambers, underwater habitats, experimental diving chambers, and surface pressure chambers are typical examples. All of these share the common problem of ventilation to limit the partial pressure of carbon dioxide to non-toxic partial pressures. The general requirement for ventilation is exactly the same as that for helmet ventilation: the rate of venting of carbon dioxide must be equal to the rate of production. During steady-state operation the helmet-ventilation equation on Page 21 applies:

$$V = \frac{P_{atm} \times O_{slm} \times R \times F}{26.2 (C_2 - C_1 \times P_{atm})}$$

Terms are defined on Page 21.

In using this equation it is necessary to apply appropriate values for the rate of oxygen consumption and the mixing factor. The rate of oxygen consumption can be estimated from the number of people in the chamber and the rate of effort expected of them. A man at rest undergoing treatment might require 0.5 slm oxygen, while an attendant working over him might require 1.0 slm. Light work may be expected in some situations, and digestion of food will generate carbon dioxide at a rate equivalent to light work over a period of an hour or more.

The applicable mixing factor is highly problematical, and every effort should be made in the chamber design and development to optimize it at near 1.0. This can be done by providing for good internal gas recirculation with a fan or ejector, and by selection of best locations for gas admission and venting. Mixing factor can be evaluated by introducing a gaseous tracer at various points in the chamber and measuring concentrations of the tracer at points within the chamber and in the vented gas. Such procedures would be well justified in chamber development as a means of minimizing required ventilation rates and avoiding excessive local carbon dioxide concentrations.

The level of carbon dioxide partial pressure that produces symptoms of toxicity depends upon the period of exposure. Partial pressure of 0.02 atm produces only mild symptoms in 30 minutes of exposure, and partial pressure of 0.01 atm produces no observable toxic symptoms in exposures of several hours. (5) However, for exposures of 12 to 38 hours, as required for longer treatment schedules in recompression chambers, levels below 0.005 atm appear necessary to avoid toxic effects, and levels of about 0.003 atm are recommended for very long exposures during saturation diving operations. (3) The Diving Manual presents ventilation schedules for recompression chambers designed to maintain carbon dioxide partial pressures at 0.015 atm or less regardless of exposure time. These schedules provide 2 cfm of air at chamber pressure ($2 \times P_{atm}$) for each occupant at rest and 4 cfm at pressure for each active attendant.

In the more elaborate hyperbaric facilities ventilation is provided by circulation of chamber gas through a conditioning system that absorbs carbon dioxide, removes water vapor, and controls gas temperature. The ventilating flow rate is controlled by a circulating fan that circulates the same volume of gas at any pressure, thus satisfying

the ventilation requirement under all conditions. Simple recompression chambers usually rely upon venting of gas from the chamber and replacement with fresh gas. The partial pressure in the chamber is determined by the rate of venting, not the rate of supply, of gas. Because of the slow rate of change of carbon dioxide partial pressure in a large chamber of 250 or 500 cubic feet of volume, it is permissible to ventilate such a chamber at intervals, exchanging a large quantity of gas in a short time instead of providing continuous flow at a low rate. This is often done to avoid the continuous noise of unsilenced gas inlets and outlets.

The oxygen content of breathing gas provided for hyperbaric chambers should be matched to circumstances; in saturation-diving habitats oxygen partial pressure should be below 0.37 atm to avoid lung irritation, but in recompression chambers, where air is used at pressures to six atm, oxygen partial pressure will be in the range of 1.2 atm. When pure oxygen is used for respiration within recompression chambers the Diving Manual recommends chamber ventilation rates of 4 cfm at pressure for each occupant breathing oxygen, which is designed to keep oxygen percentage in the chamber below 30 percent.

Helium-oxygen mixtures are recommended in the Diving Manual for treatment of serious decompression sickness. The mixture would contain about 20 percent oxygen, and the ventilation rate would be the same as for air.

CARBON DIOXIDE ABSORPTION IN DIVING OPERATIONS

Rebreathing of expired air or breathing gas is common in diving operations, including decompression chambers, underwater habitats, personnel transfer chambers, closed-circuit and semiclosed-circuit breathing apparatus, and heavy deep-sea diving rigs. In each application it is necessary to remove virtually all of the exhaled carbon dioxide before rebreathing the gas in order to avoid toxic effects. This is done in a scrubber consisting of a canister filled with a chemical absorbent through which the exhaled gas is passed. The design of the scrubber must be such that all of the carbon dioxide that enters is removed and the gas leaving it is completely free of carbon dioxide. The life of the chemical absorbent is effectively ended when carbon dioxide begins to pass through, even though it may have considerable absorptive capacity remaining. In practice, it has been found that about half of the theoretical absorptive capacity of common chemical absorbents can be utilized in diving applications.

The quantity of carbon dioxide generated by the diver, the number of divers served by the scrubber, and the period over which the scrubber must operate before replenishment determine the weight of absorbent required. The design of the canister must be such that all of the absorbent is well exposed to the gas passing through, in order to obtain efficient use of the material.

An extremely important factor in life of absorbent is the temperature at which it operates. Experiments have shown that a canister sized to perform adequately for a 4-hour period when at 70 F will reach its effective life after only 30 minutes when at 40 F. Thus, for efficient use of absorbent, it is necessary to heat the canister to maintain it at a temperature near 70 F. (8)

The quantity of carbon dioxide generated by each diver depends upon his rate of effort. The volume of carbon dioxide generated is slightly less than the volume of oxygen consumed and metabolized, the ratio being between 0.7 and 0.9 depending upon carbon-hydrogen ratio of food being metabolized. For design purposes, a value of 0.9 is recommended. The rate of effort at which a diver can work can be such as to use as much as 4 slm of oxygen on shore, but it is difficult to utilize more than 3 slm under water at shallow depths, and 2 slm at very great depths, because of the increase in breathing resistance with depth. Thus, if oxygen consumption of 3 slm is assumed, with a conversion factor of 0.9, the rate of production of carbon dioxide would be 2.7 slm. This value converts to 5.72 ft³/hr or 0.71 lb/hr.

Table 3 summarizes information on three carbon dioxide absorbents. Data are based on requirements of one diver at 3 slm oxygen usage.

TABLE 3. CHARACTERISTICS OF CARBON DIOXIDE ABSORBENTS

Absorbent	Absorbent Density, lb/ft ³	Theoretical Absorption, lb CO ₂ /lb	Useful Absorption, lb CO ₂ /lb	Pounds of Absorbent per Diver Hour, (0.71 lb CO ₂)	Canister Volume per Diver Hour, ft ³	Cost, \$/hr
Lithium ⁽¹⁰⁾ hydroxide	28	0.92	0.46	1.55	0.0552	\$6.20
Sodasorb ^(8,9)	55.4	0.49	0.245	2.90	0.0533	\$0.75
Baralyme ^(8,9)	65.4	0.39	0.195	3.65	0.0558	\$1.75

Canisters of carbon-dioxide absorbent for self-contained breathing apparatus should be sized for the planned mission duration and should provide absorptive capacity that matches breathing-gas storage capacity. Apparatus for use in very cold water should provide for heating and insulation of the canister in order to realize the potential life of the absorbent, which is greatly reduced at low temperatures(8).

Life-support systems for hyperbaric personnel chambers should incorporate carbon dioxide absorbers that can be conveniently replaced without interference with normal activities within the chamber. It is common to use dual absorbers so that one can be replaced while the other continues to operate, but it is also possible, in a large chamber, to interrupt operation of the life-support system for a short time for absorbent replacement.

The usual absorbent canister permits utilization of about 50 percent of the theoretical absorbing capacity of the absorbent before "breakthrough", when carbon dioxide begins to pass through the absorbent. The breakthrough region is usually a small local area which is exposed to a higher-than-average flow rate because of non-uniform flow distribution through the absorbent. It appears possible to improve the utilization of absorbent by careful design of canisters for ideal flow distribution.

PSYCHROMETRIC CALCULATIONS FOR DIVING APPLICATIONS

Description of Psychrometric Charts

Standard psychrometric charts are commonly used in calculating the effects of heating and cooling moist air at atmospheric pressure. Such charts relate enthalpy changes to wet- and dry-bulb temperatures, relative humidity, dew point, and specific humidity, and permit graphical solutions of usual problems. These standard charts are set up for a single gas mixture (air) and for a single pressure (14.7 psia), and are not suitable for use with a range of gas compositions and pressures. Accordingly, a special humidity chart suitable for use with diving-gas mixtures to depths of 1300 feet has been developed for use in this manual.

In order to construct a chart suitable for gases of differing molecular weight and density, the mole has been used as the unit of gas quantity. A mole, as used in this discussion, is a quantity having a weight, in pounds, equal to the molecular weight of the dry gas mixture. A useful property of this unit is that the volume of a mole of gas is the same for any gas of any molecular weight, at any specified temperature and pressure. Accordingly, the molar volume and the partial pressure of water vapor provide a means of determining the quantity of water per mole of gas, on which all other psychrometric relations are based.

Figures 5, 6, and 7 are psychrometric charts developed for use at elevated pressures with breathing-gas mixtures. Chart 1 is for use with gas mixtures at pressures from 14.7 psia to 200 psia. Chart 2 covers the pressure range from 100 psia to 600 psia. Chart 3 provides temperature corrections for moisture enthalpy, to be applied to enthalpy data from Charts 1 and 2. Charts 1 and 2 are similar except for the vertical scale.

The slopes of constant-wet-bulb lines for different gas mixtures are shown in the upper left area of Charts 1 and 2. The molar specific heats of air, nitrogen, and oxygen are approximately equal at 6.95 Btu/lb-mole, as represented by the highest line, and the molar specific heat of helium is lower, at 4.96 Btu/lb-mole, represented by the lowest line. Lines for gas mixtures containing different percentages of helium are plotted between these limits. To determine a constant-wet-bulb-temperature line, a line is drawn parallel to the appropriate line in the upper left area of the chart, using a parallel rule or drafting triangles. This line starts at the intersection of the wet-bulb temperature and the proper pressure curve and extends to the right until it reaches the desired value of dry-bulb temperature. The mole fraction of water vapor in the mixture at the dry-bulb temperature can be read from the right scale. The relative humidity can then be calculated as the ratio of this value to the value corresponding to the intersection of the dry-bulb temperature with the proper pressure curve. Alternatively, if dry-bulb temperature and relative humidity are known, wet-bulb temperature can be found by the following procedure: (1) read moisture mole fraction for saturated gas at dry-bulb temperature and proper pressure curve; (2) multiply this value by relative humidity, and plot the result on the dry-bulb temperature line; (3) draw a constant-wet-bulb-temperature line from this point to the left, to intersect with the proper pressure line. The wet-bulb temperature can be read below this intersection on the temperature scale.

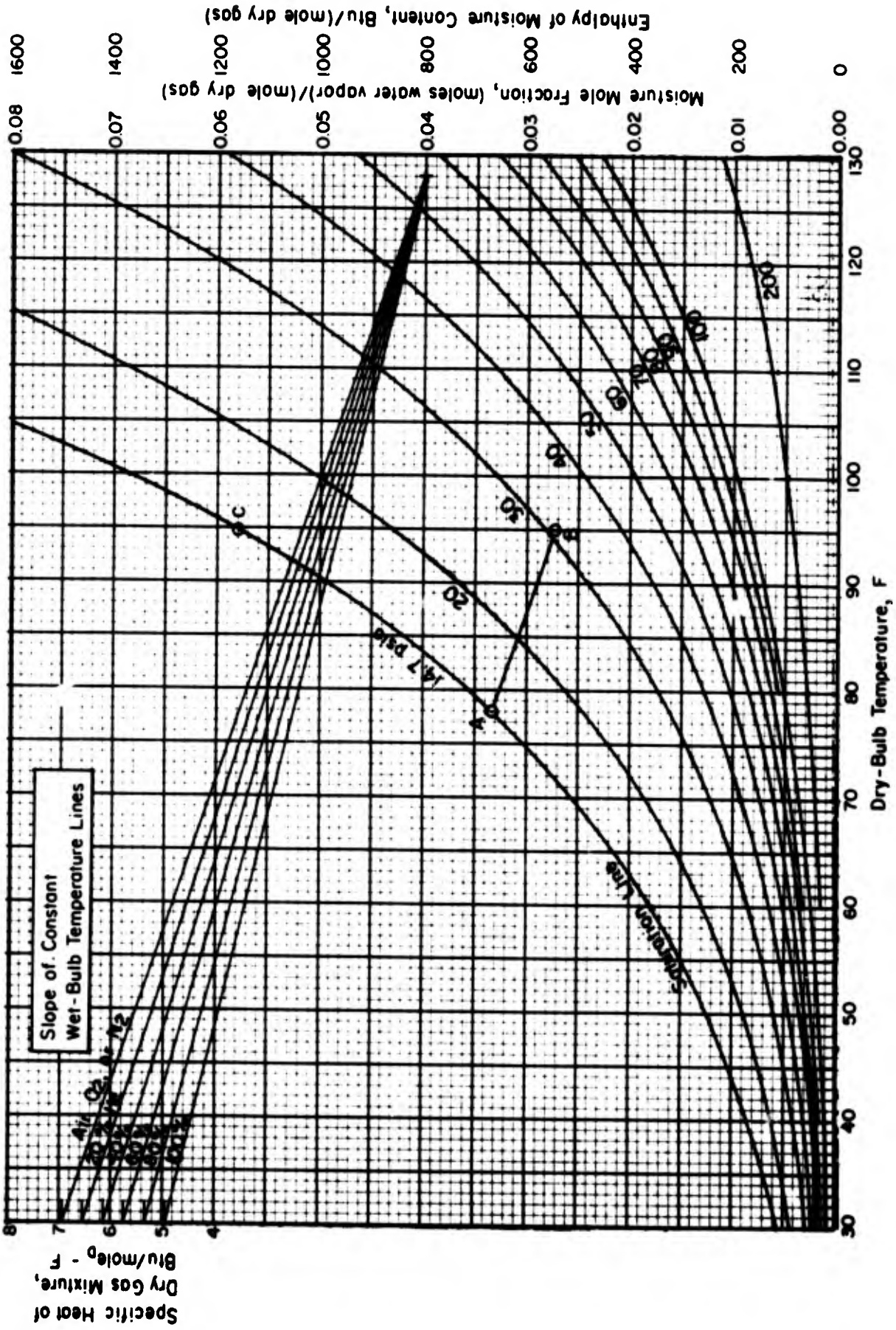


FIGURE 5. PSYCHROMETRIC CHART 1 - DATA FOR GAS MIXTURES AT PRESSURES FROM 14.7 PSIA TO 200 PSIA

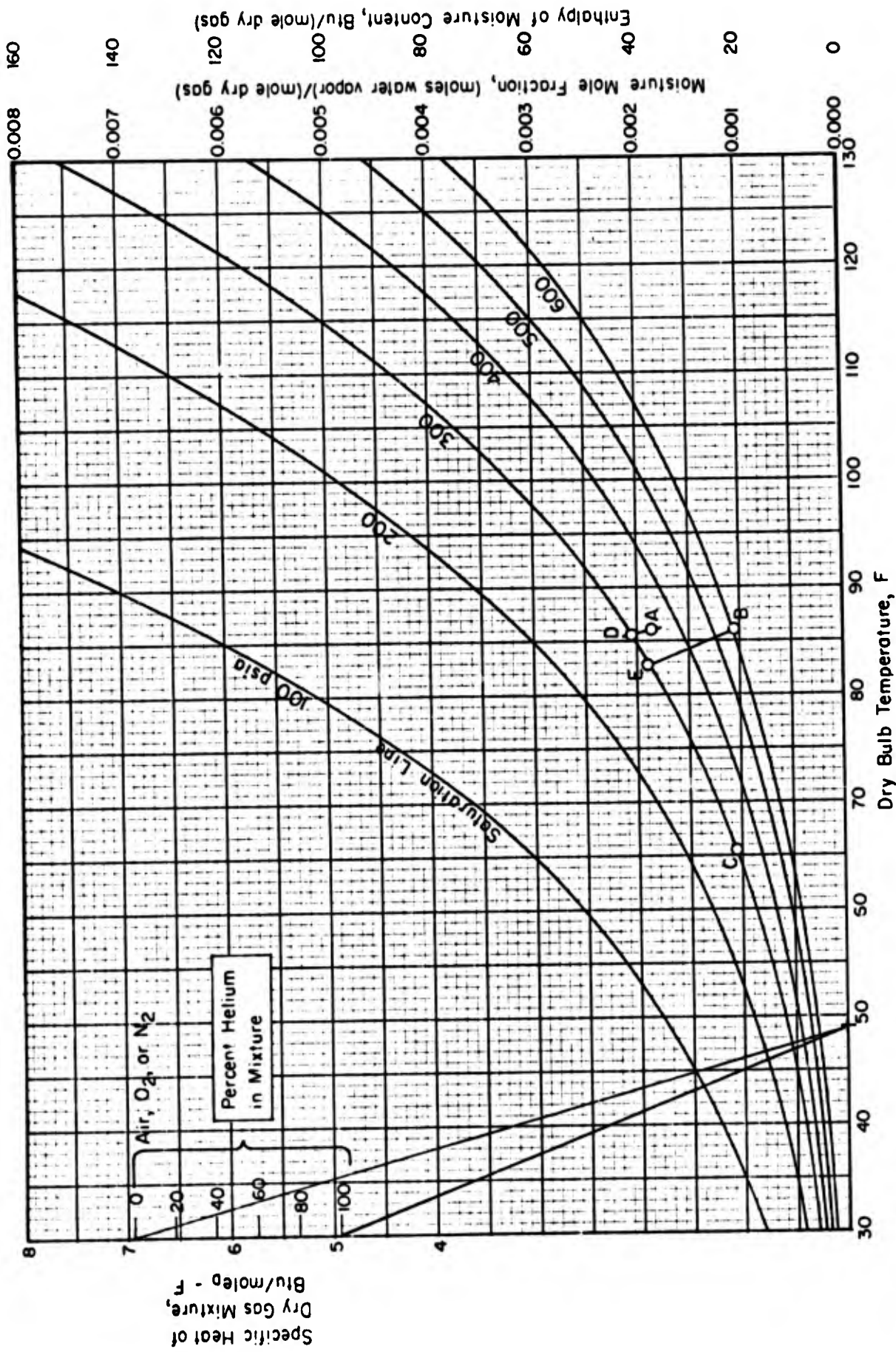


FIGURE 6. PSYCHROMETRIC CHART 2 - DATA FOR GAS MIXTURES AT PRESSURES FROM 100 PSIA TO 600 PSIA

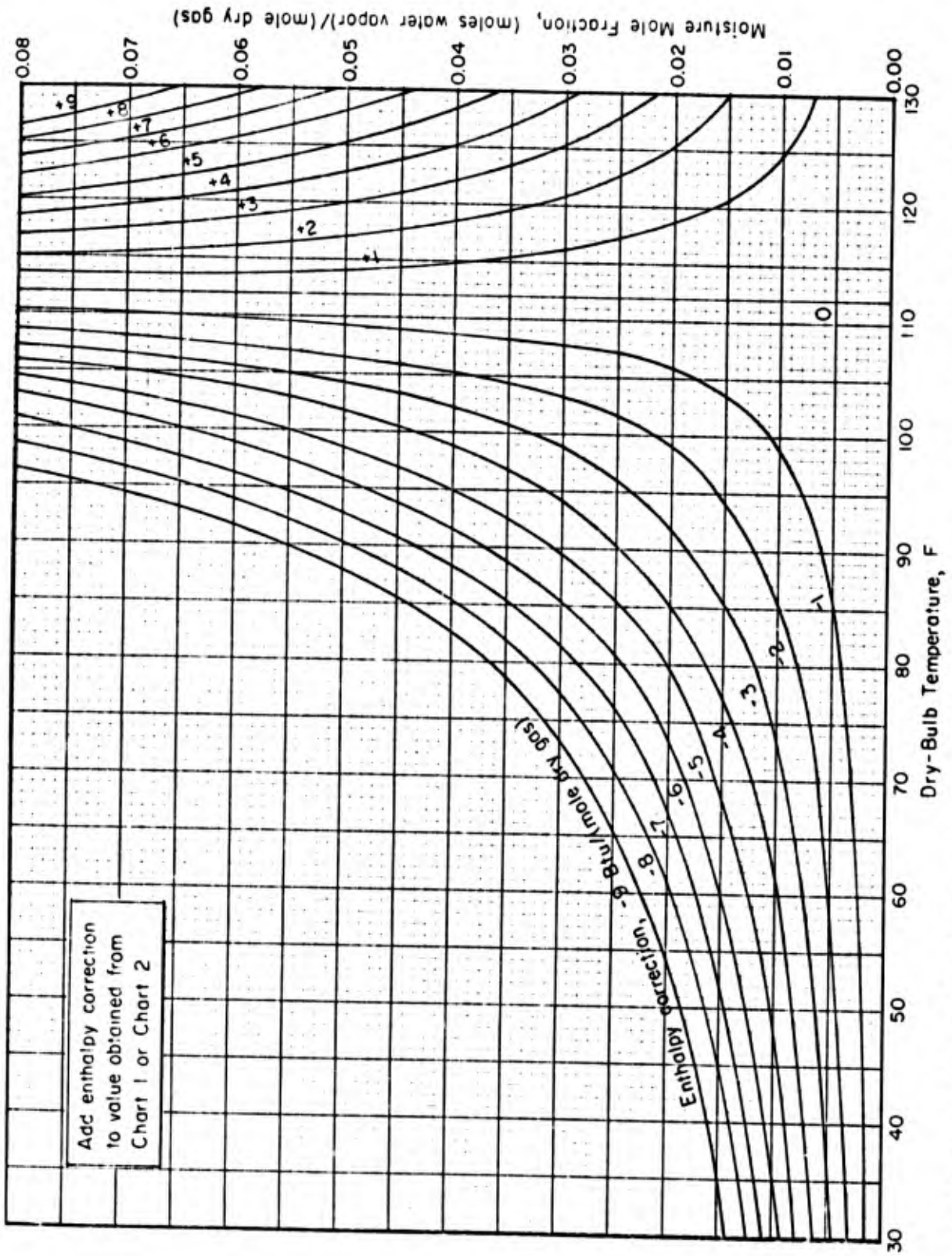


FIGURE 7. PSYCHROMETRIC CHART 3 - CORRECTION FOR ENTHALPY OF MOISTURE CONTENT

Enthalpy is given for the moisture content alone, rather than for the moist gas mixture, and is expressed in Btu/mole of dry gas. Most of the enthalpy of the water vapor is in the latent heat of vaporization. Consequently, the water-vapor enthalpy is almost directly proportional to the amount of water vapor present. For this reason, both enthalpy and moisture concentration appear on the vertical scale of Charts 1 and 2. However, there is a small effect of mixture temperatures which is presented in Chart 3 as a temperature correction.

Detailed methods for various calculations are best shown by examples, which are included below.

Data sources and computational methods involved in preparing the charts and accuracy are discussed in Appendix B.

Definitions of Psychrometric Terms

The nomenclature and definitions relating to the psychrometric charts presented in this manual are generally consistent with those presented in Chapter 6 of the ASHRAE Handbook of Fundamentals (published by the American Society of Heating, Refrigerating, and Air-Conditioning Engineers, Inc., New York, 1967). The following definitions are paraphrased from this publication:

- W - Humidity ratio is the ratio of the mass of water vapor contained in a given sample of moist gas to the mass of dry gas with which the water vapor is associated, pounds of water vapor per pound of dry gas.
- ϕ - Relative humidity is the ratio of the mole fraction of water vapor in moist gas to the mole fraction of water vapor in saturated moist gas at the same temperature and pressure. It is normally expressed as a percentage.
- x - Mole fraction pertaining to any given constituent in a mixture of gaseous substances is used herein as the number of moles of that constituent present in the mixture divided by the total number of moles of dry gas contained in the mixture. It is numerically equal to the volume fraction. x_w represents the mole fraction for water vapor, and x_{ws} represents that for water vapor at saturation.
- t_{db} - Dry-bulb temperature is the Fahrenheit temperature of moist gas at rest with respect to the temperature-measuring element.
- t_{wb} - Thermodynamic wet-bulb temperature is the temperature at which water (liquid or solid), by evaporating into moist gas at given dry-bulb temperature, t_{db} , and humidity ratio, W, can bring the gas to saturation adiabatically at the same temperature t_{wb} while the pressure p is maintained constant.

Additional nomenclature employed in following examples of the use of the diving-gas psychrometric charts are:

c_{pm} - Specific heat at constant pressure, expressed as Btu/mole of dry gas

h - Specific enthalpy, Btu/lb

h_m - Molar enthalpy, Btu/mole

M - Pound molecular weight, lb/mole

Q - Volumetric flow rate, ft³/min

q - Heat flow rate, Btu/min

w - Mass flow rate, lb/min

ρ - Density, lb/ft³.

General subscripts used are the following:

d - refers to dry gas mixture

w - refers to water vapor

m - refers to mole as unit of weight.

Example Calculations

Three example calculations which illustrate the use of the diving-gas psychrometric charts are given below.

Example 1

Find the relative humidity of moist air at 1 atm pressure (14.7 psia), 95 F dry-bulb temperature, and 78 F wet-bulb temperature.

Procedure: Using Chart 1 locate Point A at the intersection of the 78 F dry-bulb-temperature line and the 14.7-psia saturation line. (At saturation, wet-bulb and dry-bulb temperatures are the same.) Draw Line AB from Point A parallel to the constant-wet-bulb line for air, using a parallel rule or drafting triangle. Point B is the intersection of the 78 F wet-bulb line (just drawn) and the 95 F dry-bulb line.

Locate Point C by going up the 95 F dry-bulb line to the 14.7 psia saturation line.

Reading the vertical scale on the right to determine the moisture content (expressed as mole fraction) for Points B and C gives the following:

.0272 mole_w/mole_d for Point B

.0581 mole_w/mole_d for Point C.

The relative humidity is the ratio of the moisture content for Point B to the moisture content of saturated gas at the dry-bulb temperature (Point C). Therefore,

$$\begin{aligned}\text{Relative humidity} = \phi &= .0272 / .0581 \\ &= 0.468, \text{ or} \\ &46.8 \text{ percent}\end{aligned}$$

Example 2

Find the humidity ratio and the enthalpy of the gas mixture used in Example 1 (air at one atm pressure, 95 F dry-bulb temperature, and 78 F wet-bulb).

Using the vertical scale on the right for Point B, the enthalpy of the moisture content is read as 544 Btu/mole_d.

Using Chart 3, the enthalpy correction for $x_w = .0272$ and $T_{db} = 95$ F is about 3 Btu/mole_d.

The molar enthalpy of the moisture content is:

$$h_{wm} = 544 - 3 = 541 \text{ Btu/mole}_d$$

The enthalpy of the moist air mixture is given by

$$h = [(c_{pm})_d t_{db} + h_{wm}] / M_d$$

For air,

$$(c_{pm})_d = 6.95 \text{ Btu/mole}_d$$

$$M_d = 28.96 \text{ lb/mole}$$

Then

$$h = [(6.95)(95) + 541] / 28.96$$

$$h = 41.5 \text{ Btu/lb}_d$$

The humidity ratio, W , is given by

$$W = x_w (M_w / M_d)$$

$$W = .0272 (18.01 / 28.96)$$

$$W = .0169 \text{ lb water vapor per lb dry air}$$

From standard psychrometric charts for air at 1 atm, the following values are obtained directly:

$$h = 41.58 \text{ Btu/lb}_d$$

$$W = .0168 \text{ lb}_w/\text{lb}_d \text{ (approximately)}$$

$$\phi = 47 \text{ percent (approximately)}$$

The values obtained by the use of Charts 1 and 3 in this example are essentially in agreement with those obtained using standard psychrometric charts for air.

Example 3

Find the energy requirement, moisture removal rate, and coil temperature to dehumidify 100 cfm of a 1 percent O_2 - 99 percent He gas mixture at 300 psia from 86 F dry-bulb temperature and 90 percent relative humidity to 50 percent relative humidity. Also find the wet-bulb temperatures for this gas mixture at 86 F dry-bulb temperature and at 90 and 50 percent relative humidities.

Procedure:

Moisture Removal Rate. Using Chart 2, the saturation line for 300 psia shows that the moisture mole fraction at 86 F and 100 percent relative humidity is .00209 mole_w/mole_d. Since relative humidity, ϕ , is:

$$\phi = x_w/x_{ws}$$

the moisture mole fraction at $\phi = 90$ percent is

$$x_w = 0.90 (.00209)$$

$$x_w = .01881 \text{ mole}_w/\text{mole}_d$$

and at $\phi = 50$ percent is

$$x_w = 0.50 (.00209)$$

$$x_w = .001045 \text{ mole}_w/\text{mole}_d$$

The 90 and 50 percent relative humidity points are shown on Chart 2 as Points A and B. From the density tables for the 1 percent O_2 - 99 percent He mixture, the density figures given for 640 ft depth (299.14 psia) are

$$\rho_d = 0.21908 \text{ lb/ft}^3 \text{ at } 80 \text{ F}$$

$$\rho_d = 0.21514 \text{ lb/ft}^3 \text{ at } 90 \text{ F}$$

and the molecular weight of the dry gas mixture is given as 4.283. Interpolation to find the density at 86 F gives

$$\rho_{86 \text{ F}} = 0.2167 \text{ lb/ft}^3$$

(For engineering purposes, there is no need to correct for the difference between 299.14 and 300 psia).

The humidity ratio is given by

$$W = x_w (M_w/M_d) \quad ;$$

therefore, at 90 percent relative humidity

$$W = .01881 (18.01/4.283)$$

$$W = .07910 \text{ lb}_w/\text{lb}_d \quad .$$

and at 50 percent relative humidity

$$W = .01045 (18.01/4.283)$$

$$W = .04394 \text{ lb}_w/\text{lb}_d \quad .$$

The dry-gas mass flow rate at 100 cfm is

$$w_d = \rho_d Q_d$$

$$w_d = 0.2167 (100)$$

$$w_d = 21.67 \text{ lb}_d/\text{min} \quad .$$

The moisture removal rate is

$$w_w = w_d (W_{90} - W_{50})$$

$$w_w = 21.67 (0.07910 - 0.04394)$$

$$w_w = 0.7619 \text{ lb}_w/\text{min} \quad .$$

Coil Temperature. The coil temperature is found at Point C on Chart 2, a point on the 300-psia saturation line at which the moisture mole fraction is the same as that of 86 F dry-bulb gas at 50 percent relative humidity (Point B). This indicates that it will be necessary to cool the gas mixture to 65.5 F to remove the desired amount of moisture.

Energy Requirement. The moisture enthalpy values corresponding to Points A and C are:

$$h_{wm} = h_{wm} \text{ (from Chart 2)} - \Delta h_{wm} \text{ (from Chart 3)}$$

$$h_{wm} = 37.6 - 0.3$$

$$= 37.3 \text{ Btu/mole}_d \text{ at } t_{db} = 86 \text{ F, } \phi = 90 \text{ percent} \quad .$$

$$h_{wm} = 20.9 - 0.3$$

$$= 20.6 \text{ Btu/mole}_d \text{ at } t_{db} = 65.5 \text{ F, } \phi = 100 \text{ percent} \quad .$$

The corresponding moist-gas-mixture enthalpies are computed by the following expression:

$$h = [(c_{pm})_d t_{db} + h_{wm}] / M_d$$

The value of the specific heat of the dry gas is computed as follows:

$$\begin{aligned} (c_p)_d &= 6.95 x_{O_2} + 4.965 x_{He} \\ &= 6.95(0.01) + 4.965(0.99) \\ &= 4.985 \text{ Btu/mole}_d\text{-F} \end{aligned}$$

Using this value in the expression for enthalpy, the following values are obtained:

$$h_A = [(4.985)(86) + 37.3] / 4.283$$

$$h_A = 108.80 \quad \text{Btu/lb}_d$$

$$h_C = [(4.985)(65.5) + 20.6] / 4.283$$

$$h_C = 81.05 \quad \text{Btu/lb}_d$$

$$q = w_d (h_A - h_C)$$

$$q = 21.67 (108.80 - 81.05)$$

$$q = 601.3 \quad \text{Btu/min}$$

This is equivalent to $601.3/200 = 3$ tons of refrigeration, approximately. (One ton of refrigeration is defined as 200 Btu/min.)

In controlling the climate of a diving habitat, it may be necessary to reheat the gas leaving the cooling coil. If, for example, it is desirable to have the gas from the environmental control system returning at 75 F, the reheat energy required can be computed as follows:

$$\begin{aligned} \Delta h_{65.5-75} &= (c_p)_d \Delta t_{db} / M_d \\ &= 4.985 (75 - 65.5) / 4.283 \\ &= 11.06 \text{ Btu/lb}_d \end{aligned}$$

$$q = w_d \Delta h$$

$$= 21.67 (11.06)$$

$$= 239.6 \text{ Btu/min}$$

Wet-Bulb Temperatures. On Chart 2, the slope of the constant wet-bulb temperature line should be that corresponding to a specific heat of 4.985 Btu/mole_d. Drawing lines of this slope from Points A and B to the 300-psia saturation line gives Points D and E. At saturation, wet-bulb and dry-bulb temperatures are the same; therefore, the following results are obtained:

$$t_{wb} = 85.4 \text{ F when } t_{db} = 86 \text{ F and } \phi = 90 \text{ percent}$$

$$t_{wb} = 82.8 \text{ F when } t_{db} = 86 \text{ F and } \phi = 50 \text{ percent}$$

MISCELLANEOUS CALCULATIONSSeawater Pressure at Depth

The seawater pressure at depth can be calculated by several methods. Three of these follow.

Method 1

$$\text{Seawater pressure, atm} = \frac{\text{depth, ft}}{33} + 1$$

This calculation is correct for seawater density of 64.12 lb/ft³, and is used in the U. S. Navy Diving Manual.

Method 2

Select atm pressure for depth of interest from Column 2 in tables of properties.

The tables are based on seawater density of 64.0 lb/ft³, so the resulting pressure will be 0.2 percent less than for Method 1. The equation used in calculating this table is:

$$\text{atm} = \frac{\text{depth, ft} \times 64.0 \text{ lb/ft}^3}{14.696 \times 144} + 1$$

This reduces to:

$$\text{atm} = \frac{\text{depth, ft}}{33.1} + 1$$

Method 3

Measure the temperature gradient and salinity gradient at the diving site and calculate a seawater density gradient and the true pressure at the depth of interest. After mean density has been calculated, atm pressure is:

$$\text{atm} = \frac{(\text{depth, ft}) \times (\text{mean density, lb/ft}^3)}{14.696 \times 144} + \frac{(\text{barometric pressure, in Hg})}{2.041}$$

The barometric pressure is approximately 14.7 psi, and can vary by about ±0.5 psi with the weather. It will influence bottom pressure accordingly.

For normal diving calculations, the same results can be obtained within satisfactory tolerances by any of these methods.

Oxygen Percentage at Specified Partial Pressure and Depth

The percentage of oxygen in a mixture corresponding to a selected oxygen partial pressure and depth can be calculated as follows:

$$\text{Percent oxygen} = \frac{(\text{partial pressure oxygen, atm}) (100)}{(\text{seawater pressure at depth, atm})}$$

Carbon Dioxide Percentage at Specified Partial Pressure

The percentage of carbon dioxide in a mixture corresponding to a selected carbon dioxide partial pressure and depth can be calculated as follows:

$$\text{Percent carbon dioxide} = \frac{(\text{partial pressure carbon dioxide, atm}) (100)}{(\text{seawater pressure at depth, atm})}$$

Conversion of Liters to Cubic Feet

Several conversions that are frequently used can be easily confused. These involve the following conversion factors:

- (1) 1 ft³ = 28.317 liters
- (2) 1 std ft³ is mass in 1 ft³ at 60 F, 14.696 psia
- (3) 1 std liter is mass in 1 liter at 32 F, 14.696 psia
- (4) 1 std ft³ = 26.3 std liters (accounting for temperature correction).
- (5) 1 std ft³ air for compressor rating is at 14.7 psia, 36% RH, 70 F,
 $\rho = 0.075 \text{ lb/ft}^3$

Conversion Factors for Viscosity

In the tables of Part 2, viscosity is expressed as lb/ft sec. The following conversion factors can be used in converting to other units:

<u>To get:</u>	<u>Multiply table entry by:</u>
Poise	14.88
Kg/hr meter	5.357 x 10 ⁴
Lbs/ft hr	3600
Slugs/ft hr	111.89

REFERENCES

- (1) Stanley Miles, Underwater Medicine, Lippincott, Philadelphia, 1962.
- (2) U. S. Navy Diving Manual, NAVSHIPS 250-538 (1963 and 1965).
- (3) C. J. Lambertsen, "Limitations and Breakthroughs in Manned Undersea Activity", Supplementary Transactions of 2nd Annual MTS Conference, June 27-29, 1966, Marine Technology Society, 1030 15th Street, N. W., Washington, D. C. 20005.
- (4) B. Lewis and G. von Elbe, Combustion, Flames, and Explosion of Gases, Academic Press, Inc., 1951.
- (5) O. R. Hansen, "Analytical Evaluation of Diver's Breathing Air Systems on ARS Type Ships", USN Supervisor Diving Research Report 2-69, September, 1969.
- (6) P. S. Riegel and J. V. Harter, "Design of Breathing Apparatus for Diving to Great Depths", ASME Paper No. 69-De-22 available from American Society of Engineers, 345 East 47th Street, New York, New York 10017.
- (7) P. S. Riegel, The 3-color chart of Figure 4 was provided by P. S. Riegel and is to be described in detail in a future publication.
- (8) B. R. Lower, J. S. Glasgow, and D. W. Frink, "Mark VII and IX Carbon Dioxide Canister Heat Exchangers", report to U. S. Navy Experimental Diving Unit, Contract No. N00014-C-0199, April 25, 1969.
- (9) From manufacturers' brochures, NCG Division of Chemtron Corp., and Dewey and Almy Chemical Company.
- (10) R. O. Bach, "Application of Lithium Chemicals for Air Regeneration of Manned Spacecraft", AMRL-WPAFB Report, June, 1965, AD 61947.

U. S. NAVY DIVING-GAS MANUAL

PART 2.

PROPERTIES OF HELIUM-OXYGEN MIXTURES AND OF PURE GASES

October 1, 1969

PART 2. PROPERTIES OF OXYGEN-HELIUM MIXTURES AND OF PURE GASES

CONTENT AND FORMAT OF TABLES

The tables of gas properties included in this manual were prepared for convenient use in design and operation of diving systems. The pressure range covered is 0 to 1500 feet of seawater, and 500 to 5000 psi, for use in high-pressure-gas-storage calculations. All pressures are also expressed in atm as a convenience for calculations. The temperature range covered is 30 F to 130 F, which covers the usual range of air and water temperatures for diving operations.

In order to present data as a function of depth, it was necessary to select a suitable value for the density of seawater. As discussed in Appendix A, density of seawater varies slightly with temperature and salinity, so that no single value can be correct for all conditions. A value of 64.0 lb/ft³ was selected for use in preparing these tables. If another value is known to be applicable, the known value can be used to calculate the true pressure at depth, and tabulated values for this pressure can then be used.

Properties of five pure gases and of thirteen mixtures of helium and oxygen are tabulated. These gases and mixtures appear most useful in diving studies. The mixtures included in the tables are

Air
Oxygen
Helium
Nitrogen
Carbon dioxide
Oxygen-helium mixtures:

<u>O₂, percent</u>	<u>He, percent</u>
60	40
40	60
32	68
20	80
15	85
10	90
6	94
3	97
1.5	98.5
1.0	99.0
0.7	99.3
0.5	99.5
0.3	99.7

For each of the 18 gases and mixtures listed above, the following properties are tabulated:

Property	Units
Specific volume	ft ³ /lb
Density	lb/ft ³
Enthalpy	Btu/lb
Entropy	Btu/lb F
C _p	Btu/lb F
C _v	Btu/lb F
Thermal conductivity	Btu/sec ft F
Prandtl number	None
C _p /C _v (k) or (γ)	None
Viscosity	lb/ft sec
Sonic velocity	ft/sec

The increments of pressure used for specific volume and density are small, to minimize the need for interpolation in use of tables. Most of the other properties vary only a little through the pressure and temperature range covered, so that large increments of pressure and temperature are used in the tables.

The content of the tables is discussed in some detail in Appendix C. At this point it is appropriate to note that very few published experimental data were found for oxygen-helium mixtures, so that comparisons of the table contents with experimental information were not generally possible. Moreover, some of the correlation schemes used to provide information for the pure gases could not be used for mixtures and required abbreviation of some tables. Accordingly, the possibility of some indeterminate error in helium-oxygen tables should be recognized, with the possible magnitude of error increasing with increasing pressure.

APPENDIX A

DENSITY OF SEAWATER

APPENDIX A

DENSITY OF SEAWATER

For practical purposes, in diving practice it can be assumed that the density of seawater is 64.0 lb/ft^3 , equivalent to 0.4444 psi for each foot of depth. The true value of seawater density varies from point to point in the oceans, but the error resulting from the assumption of a constant value should normally be less than 0.5 percent, equivalent to 5 feet in 1000 feet.

The true density of seawater depends upon its salinity, or salt content, its temperature, and its depth. Increasing salinity results in increasing density; increasing temperature results in decreasing density except for the range of temperature below 39 F, and water is slightly compressible so that increasing depth results in increasing density. The effects of these three variables are interrelated in a complex way so that it is difficult to calculate an exact value of density. Accordingly, calculation methods based on use of nomograms, or upon tabulated values, are commonly used by oceanographers needing high accuracy. All of these tables and nomograms are based on use of metric or international units. Illustrative density data that follow are shown in both metric and engineering units to provide a basis of comparison.

The surface temperature of seawater varies widely both with latitude and with the seasons, and can range from about 32 F to about 85 F. However, at a depth of 600 feet the worldwide variation of seawater temperature is smaller, 34 F to 68 F, and at 1500 feet the range is further reduced to 39 F to 54 F. The lower temperatures are found in northern areas and the higher temperatures are found near the equator. The difference in density at constant salinity of 35 parts per thousand for a temperature range of 39 F to 85 F would be 0.59 percent.

The salinity of seawater is defined as the total quantity of dissolved solids expressed as parts per thousand by weight. Salinity can vary widely from place to place, and is typically low in regions of high rainfall, or inflow from coastal rivers, and high in regions of high evaporation and low rainfall, as in the Sargasso Sea. The usual range of salinities is from 33 to 37 parts per thousand; the equivalent range of seawater density at a constant water temperature is 0.3 percent.

The third variable that affects seawater density is the compressibility of water under pressure. However, this factor is very small for present diving depths, and density is increased by only 0.01 percent at 1500 feet.

The naturally occurring gradients of temperature and salinity in the oceans are such that high temperatures and high salinities are found near the surface, and lower temperatures and lower salinities are found with increasing depth. The net result is a very small gradient of density, which increases with depth within a water mass of a common origin. Discontinuities of temperature-salinity gradients occur as the result of formation of layers of water from different sources and can be clearly distinguished when plotted on a temperature-salinity diagram.

Figure A-1 is a temperature-salinity diagram upon which are plotted some typical values of August surface salinities and temperatures for various coastal locations. (A-1)

The diagonal curves represent constant-density lines. The dashed lines are seawater density, in lb/ft^3 , and the solid lines are density anomaly, or the last two places in the specific gravity value; for example, the value "25" means that 1.025 is the specific gravity. Figure 1 can be used to interpret density in engineering units, simply by plotting salinity and temperature data in either English or metric units.

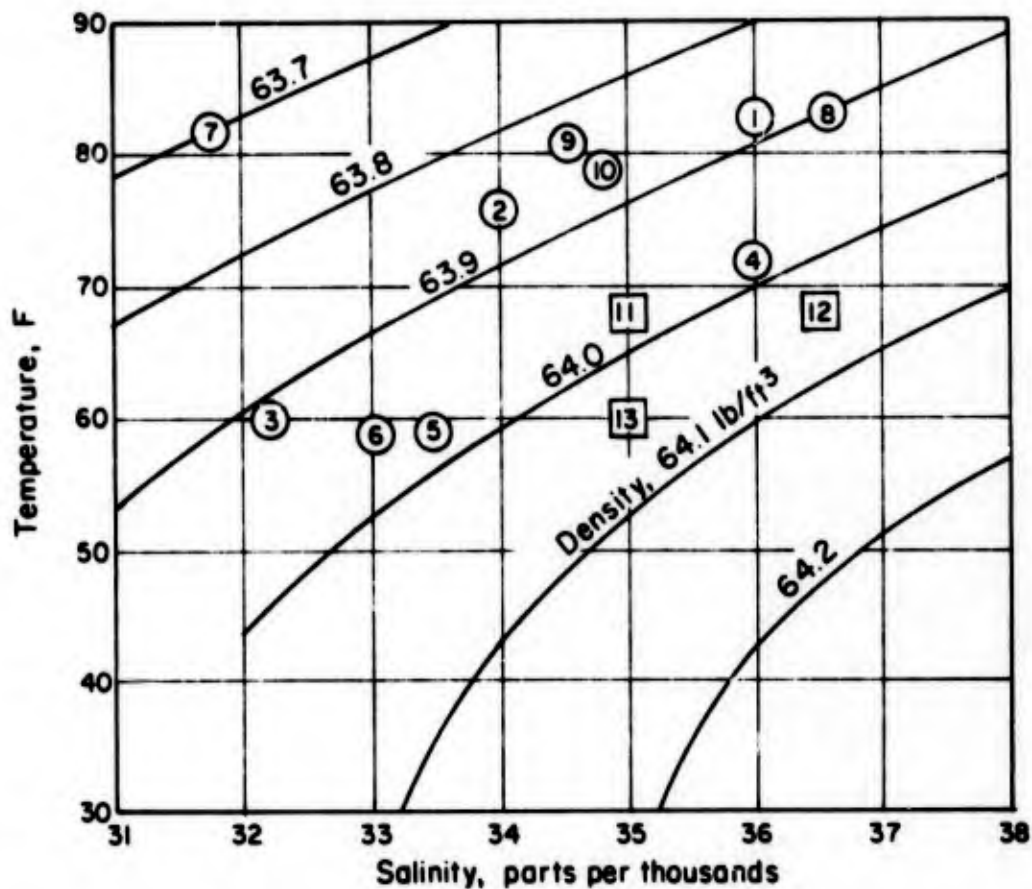


FIGURE A-1. TEMPERATURE-SALINITY-DENSITY PLOT SHOWING SEAWATER SURFACE CONDITIONS AT SELECTED LOCATIONS, AUGUST

<u>No.</u>	<u>Density</u>	<u>Location</u>
(1)	63.88	Florida East Coast
(2)	63.85	Atlantic Coast off Chesapeak Bay
(3)	63.91	Atlantic Coast off Maine
(4)	63.99	Atlantic Coast off Gibraltar
(5)	63.97	Pacific Coast off San Diego
(6)	63.95	Pacific Coast off Seattle
(7)	63.7	Panama - Pacific Coast
(8)	63.9	Panama - Atlantic Coast (Gulf)
(9)	63.82	Pacific - off South Japan
(10)	63.86	Pacific - at Hawaii
(11)	63.97	South Central Atlantic
(12)	64.03	North Central Atlantic
(13)	64.04	Central Indian Ocean

A-3 and A-4

In Figure A-1 it will be noted that the density of surface water under summer conditions is generally somewhat less than 64.0 lb/ft³. However, the density increases as the water cools in other months, and is greater with increasing depth, so that the value of 64.0 appears suitable as an all-around number for engineering use.

References

- (A-1) Von Arx, W. S., An Introduction to Physical Oceanography, Addison-Wesley Publishing Co., Inc., Reading, Mass. U.S. and London, England; 1962
Library of Congress Card No. 61-5026.
- (A-2) Williams, Jerome, Oceanography, Little, Brown and Company, Boston and Toronto, 1962; Library of Congress Card No. 62-18105.
- (A-3) Sverdrup, H. U., Johnson, Martin W., and Fleming, Richard H., The Oceans, Their Physics, Chemistry, and General Biology, Prentice-Hall, Inc., Englewood Cliffs, N. J., 1942.

APPENDIX B

DATA SOURCES AND CALCULATIONS FOR PSYCHROMETRIC CHARTS

APPENDIX B

DATA SOURCES AND CALCULATIONS FOR PSYCHROMETRIC CHARTS

In constructing the psychrometric charts for diving-gas mixtures, the following assumptions are made:

- (1) The molar specific heats of the dry-gas and constituents are constant over the temperature and pressure ranges covered by the charts. Specifically, the following values are assumed:

$$c_{pm} = 6.951 \text{ Btu/mole for air, oxygen, and nitrogen}$$

$$c_{pm} = 4.965 \text{ Btu/mole for helium.}$$

- (2) The molar specific heat of the dry-gas mixture is calculable by the ideal additive rule,

$$(c_{pm})_d = 6.951 (x_{O_2} + x_{N_2}) + 4.965 x_{He} \quad (\text{Btu/mole}_d)$$

- (3) The enthalpy of the moist-gas mixture is equal to the sum of the individual enthalpies of the dry-gas constituents and the water vapor, that is,

$$\begin{aligned} h_{\text{mixture}} &= h_d + h_w \\ &= (c_{pm})_d t_{db} + h_w \end{aligned}$$

Further, the value of enthalpy of the water vapor is taken as the value for saturated water vapor in the absence of other gases. No correction for effect of total pressure is used.

- (4) For total mixture pressures up to 100 psia, no correction for the effect of total pressure on the water-vapor saturation pressure is used. Above 100 psia, Poynting's equation,

$$\frac{\Delta p_{ws}}{p_{ws}} = \frac{v_l}{RT} \Delta p_{\text{total}}$$

where v_l is the specific volume of saturated liquid, is used as an estimated correction for total pressure. A derivation of Poynting's equation is presented in Reference B-1, p 236.

By limiting the psychrometric charts to 600-psia pressure, and with the assumption that oxygen and nitrogen concentration will be low at the higher pressures, it is believed that the accuracy of the charts will be adequate for normal engineering calculations, and that the effort required to obtain an estimate of the error is not justified at this time.

Charts 1 and 2 are primarily graphs of the relationship

$$x_w(t_{db}, p) = \frac{P_s}{P_{total} - P_s} ,$$

where the values of P_s have been taken from Reference (B-2), Chapter 21, Table 2. Values used in Reference B-2 are taken from an earlier paper by Goff and Gratch.

At pressures above 100 psia, values for P_s have been modified using Poynting's correction.

The enthalpy scale on Charts 1 and 2 (Figures 5 and 6) was set up using an approximate water-vapor enthalpy value of 20,000 Btu/mole_w. (This value is correct for saturated vapor at 112 F.) The approximate water-vapor enthalpy scale, expressed in Btu/mole of dry gas, is obtained by simply multiplying the moisture mole fraction, x_w , by 20,000 or

$$h_{m,w} = 20,000 \left(\frac{\text{Btu}}{\text{mole}_w} \right) x_w \left(\frac{\text{moles}_w}{\text{mole}_d} \right) .$$

Chart 3 (Figure 7) for the enthalpy correction is derived as follows: If Δh_w is defined as the difference between the true water-vapor enthalpy at t_{db} and the reference enthalpy value at 112 F, then the error is given by

$$\Delta h_{m,w} \left(\frac{\text{Btu}}{\text{mole}_d} \right) = x_w \Delta h_{m,w} \left(\frac{\text{Btu}}{\text{mole}_w} \right) .$$

To set up constant- $\Delta h_{m,w}$ lines, the above expression was simply rearranged to obtain

$$x_w = \Delta h_{m,w} \left(\frac{\text{Btu}}{\text{mole}_d} \right) / \Delta h_{m,w} \left(\frac{\text{Btu}}{\text{mole}_w} \right) ,$$

where Δh_w is a function of dry-bulb temperature.

Enthalpy values for saturated water vapor were also taken from Reference (B-2), Chapter 21, Table 2. As previously explained, no correction was made for the effect of total mixture pressure on the water-vapor enthalpy.

Because the constant-wet-bulb-temperature lines are lines of constant mixture enthalpy, their slope is determined as follows: The expression for mixture enthalpy was taken previously to be

$$h_{m, \text{mixture}} = (c_{pm})_d t_{db} + h_{m,w} .$$

Differentiating with respect to dry-bulb-temperature results in

$$\left. \frac{dh_{m, \text{mixture}}}{dt_{db}} \right]_{t_{wb}} = (c_{pm})_d + \frac{dh_{m,w}}{dt_{db}} = 0 ,$$

B-3 and B-4

Therefore,

$$\frac{dh_{m,w}}{dt_{db}} = - (c_{pm})d$$

along a constant wet-bulb temperature line.

References

- (B-1) Glasstone, Samuel, Thermodynamics for Chemists, 8th Printing, D. Van Nostrand, New York (1947).
- (B-2) ASHRAE Handbook of Fundamentals, American Society of Heating, Refrigerating, and Air-Conditioning Engineers, Inc., New York (1967).

APPENDIX C

DATA SOURCES AND INTERPOLATION TECHNIQUES FOR
CALCULATING TABLES OF PHYSICAL
PROPERTIES OF GASES

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APPENDIX C

DATA SOURCES AND INTERPOLATION TECHNIQUES
FOR CALCULATING TABLES OF PHYSICAL
PROPERTIES OF GASES

For the purpose of this discussion, the physical properties will be classed in two groups: (1) volumetric properties and (2) thermodynamic and transport properties.

Volumetric Properties

Density and specific volume values for the tables were calculated on a computer by manipulation of the virial equation of state:

$$\frac{PV}{RT} = 1 + \frac{B}{v} + \frac{C}{v^2} + \frac{D}{v^3} + \frac{E}{v^4} \dots$$

As a matter of convenience, all substances were treated in this general fashion.

At first glance this might be considered a poor technique for obtaining such information. For example, it can be noted that, because of the sizable pressure range, as many virial coefficients as possible need to be included and present knowledge of the higher ones is poor. This is indeed true, but the difficulty is lessened by two facts:

- (1) The higher coefficients can be estimated for pure substances and comparisons made of calculated and experimental compressibility results.
- (2) The virial expansion in $\frac{1}{v}$ converges more rapidly than the pressure expansion which would have been obviously more convenient in this situation. Therefore the $\frac{1}{v}$ expansion was used to minimize the effect of errors in the higher coefficients.

The balance was swung in favor of this technique by the He-O₂ mixtures. No PVT data were found for this system in the temperature range of interest. The temperature range is low enough that the quantum effects of helium cannot be dismissed summarily as insignificant. Under these conditions it seemed best to estimate the second and third virial coefficients where some precedent is available for making estimates from knowledge of pure-component behavior. This feeling was reinforced when it became clear that the other major alternative would fall short of the goal of producing desired information. This alternative would have been to determine pseudocritical constants of the mixtures by some scheme and from these calculate compressibilities (or for that matter other thermodynamic properties) by employing some corresponding-states treatment. For the very helium-rich mixtures, it developed that, even overlooking problems in making choices of pseudocritical constants, tabulations of Pitzer et al. and Hougen et al. were not capable of producing data over the temperature range of interest. (1)

Having selected the virial formalism, then the problem was to evaluate the virial coefficients. Consider the pure gases only for a moment. It was elected to evaluate B and C for pure substances using a Kihara potential function with a spherical core. Tables of values of functions are available for these. The basic information is contained in three references. (C-2 - C-4) The same sort of information for the fourth and fifth virial coefficients is to be found in one paper. (C-5) The information for the fourth and fifth coefficients is presented for the Lennard-Jones and Devonshire potential function only. However, this is the same as a Kihara potential with a spherical core of radius zero.

In order to use the information in References C-2 through C-5 it is necessary to have parameters which characterize the particular substance. These parameters were obtained by a variety of techniques from numerous sources, which are summarized in the listing at the end of this section.

For helium-oxygen mixtures it was necessary to calculate virial coefficients for the mixture. In the case of the second virial coefficient, this involves taking a sum of terms which include the composition of the given mixture, virial coefficients for both pure substances, and one virial coefficient characteristic of the two substances in the mixture. This is a common procedure described in most texts. In the listing which follows, the mixing rules used to calculate the term characteristic of the two substances in the mixture are given. Quantum corrections are included.

The procedure for estimating the third virial coefficient of the mixture is similar, although the rules stated for calculation of the mixed coefficients are empirical. No estimate for other virial coefficients was possible. Thus, results for all mixtures became more suspect with increasing pressure.

For pure substances it was possible to compare compressibility values calculated by this scheme with some experimental values. Experimental results were chosen with an eye toward getting a single piece of work which spanned a significant portion of the pressure-temperature range of interest. As it turned out, the very precise and generally well regarded work of Michels and co-workers was often suitable.

In the listing below, then, may be found statements regarding the source material for parameters of the virial coefficients and compressibility data and something about the difference between calculated and experimental compressibilities. Following that is a listing of the parameters used for each gas and listings of Fortran IV statements for two computer routines - one for pure substances, the other for mixtures. The Fortran statements supply the detailed information on the calculational procedure.

Air

Parameters from Hirschfelder, Curtiss and Bird, "The Molecular Theory of Gases and Liquids", Wiley, New York (1954), p 1111 (parameters from Ref. L).

Data from Michels, A., Wassenaar, T., and Van Seventer, W., "Isotherms of Air Between 0 and 75 C and Pressures up to 2200 Atm", Appl. Sci. Rev., A4, 52 (1953).

Difference between calculated and experimental compressibilities is 0.6 percent or less for the points examined.

Nitrogen

Parameters from Holleran, E. M., "A Dimensionless Constant Characteristic of Gases, Equations of State and Intermolecular Potentials", J. Phys. Chem., 73, 167 (1969).

Data from Michels, A., Lunbeck, R. J., and Wolkers, G. J., "Thermodynamical Properties of Nitrogen as Functions of Density and Temperature Between -125 and +150 C and Densities up to 760 Amagat", Physica, 17, 801 (1951).

Difference between calculated and experimental compressibilities is 0.6 percent or less for the points examined.

Oxygen

Parameters from Holleran, E. M., "A Dimensionless Constant Characteristic of Gases, Equations of State and Potentials", J. Phys. Chem., 73, 167 (1969) and "Linear Relation of Temperature and Density at Unit Compressibility Factor", J. Chem. Phys., 47, 5318 (1967) (with slight scaling).

Data from Michels, A., Schamp, H. W., and De Graaf, W., "Compressibility Isotherms of Oxygen", Physica 20, 1209 (1954).

Difference between calculated and experimental compressibilities is 0.6 percent or less for points examined.

Helium

Parameters from Hirschfelder, Curtiss and Bird, "The Molecular Theory of Gases and Liquids", Wiley, New York (1964), (quantum mechanical with slight scaling), p 1110.

Data from Canfield, F. B., Leland, T. W., and Kobayashi, R., "Compressibility Factors for Helium-Nitrogen Mixtures", J. Chem. Engr. Data, 10, 92 (1965) and Miller, J. G., Brandt, L. N., and Stroud, L., "Compressibility Factors for Helium and Helium-Nitrogen Mixtures", Bureau of Mines Rept. of Invest., No. 5845 (1961).

Difference between calculated and experimental compressibilities is 0.5 percent or less for points examined.

Carbon Dioxide

Parameters from Sherwood, A. E., and Prausnitz, J. M., "Intermolecular Potential Functions and the Second and Third Virial Coefficients", J. Chem. Phys., 41, 429 (1964) and Holleran, E. M., J. Phys. Chem., 73, 167 (1969) (for 4th and 5th virials).

Data from Van Huff, V. E., Houghton, G., and Coull, J., "Equation of State and Compressibilities for Gaseous CO₂ in the Range 0-600 C and 0 to 150 Atm", J. Chem. and Engr. Data, 8, 336 (1963); and Hilsenrath, et al., National Bureau of Standards Circular 564, U. S. Govt. Printing Office, Washington, D. C. (1955).

Difference between calculated and experimental compressibilities is 0.5 percent or less for points examined.

Helium-Oxygen Mixtures

Parameters -

Mixing rules to estimate second virial coefficient characteristic of the two substances:

$$2\sigma_{12} = \sigma_{11} + \sigma_{22}$$

$$\epsilon_{12}^2 = \epsilon_{11} \epsilon_{22}$$

$$2\rho_{12} = \rho_{11} + \rho_{22}$$

$$\rho = 2^{1/6} (\sigma - 2a)$$

$$\Lambda^* = \frac{6.6 \times 10^{-27}}{\sigma \sqrt{2\epsilon \left(\frac{1}{m_i} + \frac{1}{m_j} \right)^{-1}}}$$

In Fortran Statements

$$\text{BGAS} = 1.262 \times 10^{24} \sigma^3$$

$$\text{EGAS} = \epsilon$$

$$\text{ASTAR} = \frac{2a}{\sigma - 2a}$$

$$\text{LAMBDA} = (\Lambda^*)^2$$

m_i is the mass of molecule i .

Thermodynamic and Transport Properties

Values for thermodynamic and transport properties came from a variety of sources. In a program of this size there was no possibility of evaluation, correlation, and refinement of original data. Thus, reliance had to be placed in published efforts which did this. It might be thought then that NBS Circular 564 was uniformly adequate for the pure gases. However, the information in that publication falls quite short of the maximum pressure of interest.

In the list which follows this section, comments are made about the sources of data for the pure gases. Generally, thermodynamic data in these sources were merely converted into desired units. Usually, however, sonic velocities and Prandtl numbers had to be calculated. The Prandtl number calculation is merely the straightforward combination of other quantities in the tables. Sonic velocities were calculated using an approximate but convenient equation for nonideal gases. (6) Exceptions are noted in the list.

Also, usually, low-pressure viscosity and thermal-conductivity data for pure gases came from two sources. (7, 8) Pressure coefficients were generally estimated

using the scheme presented in two sources. (9, 10) The major exceptions to this were helium and air. For air there were other sources. For helium the schemes are invalid.

No data on the thermodynamic properties of helium-oxygen mixtures were found. Sonic velocity was calculated in the same manner as that used for pure substances. Data were found for low-pressure thermal conductivity⁽¹¹⁾ and the low-pressure viscosity was estimated using the reliable Wilke technique.⁽¹²⁾ Pressure coefficients for viscosity and thermal conductivity were not estimated because of a lack of confidence in available techniques in handling helium-containing mixtures. Estimates for the other thermodynamic properties of the mixtures were made assuming that ideal solutions were formed at every pressure and temperature. More refined procedures are not justified by present levels of knowledge.

Air

Enthalpy, entropy, C_p , and C_v , from Din, F., "Air", Thermodynamic Functions of Gases, F. Din, Editor, Butterworths, London (1962), Vol 2, p 1.

Sonic velocity calculated using Reference (6).

Thermal conductivity from Carroll, D. L., Hing, Y. L., and Stiel, L. I., "Thermal Conductivity of Air at Moderate and High Pressures", J. Chem. and Engr. Data, 13, 53 (1968).

Viscosity from Hing, Y. L., Carroll, D. L., and Stiel, L. I., "Viscosity of Gaseous Air at Moderate and High Pressures", J. Chem. and Engr. Data, 11, 540 (1966).

Nitrogen

Enthalpy, entropy; C_p , and C_v from Din, F., "Nitrogen", Thermodynamic Functions of Gases, F. Din, Editor, Butterworths, London (1961), Vol 3, p 72.

Sonic velocity calculated using Reference (6).

Thermal conductivity from References (8) and (10).

Viscosity from References (7) and (9) and Hilsenrath, et al., National Bureau of Standards Circular 564, U. S. Government Printing Office, Washington, D. C. (1955).

Oxygen

Enthalpy, entropy, sonic velocity, C_p/C_v , and C_v up to 100 atmospheres from Hilsenrath, et al., National Bureau of Standards Circular 564, U. S. Government Printing Office, Washington, D. C. (1955).

Enthalpy and entropy above 100 atm from Curl, R. F., Jr., and Pitzer, K. S., "Volumetric and Thermodynamic Properties of Fluids - Enthalpy, Free Energy and Entropy", Ind. Eng. Chem., 50, 265 (1958).

C_p and C_v above 100 atm from Edmister, W. C., "Application of Thermodynamics to Hydrocarbon Processing, Part XIII-Heat Capacities", Petrol. Refiner, 27, 609 (1948).

Viscosity from NBS Circular 564 and Reference (9).

Thermal Conductivity from References (8) and (10).

Helium

Enthalpy, entropy, C_p , C_v , and sonic velocity from Wilson, M. P., Jr., "Thermodynamic and Transport Properties of Helium", USAEC Report GA-1355 (1960).

Viscosity from USAEC Report GA-1355 (1960), Reference (8), and Ross, J. F., and Brown, G. M., "Viscosities of Gases at High Pressures", Ind. Eng. Chem., 49, 2027 (1957).

Thermal conductivity from Reference (8) and USAEC Report GA-1355 (1960).

Carbon Dioxide

Enthalpy, entropy, C_p , C_v , and sonic velocity from Hilsenrath, et al., National Bureau of Standards Circular 514, U. S. Government Printing Office, Washington, D. C. (1955).

Viscosity from NBS Circular 564 and References (7) and (9).

Thermal conductivity from References (8) and (10).

Helium-Oxygen Mixtures

X = mole fraction of helium

Enthalpy (Btu/mole) = X(Enthalpy He) + (1-X)(Enthalpy, O₂).

Entropy (Btu/mole R) = X(Entropy He) + (1-X)(Entropy O₂) -
R [X ln X + (1-X) ln (1-X)]

C_p (Btu)/mole, F) = X(C_{pHe}) + (1-X)(C_{pO_2})

C_v (Btu)/mole F) = X (C_{vHe}) + (1-X)(C_{vO_2})

Sonic velocity from Reference (6).

Viscosity from Reference (1), p 421 .

Thermal conductivity from Reference (11).

Note: Equations written here are on a mole basis for simplicity. Tables contain entries on a per pound basis.

References

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- (4) Sherwood, A. E., and Prausnitz, J. M., "Third Virial Coefficients for the Kihara, Exp-6, and Square Well Potentials", *J. Chem. Phys.*, 41, 413 (1964).
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- (7) DiPippo, R., and Kestin, J., "The Viscosity of Seven Gases Up to 500 C and Its Statistical Interpretation", Proceedings of the Fourth Symposium on Thermo-physical Properties, I. R. Moszynski, Editor, A. S. M. E., New York (1968), p 304.
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- (9) Jossi, J. A., Stiel, L. I., and Thodos, G., "The Viscosity of Substances in the Dense Gaseous and Liquid Phases", *A. I. Ch. E. Jour.*, 8, 59 (1962).
- (10) Stiel, L. I., and Thodos, G., "The Thermal Conductivity of Nonpolar Substances in the Dense Gaseous and Liquid Regions", *A. I. Ch. E. Jour.*, 10, 26 (1964).
- (11) Srivastava, B. N., and Barua, A. K., "Thermal Conductivity of Mixtures of Diatomic and Monatomic Gases", *J. Chem. Phys.*, 32, 427 (1960).


```

000573      CUEP=PP(J)/102.0570*ALV(I)
000577      IF (CUEP) 104*150*104
000580      104 CONTINUE
000583      AL=ALL/PP(J)
000586      AM=AM/PP(J)
000589      FFL=FUNC1(AL)
000592      FF=FUNC1(AM)
000595      2000 FORMAT (1X,10E12.4)
000598      CALL MACT (FUNC1,AL,AM,30*0.0*0.0*0.0*00010*V(J),IEM)
000601      10 IEM=15*150*105
000604      105 PRINT 2010*J,IEM
000607      2010 FORMAT (///10X,10HMINIMUM AT POINT I=,13,2X,2HJ=,13)/M IEM=,12 /
000610      /
000613      V(J)=1.0
000616      150 CONTINUE
000619      200 CONTINUE
000622      DO 210 J=1,NUM
000625      UN 210 J=1,NUM
000628      210 V(I,J)=CONST/V(I,J)
000631      IS=1
000634      300 CONTINUE
000637      NPT=
000640      ICOUNT=
000643      PRINT 2222
000646      2222 FORMAT (1X)
000649      PRINT 2021*ITITL(I,J),J,1.0)
000652      2021 FORMAT(1X) /// 27X,0A10)
000655      PRINT 2022
000658      2022 FORMAT( / 4X,5HTEMP IN TA,0HPRESSURE IN TA,10HTEMP IN TA, /
000661      1 7A,2HPT,5A,0HPST IN SA,3HATN,7A,2HJC,7A,2HWC,7A,2HSC,7A,2HWC,
000664      2 7A,2HFC,7A,2HWC,7A,2HWC,7A,2HWC,7A,2HWC,7A,2HWC,7A,2HWC,7A,2HWC,
000667      3 /)
000670      DO 301 J=1,NUM
000673      VMIN=1000.0
000676      DO 320 J=1,11
000679      IF (V(I,J)-VMIN) 315,320,320
000682      315 VMIN=V(I,J)
000685      320 CONTINUE
000688      IF (VMIN-10) 330,300,300
000691      300 FMA1(I)=10H11F9.3)
000694      FMA2(I)=FMA1(I)
000697      GO TO 300
000700      330 IF (VMIN-1) 350,300,300
000703      300 FMA1(I)=10H11F9.6)
000706      FMA2(I)=FMA1(I)
000709      GO TO 300
000712      350 FMA1(I)=10H11F9.5)
000715      FMA2(I)=FMA1(I)
000718      300 CONTINUE
000721      IF (I-1) 302,303,300
000724      302 PVP=16.49*PP(I)
000727      PRINT FMA1(I),P(I),PP(I),V(I,J),J=1,NUM)
000730      GO TO 305
000733      303 PRINT 2111)
000736      2111 FORMAT (///)
000739      305 CONTINUE
000742      ICOUNT=
000745      306 PRINT 2012* V(I),PP(I),V(I,J),J=1,NUM)
000748      2012 CONTINUE
000751      NPT=NPT+1
000754      10 INP=001 300,307,300
000757      307 PRINT 2021*ITITL(I,J),J,1.0)
000760      PRINT 2022
000763      NPT=
000766      ICOUNT=
000769      GO TO 301
000772      306 CONTINUE
000775      ICOUNT=ICOUNT+1
000778      IF (ICOUNT-5) 311,310,311
000781      310 PRINT 2014)
000784      2014 FORMAT (1X)
000787      ICOUNT=
000790      311 CONTINUE
000793      301 CONTINUE
000796      GO TO 1000500: 150
000800      DO 410 J=1,NUM
000803      DO 420 J=1,NUM
000806      410 V(I,J)=0.0*V(I,J)
000809      DO 420 I=1,0
000812      420 ITITL(I)=ITITL2(I)
000815      IS=2
000818      GO TO 300
000821      300 CONTINUE
000824      GO TO 5
000827      END

```

```

SUBROUTINE ATTN (X,Y,HA,AB,YO,TEMP,IER)
CALLRN INTERPOLATION SUBROUTINE
C THIS SUBROUTINE ALLOWS THE INDEPENDENT VARIABLE TO BE IN EITHER
C INCREASING OR DECREASING SEQUENCE. IF INCREASING SEQUENCE ONLY IS
C NEEDED THEN CARDS 16 AND 32 THRU 42 CAN BE REMOVED
C CALLING SEQUENCE...
C CALL ATTN (X,Y,HA,AB,YO,TEMP,IER)
C A IS A ONE DIMENSIONAL ARRAY OF INDEPENDENT
C VARIABLE INCREASING OR DECREASING
C Y IS A ONE DIMENSIONAL ARRAY OF DEPENDENT
C VARIABLE
C N IS NO. OF X,Y PAIRS
C K IS DEGREE OF INTERPOLATING POLYNOMIAL
C HA IS INDEP. VARIABLE ARGUMENT
C IF HA IS NOT WITHIN THE RANGE OF X, EXTRAPOLATION
C WILL BE PERFORMED
C YA IS INTERPOLATED RESULT
C TEMP ONE DIMENSIONAL ARRAY OF AT LEAST 200 WORDS PROVIDING
C SCRATCH STORAGE
C IER INDICATES WHETHER OR NOT EXTRAPOLATION WAS PERFORMED
C IER=1. INDICATES IT WAS PERFORMED
C IER=0. INDICATES IT WAS NOT PERFORMED

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```



```

000013      DIMENSION A(1),V(1),TEMP(1)
000013      IFA=0
000013      K=K+1
000015      IF (X(N)-X(1))100,10,10
000021  10  IF (X(N)-X(1)) 15,15,30
000024  15  IEX=1
000026  20  LL=8
000027      GO TO 200
000030  30  IF (X(N)-X0) 35,35,50
000033  35  IES=1
000035  40  LL=N-K
000037      GO TO 200
000037  50  LL=1
000040      LU=N
000041  60  IF (LU-LL-1) 100,100,70
000044  70  LI=(LL+LU)/2
000046      IF (X(LI)-X0) 80,80,90
000052  80  LL=LI
000054      GO TO 60
000054  90  LU=LI
000056      GO TO 60
000056  100 IF (X(N)-X(1)) 120,15,15
000061  120 IF (X(N)-X0) 170,35,35
000064  130  LL=1
000065      LU=N
000066  140 IF (LU-LL-1) 140,100,150
000071  150 LI=(LL+LU)/2
000073      IF (X(LI)-X0) 160,170,170
000077  160 LU=LI
000077      GO TO 140
000101  170 LL=LI
000101      GO TO 140
000103  180 LL=(LI+1)/2
000106  19  (LL) 20,200,150
000110  190 IF (LL+1-N) 200,200,40
000113  200 DO 210 I=1,K
000115      I=LL+1
000117      TEMP(I)=X(I)-X0
000124  210 TEMP(N+1)=V(I)
000132  DO 220 I=1,K
000133  DO 220 J=1,K
000134  220 TEMP(N+J)=((I,0/TEMP(J,1))-TEMP(I,1))/TEMP(N+1,1)+TEMP(J,1)-TEMP(I,1)
000157      V(N+K)=1
000162      RETURN
000162      END

```

```

000003      FUNCTION FUNCT(A)
000003      CUMPRN =V,CV,DV,EV,ISW,COEF
000011      F=X*(A*(COEF*A-1.0)+DV)-CV
000013      IF (ISW=1) 10,5,10
000015  5  FUNCT=F
000015      RETURN
000022  10  FUNCT=X*(A*F-DV)-EV
000022      RETURN
000022      END

```

```

SUBROUTINE BISECT(F,A,B,N,EF,EA,EB,AR,IFRR)
C
C      INITIALIZATIONS
000014      IT=0
000014      IERR=0
000015      XA=A
000017      XB=B
C
C      CHECK THAT F(A) AND F(B) DIFFER IN SIGN
000020      FA=F(A)
000031      IF (FA*F(B))1,1,1
000044  1  IERR=1
000046  2  RETURN
C
C      BISECT AND PERFORM CONVERGENCE TESTS
000047  3  X=(XA+XB)/2.0
000052      IT=IT+1
000056      F=X*F(AR)
000065  4  IF (ABS(F(AR)-EF)2,2,5
000070  5  IF (ABS(XA-AB)-ABS(XA-EB))12,2,6
C
C      TEST FOR LIMIT ON NUMBER OF BISECTIONS
000102  6  IF (IT-NI0,7,7
000104  7  IERR=2
000106      RETURN
C
C      SELECT NEW INTERVAL AND LOOP BACK FOR NEXT BISECTION
000106  8  IF (FA*F(X))9,2,10
000111  9  XA=X
000113      GO TO 3
000113  10  XB=X
000115      GO TO 3
000115      END

```

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BSECT001
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BSECT037
BSECT038

OXYGEN TEST CASE

ASIAN .5000E-01
 ALPHA 0.
 LAMMUA 0.
 LGAS .1271E+03
 MGAS .95A+0E+02
 NGAS .1100E+03
 GGAS .9420E+02

T	.19000E+01	D	.13130E+00	E	-.21700E-01		
T	.20000E+01	D	.12200E+00	E	-.10000E-01		
T	.21000E+01	D	.11750E+00	E	.15000E-02		
T	.22000E+01	D	.11450E+00	E	.12000E-01		
T	.23000E+01	D	.11310E+00	E	.21000E-01		
T	.24000E+01	D	.11200E+00	E	.24500E-01		
T	.25000E+01	D	.11310E+00	E	.34500E-01		
T	.30000E+01	D	.11900E+00	E	.57000E-01		
T	.40000E+01	D	.13110E+00	E	.64000E-01		
T	.50000E+01	D	.13410E+00	E	.62000E-01		
T	.15344E+01	F1	.94557E+00	F2	-.12012E-01	F3	-.00000E+00
T	.166667E+01	F1	.97291E+00	F2	.53500E-01	F3	-.00140E+00
T	.20000E+01	F1	.62600E+00	F2	.17714E+00	F3	-.04370E+00
T	.27222E+01	F1	.66772E+00	F2	.23534E+00	F3	-.33677E+00
T	.25000E+01	F1	.67015E+00	F2	.24003E+00	F3	-.22145E+00
T	.27777E+01	F1	.60710E+00	F2	.33301E+00	F3	-.13641E+00
T	.10710E+01	F1	.69017E+00	F2	.34610E+00	F3	-.74700E-01
T	.33333E+01	F1	.71045E+00	F2	.34370E+00	F3	-.10611E-01
T	.40000E+01	F1	.72020E+00	F2	.44020E+00	F3	.01010E-01
T	.10000E+01	C	.43250E+00	DC	.39220E-01		
T	.20000E+01	C	.41737E+00	DC	.36100E-01		
T	.22000E+01	C	.39325E+00	DC	.31130E-01		
T	.24000E+01	C	.37540E+00	DC	.27360E-01		
T	.26000E+01	C	.36240E+00	DC	.24410E-01		
T	.28000E+01	C	.35250E+00	DC	.22050E-01		
T	.30000E+01	C	.34500E+00	DC	.20110E-01		

MINIMUM TEST CASE

ALPHA .7500E-01
 BETA 0.
 GAMMA .1000E+03
 DELTA .6130E+02
 EPSILON .4540E+02
 ZETA .5000E+02

T	.10000E+01	D	.13130E+00	C	-.21700E-01		
T	.20000E+01	D	.12200E+00	C	-.10000E-01		
T	.21000E+01	D	.11750E+00	C	.15000E-02		
T	.27000E+01	D	.11450E+00	C	.17000E-01		
T	.23000E+01	D	.11310E+00	C	.21000E-01		
T	.24000E+01	D	.11200E+00	C	.20500E-01		
T	.25000E+01	D	.11310E+00	C	.30500E-01		
T	.30000E+01	D	.11400E+00	C	.57000E-01		
T	.40000E+01	D	.13110E+00	C	.00000E-01		
T	.50000E+01	D	.13400E+00	C	.07000E-01		
T	.10000E+01	F1	.319017E+00	F2	-.502610E+00	F3	-.17900E+01
T	.111111E+01	F1	.390730E+00	F2	-.37730E+00	F3	-.10000E+01
T	.17500E+01	F1	.056302E+00	F2	-.223450E+00	F3	-.120075E+01
T	.13000E+01	F1	.505300E+00	F2	-.10700E+00	F3	-.90700E+00
T	.15300E+01	F1	.505573E+00	F2	-.120120E-01	F3	-.00000E+00
T	.10000E+01	F1	.572910E+00	F2	.57500E-01	F3	-.001900E+00
T	.101010E+01	F1	.599050E+00	F2	.11050E+00	F3	-.50000E+00
T	.20000E+01	F1	.620010E+00	F2	.17710E+00	F3	-.00300E+00
T	.22222E+01	F1	.607727E+00	F2	.23530E+00	F3	-.330077E+00
T	.25000E+01	F1	.670155E+00	F2	.29000E+00	F3	-.221050E+00
T	.27770E+01	F1	.607101E+00	F2	.37300E+00	F3	-.130010E+00
T	.303030E+01	F1	.699177E+00	F2	.30010E+00	F3	-.70700E-01
T	.33333E+01	F1	.710050E+00	F2	.30370E+00	F3	-.100110E-01
T	.40000E+01	F1	.72020E+00	F2	.00200E+00	F3	.00100E-01
T	.22000E+01	C	.305050E+00	DC	.20000E-01		
T	.24000E+01	C	.30930E+00	DC	.20270E-01		
T	.26000E+01	C	.35740E+00	DC	.23050E-01		
T	.20000E+01	C	.30000E+00	DC	.21100E-01		
T	.30000E+01	C	.30200E+00	DC	.10320E-01		
T	.32000E+01	C	.337030E+00	DC	.17770E-01		

AIR TEST CASE

ALPHA 0.
 BETA 0.
 GAMMA 0.
 UGAS .10200E+03
 MGAS .00300E+02
 NGAS .10200E+03
 OGAS .00300E+02

T	.190000E+01	D	.131300E+00	E	-.217000E-01		
T	.200000E+01	D	.122000E+00	E	-.100000E-01		
T	.210000E+01	D	.117500E+00	E	.150000E-02		
T	.220000E+01	D	.114500E+00	E	.120000E-01		
T	.230000E+01	D	.113100E+00	E	.214000E-01		
T	.240000E+01	D	.112000E+00	E	.245000E-01		
T	.250000E+01	D	.113100E+00	E	.345000E-01		
T	.300000E+01	D	.119000E+00	E	.570000E-01		
T	.400000E+01	D	.132100E+00	E	.640000E-01		
T	.500000E+01	D	.134100E+00	E	.679000E-01		
T	.153000E+01	F1	.545573E+00	F2	-.120120E-01	F3	-.000000E+00
T	.166667E+01	F1	.572910E+00	F2	.515000E-01	F3	-.601000E+00
T	.200000E+01	F1	.624000E+00	F2	.177192E+00	F3	-.443197E+00
T	.222222E+01	F1	.647727E+00	F2	.235343E+00	F3	-.330477E+00
T	.250000E+01	F1	.670155E+00	F2	.240039E+00	F3	-.221050E+00
T	.277777E+01	F1	.687101E+00	F2	.333401E+00	F3	-.136410E+00
T	.303030E+01	F1	.699177E+00	F2	.366190E+00	F3	-.747040E-01
T	.333333E+01	F1	.710456E+00	F2	.343760E+00	F3	-.146110E-01
T	.400000E+01	F1	.724262E+00	F2	.440245E+00	F3	.016100E-01
T	.200000E+01	C	.437130E+00	DC	.393500E-01		
T	.220000E+01	C	.409900E+00	DC	.374600E-01		
T	.240000E+01	C	.389450E+00	DC	.247800E-01		
T	.260000E+01	C	.373000E+00	DC	.245500E-01		
T	.280000E+01	C	.361760E+00	DC	.239600E-01		
T	.300000E+01	C	.352370E+00	DC	.218500E-01		
T	.320000E+01	C	.344940E+00	DC	.200800E-01		
T	.340000E+01	C	.338000E+00	DC	.145000E-01		

CARBON DIOXIDE TEST CASE

ASTAR .65000E+00
 ALPHA .25000E-01
 LAMBDA 0.
 EGAS .44170E+03
 HUFFS .64440E+02
 M .21110E+03
 GWA .50500E+02

T	.750000E+00	D	-.100400E+02	E	-.145900E+03		
T	.800000E+00	D	-.930000E+01	E	-.774100E+02		
T	.875000E+00	D	-.324300E+01	E	-.274600E+02		
T	.950000E+00	D	-.944000E+00	E	-.640000E+01		
T	.100000E+01	U	-.276400E+00	E	-.246000E+01		
T	.105000E+01	D	.691000E-01	E	-.116500E+01		
T	.110000E+01	D	.237400E+00	E	-.416700E+00		
T	.120000E+01	D	.335400E+00	E	.150000E-01		
T	.130000E+01	U	.315700E+00	E	.361000E-01		
T	.140000E+01	D	.269500E+00	E	-.220000E-02		
T	.150000E+01	D	.225000E+00	E	-.303000E-01		
T	.160000E+01	D	.184400E+00	E	-.413000E-01		
T	.500240E+00	F1	-.360134E+00	F2	-.208537E+01	F3	-.454377E+01
T	.600060E+00	F1	-.298657E+00	F2	-.194740E+01	F3	-.430352E+01
T	.625000E+00	F1	-.279493E+00	F2	-.181531E+01	F3	-.407136E+01
T	.645140E+00	F1	-.182553E+00	F2	-.168741E+01	F3	-.384064E+01
T	.689660E+00	F1	-.749400E-01	F2	-.144495E+01	F3	-.341064E+01
T	.714240E+00	F1	-.242010E-01	F2	-.133004E+01	F3	-.321478E+01
T	.735290E+00	F1	-.150660E-01	F2	-.124099E+01	F3	-.305020E+01
T	.550000E+00	C	-.500110E+00	UC	.417930E+02		
T	.600000E+00	C	.175200E+00	UC	.303720E+02		
T	.650000E+00	C	.468940E+00	UC	.223560E+02		
T	.700000E+00	C	.509030E+00	UC	.177120E+02		
T	.750000E+00	C	.568940E+00	UC	.117300E+02		
T	.800000E+00	C	.509420E+00	UC	.112600E+02		

HELIUM TEST CASE

ASTAR 0.
 ALPHA 0.
 LAMBDA .71300E+01
 EGAS .10720E+02
 HGAS .21640E+02
 MGAS .10770E+02
 GUS .21640E+02

T .100000E+02 D .115000E+00 L .340000E-01
 T .150000E+02 D .964000E-01 L .275000E-01
 T .200000E+02 D .832000E-01 L .204000E-01
 T .200000E+02 F1 .752000E+00 F2 .541000E+00 F3 .371000E+00
 T .250000E+02 F1 .746400E+00 F2 .545400E+00 F3 .373700E+00
 T .333333E+02 F1 .737400E+00 F2 .536100E+00 F3 .371000E+00
 T .500000E+02 F1 .722400E+00 F2 .514100E+00 F3 .354500E+00
 T .600000E+01 C .307700E+00 UC .445000E+00
 T .100000E+02 C .206100E+00 UC .576000E+00
 T .100000E+03 C .142500E+00 UC .790000E-01

```

PROGRAM PROG2 (INPUT,OUTPUT,TAPL60=INPUT)
DIMENSION IMAT(3)
DIMENSION ID(4),IM(25),ASTAR(25),ASTAR(J),LAMBDA(3),EGAS(3),HGAS(3)
DIMENSION TC(25,2),C(25,2),TF(25),F1(25),F2(25),F3(25), TEMP(50)
DIMENSION I(11),CC(12),B(3),P(300),PPP(300),V(300,11)
DIMENSION FMAT(3),FMAT2(3),X(13)
COMMON HM,CM,CDEF
REAL LAMBDA,LENDPT
EXTERNAL FUNCT
DATA (IM(1),I(1),A)/1,0,2,0,0,0,0,0,0,0,12,0,25,0,100,0/
DATA (IMAT(1),I(1),B)/0,0,145,0,00203,0,03247,0,01026,0,01057,
10,00732,0,007025,0,0006126/
PRINT 9999
9999 FORMAT ('PM REVERSE PAPER')
FMAT(1)=10M(F9.0,F9.2)
FMAT(2)=10M (F8.2,2X)
FMAT(3)=10M(F8.3)
FMAT2(1)=10M(9,F9.0)
FMAT2(2)=10M(F8.2,2X)
FMAT2(3)=10M(F8.3)
5 READ 1000, (ID(I),I=1,6)
1000 FORMAT (I4)
IF (EOF,6) 10,15
10 CALL EXIT
15 CONTINUE
PRINT 2000, (ID(I),I=1,6)
2000 FORMAT (I4,4I4)
READ 1001,NUMT,NUMP,NUMR,LENDPT,RENDPT
1001 FORMAT (3I10,7E10.0)
ON 20 I=3
READ 1002,ASTAR(1),LAMBDA(1),EGAS(1),HGAS(1)
1002 FORMAT (E10.0)
PRINT 2001,ASTAR(1),LAMBDA(1),EGAS(1),HGAS(1)
2001 FORMAT (//5X,5HASTAR,3X,E15.5,5X,6HLAMBDA,2X,E15.5,5X,4HEGAS,4X,
1 F15.5,5X,4MHGAS,4X,E15.5 /)
20 CONTINUE
ON 30 J=1,2
HEAD 1003,NTC(J)
1003 FORMAT (I10)
INDEX=NTC(J)
READ 1002, (TC(I,J),C(I,J),I=1,INDEX)
PRINT 2002, (TC(I,J),C(I,J),I=1,INDEX)
2002 FORMAT (10X,1HT,E15.6,10X,1MC,E15.6)
30 CONTINUE
READ 1003,NTF
HEAD 1002, (TF(1),F1(1),F2(1),F3(1),I=1,NTF)
000204 X(1)=0.097
000725 X(2)=0.095
000277 X(3)=0.093
000730 X(4)=0.09
000732 X(5)=0.085
000233 X(6)=0.07
000235 X(7)=0.06
000236 X(8)=0.9
    
```

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000241      X(1)=0.05
000243      X(2)=0.0
000244      X(3)=0.00
000246      X(4)=0.04
000247      X(5)=0.04
000251      N=13
000252      INTDEG=7
000253      DN 24 I=1,11
000254      TRELVM(1)=272.00*(1-1)*5.556
000262      26 CONTINUE
000264      NP=117
000265      DO 27 I=1,100
000266      27 P(I)=(1-I)*10.0
000274      DO 28 I=1,11
000276      28 P(I)=100.0*(1-1)*50.0
000280      DO 29 I=1,111
000281      29 PPP(I)=(P(I)*0.0000010.006)/10.006
000281      P(112)=400.0
000281      PPP(112)=34.071
000281      DO 25 I=1,65
000282      P(I)=112-P(I)-111+100.0
000285      25 PPP(I)=112-P(I)-112/10.006
000286      DN 40 I=1,NTP
000287      PHINT 2003,TF(1),F(11),F(11),F(11)
000287      2003 F0HMT 1/10R,1MT,E15.0,0R,2MF1,E15.0,0R,2MF2,E15.0,0R,2MF3,E15.0)
000287      40 CONTINUE
000287      PHINT 2222
000287      2222 FORMAT(1NA)
000287      DO 200 J=1,NX
000287      DO 200 I=1,NUMT
000287      DO 100 L=1,3
000287      TSTAR=TRELVM(L)/EGAS(L)
000287      CALL AITRN(TH,ASTAR,0,INTDEG,TSTAR,HT,TEMP,(FR)
000287      CALL AITRN(TH,F3,NTP,INTDEG,TSTAR,F3I,TEMP,(FN)
000287      110 CALL AITRN(TH,F2,NTP,INTDEG,TSTAR,F2I,TEMP,(FN)
000287      CALL AITRN(TH,F1,NTP,INTDEG,TSTAR,F1I,TEMP,(FR)
000287      TRP=1.4162140F37
000287      TRP2=1.7707630ASTAR(1)0F27
000287      TRP3=1.7673640ASTAR(1)0020F1*ASTAR(1)003
000287      DENOM(1)=0*ASTAR(1)003
000287      120 H(1)=RGAS(1)0*(TH)-TRP-TRP3)/DENOM(1)0LAMBDA(1)0MT
000287      IF (I-1) 130,100,190
000287      130 CALL AITRN(TH,C(1),C(1),C(1),C(1),INTDEG,TSTAR,CT,TEMP,(FR)
000287      CC(1)=RGAS(1)0020CT
000287      100 CONTINUE
000287      HMAX(TH)0020R(1)02.00R(1)0*(1.0-X(1)0)0(3)0(1),0-X(1)0020R(2)
000287      CMX(1)0030CC(1)03.00X(1)0020(1.0-X(1)0)0 (CC(1)0020CC(2)000.33
000287      1) 03.00X(1)0*(1.0-X(1)0)0020CC(1)002000.333 0(1.0-X(1)0)
000287      2 030CC(2)
000287      ALL=(FNPT002.00TRELVM(1)
000287      ARR=FNPT002.00TRELVM(1)
000287      DO 180 J=1,NUMP
000287      COEF=PPP(J)/(02.057*TRELVM(1)
000287      AL=ALL/PPP(J)
000287      AR=ARR/PPP(J)
000287      CALL NSCT (FNCT,AL,AR,30.0,0.0,0.0,0.0005,VIJ,1),IEMH)
000287      IF (IEMH) 170,100,170
000287      170 PRINT 210,VIJ,IEMH
000287      2010 FORMAT(//10R,2HINTROUBLE AT POINT 11=12,2X,2M1=12,2X,2M2=12,2X,
000287      1 2H1ERR=12 /)
000287      180 CONTINUE
000287      200 CONTINUE
000287      CONV=07.028*(1.003)*R(1)032.00*(1.0-R(1)0)
000287      DO 210 J=1,NUMP
000287      DO 210 I=1,NUMT
000287      VIJ(I)=CONV/VIJ(I)
000287      210 CONTINUE
000287      ICM=1
000287      IMAT(1)=10M DENS1
000287      IMAT(2)=10MTY*LOS/CUM
000287      IMAT(3)=10MIC FT
000287      300 CONTINUE
000287      NP=0
000287      ICCOUNT=0
000287      XM(1)=0-X(1)0
000287      AM=0.0030X(1)032.00XM
000287      PRINT 2100,IMAT
000287      2100 FORMAT(1M),50R,3A10)
000287      PHINT 2100 ,XM),X(1)0,AM)
000287      2100 FORMAT 1 /10R,FS,3,7M OXYGEN,3X,FS,3,7M HELIUM,3X,20MAVENAGE M
000287      MOLECULAR WEIGHT ,FS,3 )
000287      PHINT 2011
000287      2011 FORMAT 1 / 4X,5HDEPTH,7X,AMHPRESSURE,40X,10HTEMPERATURE,F /
000287      1 7X,2MFT,5X,6MPSIA,5X,3MATH,7X,2M30,7X,2M40,7X,2M50,7X,2M60
000287      2 7X,2M70,7X,2M80,7X,2M90,7X,3M100,6R,3M110,6A,3M120,6A,3M13
000287      30 /)
000287      DO 301 I=1,NUMP
000287      VMIN=10000.0
000287      DO 320 J=1,11
000287      IF (V(I,J)-VMIN) 315,320,320
000287      315 VMIN=V(I,J)
000287      320 CONTINUE
000287      IF (VMIN=10) 330,340,340
000287      340 FMAT1(3)=10M11F9.3 )
000287      FMAT2(3)=FMAT1(3)
000287      GO TO 340
000287      330 IF (VMIN=1) 350,360,360
000287      360 FMAT1(3)=10M11F9.6 )
000287      FMAT2(3)=FMAT1(3)
000287      GO TO 360
000287      350 FMAT1(3)=10M11F9.5 )
000287      FMAT2(3)=FMAT1(3)
000287      300 CONTINUE
000287      IF (I=NP) 302,303,304
000287      302 PP=10.000*PPP(1)
000287      PRINT FMAT1,P(1),PP,PPP(1), (VIJ,J),J=1,NUMT)
000287      GO TO 305
000287      303 PRINT 2111
000287      2111 FORMAT (/// )
000287      NP=NP+2

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C-17 and C-18

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001012      ICOUNT=0
001013 304 PRINT F=ATZ,P(1),PPP(1),V(1),J1,J2,NUMT)
001034 305 CONTINUE
001034      NP=NP+1
001034      IF (NP=0) 304,307,306
001040 307 PRINT 2300,IMAT
001040      PRINT 2100 ,AM1,AM(1),AMW
001040      PRINT 2011
001044      NP=0
001045 ICOUNT=0
001046 GO TO 301
001046 306 CONTINUE
001046 ICOUNT=ICOUNT+1
001070 IF (ICOUNT=5) 311,310,311
001072 310 PRINT 2014
001076 2014 FORMAT (1X)
001076 ICOUNT=0
001077 311 CONTINUE
001077 301 CONTINUE
001102 GO TO (400,290) ISU
001110 400 DO 410 1=1,NUMP
001112 DO 410 J=1,NUMT
001113 *10 V(1,J)=1.0/V(1,J)
001125 ISU=2
001126 IMAT(1)=10M SPECIFIC
001130 IMAT(2)=10MVOLUME/CUB
001131 IMAT(3)=10MIC FT/LB
001133 GO TO 300
001133 290 CONTINUE
001136 GO TO 5
001136 END
000003      FUNCTION FUNCT(X)
000003      COMMON AM,CM,CNEF
000003      FUNCT=X*(X*(CNEF*X-.01-AM))-CM
000011      RETURN
000012      END
    
```

MINIATURE TEST CASE

ASTAR	0.	LAMBDA	.71300E+01	EGAS	.10220E+02	00AS	.21400E+02
ASTAR	.50000E-01	LAMBDA	0.	EGAS	.12710E+03	00AS	.55000E+02
ASTAR	.25000E-01	LAMBDA	.01100E+00	EGAS	.30040E+02	00AS	.36110E+02
T	.400000E+01	C	.307700E+00				
T	.100000E+02	C	.296120E+00				
T	.100000E+03	C	.142540E+00				
T	.190000E+01	C	.432550E+00				
T	.200000E+01	C	.417370E+00				
T	.220000E+01	C	.393250E+00				
T	.240000E+01	C	.375400E+00				
T	.260000E+01	C	.362420E+00				
T	.280000E+01	C	.352400E+00				
T	.300000E+01	C	.345090E+00				
T	.200000E-01	F1	.626010E+00	F2	.17710E+00	F3	-.463000E+00
T	.222227E-01	F1	.647730E+00	F2	.23936E+00	F3	-.330400E+00
T	.250000E-01	F1	.670150E+00	F2	.29094E+00	F3	-.221050E+00
T	.277770E-01	F1	.697100E+00	F2	.33340E+00	F3	-.136010E+00
T	.303730E-01	F1	.69910E+00	F2	.36610E+00	F3	-.747000E-01
T	.333333E-01	F1	.710660E+00	F2	.39376E+00	F3	4.166100E-01
T	.400000E-01	F1	.72820E+00	F2	.44028E+00	F3	-.016100E-01
T	.454545E-01	F1	.737730E+00	F2	.46625E+00	F3	.136070E+00
T	.500000E-01	F1	.743470E+00	F2	.48259E+00	F3	.172070E+00
T	.62500E-01	F1	.753250E+00	F2	.51234E+00	F3	.239100E+00
T	.714290E-01	F1	.757010E+00	F2	.52926E+00	F3	.270270E+00
T	.833330E-01	F1	.759740E+00	F2	.53673E+00	F3	.299610E+00
T	.900000E-01	F1	.760610E+00	F2	.54163E+00	F3	.313030E+00
T	.100000E+02	F1	.76100E+00	F2	.54509E+00	F3	.325000E+00
T	.125000E+02	F1	.760410E+00	F2	.55195E+00	F3	.340710E+00
T	.200000E+02	F1	.752750E+00	F2	.55099E+00	F3	.371500E+00
T	.250000E+02	F1	.746910E+00	F2	.54586E+00	F3	.373720E+00
T	.333333E+02	F1	.73760E+00	F2	.53697E+00	F3	.370950E+00
T	.500000E+02	F1	.72294E+00	F2	.51407E+00	F3	.359470E+00

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SUMMARY OF DATA TABULATIONS

<u>Gas or Mixture</u>	<u>Page Numbers</u>			<u>Molecular Weight</u>	<u>Oxygen, wt %</u>
	<u>Density</u>	<u>Volume</u>	<u>Others</u>		
Air	2-3	4-5	6-7	28.97	
Oxygen	8-9	10-11	12-13	32.00	
Helium	14-15	16-17	18-19	4.00	
Nitrogen	20-21	22-23	24-25	28.00	
Carbon dioxide	26	27	30-31	44.01	
60% O ₂ , 40% He	32-33	34-35	36-37	20.80	92.308
40 60	38-39	40-41	42-43	15.20	84.211
32 68	44-45	46-47	48-49	12.96	79.012
20 80	50-51	52-53	54-55	9.60	66.667
15 85	56-57	58-59	60-61	8.26	58.537
10 90	62-63	64-65	66-67	6.80	47.059
6 94	68-69	70-71	72-73	5.68	33.803
3 97	74-75	76-77	78-79	4.8	19.835
1.5 98.5	80-81	82-83	--	4.42	10.860
1.0 99.0	84-85	86-87	88-89	4.28	7.477
0.7 99.3	90-91	92-93	--	4.20	5.338
0.5 99.5	94-95	96-97	98-99	4.14	3.865
0.3 99.7	100-101	102-103	--	4.09	2.350

DENSITY

AIR

AIR DENSITY LBS./CUBIC FT.

Table with columns: DEPTH FT, PRESSURE PSIA, ATM, 30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130. Rows represent depth intervals from 0 to 740 feet.

AIR

T-6

AIR
ENTHALPY-BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	171.06	176.66	181.46	186.27	191.08	195.89
30.00	171.75	176.56	181.37	186.18	191.00	195.82
50.00	171.60	176.62	181.24	186.07	190.90	195.73
100.00	171.23	176.08	180.93	185.76	190.64	195.49
200.00	170.49	175.61	180.31	185.24	190.12	195.01
300.00	169.74	174.73	179.69	184.65	189.60	194.56
400.00	168.99	174.04	179.07	184.00	189.07	194.06
500.00	168.23	173.36	178.45	183.51	188.55	193.58
1000.00	164.53	169.49	175.39	180.72	186.02	191.26
2000.00	157.83	163.92	169.89	175.73	181.49	187.14
3000.00	153.00	159.59	165.95	172.15	178.26	184.24
4000.00	150.26	156.96	163.50	169.91	176.24	182.44
5000.00	148.05	155.61	162.25	168.70	175.25	181.61

AIR
CV-BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.172	.172	.172	.172	.172	.172
30.00	.172	.172	.172	.172	.172	.172
50.00	.172	.172	.172	.172	.172	.172
100.00	.172	.172	.172	.172	.172	.172
200.00	.173	.172	.172	.172	.172	.172
300.00	.174	.173	.173	.173	.173	.173
400.00	.174	.174	.174	.174	.174	.174
500.00	.175	.175	.175	.175	.175	.175
1000.00	.179	.178	.177	.177	.176	.176
2000.00	.183	.182	.180	.179	.179	.178
3000.00	.186	.184	.182	.181	.181	.181
4000.00	.188	.187	.185	.184	.183	.183
5000.00	.190	.189	.186	.185	.184	.183

AIR
ENTROPY-BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.0053	.0169	.0242	.0330	.0416	.0499
30.00	.0764	.0860	.0953	.1042	.1128	.1210
50.00	.0306	.0403	.0576	.0664	.0750	.0833
100.00	.7768	.7865	.7958	.8047	.8134	.8217
200.00	.7300	.7390	.7493	.7583	.7671	.7755
300.00	.7002	.7101	.7197	.7280	.7377	.7462
400.00	.6755	.6855	.6952	.7044	.7134	.7220
500.00	.6611	.6714	.6812	.6905	.6996	.7082
1000.00	.6070	.6179	.6283	.6381	.6476	.6566
2000.00	.5437	.5558	.5673	.5780	.5883	.5980
3000.00	.5053	.5193	.5315	.5429	.5538	.5641
4000.00	.4801	.4935	.5061	.5179	.5292	.5399
5000.00	.4602	.4737	.4865	.4985	.5101	.5209

AIR
THERMAL CONDUCTIVITY-BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000010)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3.93	4.06	4.19	4.32	4.45	4.58
30.00	3.93	4.07	4.20	4.33	4.46	4.59
50.00	3.95	4.08	4.21	4.34	4.47	4.59
100.00	3.98	4.11	4.24	4.37	4.49	4.62
200.00	4.03	4.16	4.29	4.42	4.54	4.66
300.00	4.09	4.22	4.34	4.47	4.59	4.71
400.00	4.15	4.27	4.40	4.52	4.64	4.76
500.00	4.21	4.33	4.45	4.57	4.68	4.80
1000.00	4.54	4.64	4.76	4.85	4.95	5.05
2000.00	5.30	5.35	5.41	5.47	5.53	5.61
3000.00	6.21	6.17	6.16	6.17	6.19	6.22
4000.00	7.20	7.08	7.01	6.95	6.91	6.90
5000.00	8.17	7.99	7.87	7.76	7.67	7.61

AIR
CP-BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.2401	.2431	.2462	.2493	.2504	.2506
30.00	.2406	.2406	.2406	.2407	.2408	.2409
50.00	.2413	.2412	.2412	.2414	.2413	.2413
100.00	.2430	.2424	.2426	.2425	.2424	.2424
200.00	.2465	.2454	.2454	.2450	.2444	.2446
300.00	.2494	.2440	.2442	.2470	.2471	.2467
400.00	.2533	.2521	.2510	.2501	.2495	.2489
500.00	.2500	.2542	.2530	.2527	.2518	.2510
1000.00	.2750	.2714	.2683	.2657	.2634	.2615
2000.00	.3077	.3004	.2947	.2895	.2848	.2809
3000.00	.3243	.3215	.3143	.3080	.3021	.2972
4000.00	.3371	.3332	.3230	.3160	.3125	.3075
5000.00	.3410	.3352	.3249	.3251	.3204	.3150

AIR
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.70	.70	.70	.69	.69	.69
30.00	.70	.70	.70	.69	.69	.69
50.00	.70	.70	.70	.69	.69	.69
100.00	.70	.70	.70	.69	.69	.69
200.00	.71	.71	.70	.70	.70	.70
300.00	.72	.71	.71	.70	.70	.70
400.00	.72	.72	.71	.70	.70	.70
500.00	.73	.72	.72	.71	.70	.70
1000.00	.76	.75	.74	.73	.72	.72
2000.00	.82	.80	.79	.77	.75	.74
3000.00	.86	.83	.81	.80	.79	.78
4000.00	.85	.84	.83	.82	.81	.80
5000.00	.84	.84	.83	.82	.81	.80

AIR
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.40	1.40	1.40	1.40	1.40	1.40
30.00	1.40	1.40	1.40	1.40	1.40	1.40
50.00	1.41	1.40	1.40	1.40	1.40	1.40
100.00	1.41	1.41	1.41	1.41	1.41	1.41
200.00	1.43	1.42	1.42	1.42	1.41	1.41
300.00	1.44	1.43	1.43	1.43	1.42	1.42
400.00	1.45	1.44	1.44	1.44	1.43	1.43
500.00	1.46	1.45	1.45	1.45	1.44	1.44
1000.00	1.53	1.52	1.52	1.51	1.49	1.49
2000.00	1.69	1.66	1.64	1.61	1.59	1.57
3000.00	1.77	1.75	1.72	1.70	1.66	1.64
4000.00	1.80	1.77	1.76	1.74	1.71	1.68
5000.00	1.80	1.78	1.77	1.76	1.73	1.71

AIR
VISCOSITY LB/FT SEC
(MULTIPLY TABLE ENTRY BY .000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.148	1.184	1.212	1.244	1.275	1.310
30.00	1.149	1.184	1.213	1.245	1.276	1.311
50.00	1.151	1.187	1.215	1.246	1.278	1.313
100.00	1.156	1.192	1.220	1.251	1.282	1.317
200.00	1.165	1.200	1.228	1.259	1.290	1.324
300.00	1.174	1.208	1.237	1.267	1.298	1.332
400.00	1.183	1.217	1.246	1.276	1.306	1.339
500.00	1.192	1.225	1.254	1.284	1.314	1.347
1000.00	1.248	1.277	1.305	1.332	1.359	1.390
2000.00	1.400	1.416	1.435	1.456	1.476	1.499
3000.00	1.501	1.503	1.508	1.507	1.606	1.623
4000.00	1.790	1.784	1.774	1.771	1.772	1.779
5000.00	2.016	1.986	1.965	1.948	1.934	1.932

AIR
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1084	1104	1124	1149	1170	1190
30.00	1084	1107	1129	1150	1170	1191
50.00	1085	1107	1129	1150	1171	1191
100.00	1085	1104	1131	1152	1172	1193
200.00	1086	1110	1133	1155	1176	1197
300.00	1088	1112	1136	1158	1179	1200
400.00	1089	1113	1138	1161	1182	1204
500.00	1090	1114	1141	1164	1185	1207
1000.00	1106	1134	1162	1185	1207	1231
2000.00	1165	1194	1224	1248	1270	1293
3000.00	1232	1265	1295	1319	1340	1361
4000.00	1313	1344	1375	1400	1421	1441
5000.00	1401	1432	1461	1491	1523	1544

OXYGEN

T-12

OXYGEN
ENTHALPY-BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	106.24	110.62	115.01	119.41	123.81	128.22
30.00	106.11	110.50	114.89	119.30	123.70	128.12
50.00	105.93	110.33	114.74	119.15	123.57	128.00
100.00	105.69	109.93	114.36	118.80	123.26	127.69
200.00	104.6	109.1	113.6	118.1	122.6	127.1
300.00	103.7	108.3	112.8	117.4	121.9	126.5
400.00	102.8	107.4	112.1	116.7	121.3	125.9
500.00	102.0	106.6	111.3	116.0	120.6	125.3
1000.00	97.5	102.6	107.6	112.5	117.4	122.2
2000.00	89.7	95.4	101.2	106.7	111.8	116.8
3000.00	82.4	88.9	95.5	101.5	106.9	111.9
4000.00	78.4	85.2	91.8	98.1	103.9	109.6
5000.00	75.2	81.6	88.0	95.3	101.1	107.3

OXYGEN
CV-BTU/LB-F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.1565	.1564	.1572	.1576	.1581	.1585
30.00	.1566	.1569	.1573	.1577	.1581	.1586
50.00	.1567	.1569	.1573	.1577	.1582	.1586
100.00	.1568	.1571	.1574	.1578	.1583	.1588
200.00	.1572	.1574	.1577	.1580	.1584	.1589
300.00	.1575	.1577	.1580	.1584	.1588	.1592
400.00	.1577	.1581	.1583	.1587	.1591	.1595
500.00	.1583	.1584	.1588	.1591	.1595	.1599
1000.00	.1580	.1580	.1580	.1582	.1583	.1586
2000.00	.17	.17	.17	.16	.16	.16
3000.00	.18	.18	.18	.17	.17	.17
4000.00						
5000.00						

OXYGEN
ENTROPY-BTU/LB-F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.5185	1.5192	1.5276	1.5358	1.5437	1.5512
30.00	1.4884	1.4891	1.4976	1.5057	1.5138	1.5212
50.00	1.4611	1.4619	1.4703	1.4784	1.4864	1.4920
100.00	1.3966	1.4055	1.4141	1.4222	1.4307	1.4378
200.00	1.3546	1.3636	1.3723	1.3805	1.3886	1.3963
300.00	1.3337	1.3428	1.3516	1.3599	1.3681	1.3759
400.00	1.3128	1.3220	1.3309	1.3394	1.3476	1.3555
500.00	1.2918	1.3012	1.3102	1.3188	1.3271	1.3350
1000.00	1.2352	1.2453	1.2549	1.2640	1.2727	1.2810
2000.00	1.182	1.194	1.205	1.215	1.224	1.233
3000.00	1.142	1.155	1.167	1.179	1.188	1.198
4000.00	1.119	1.132	1.145	1.155	1.164	1.174
5000.00	1.098	1.111	1.124	1.136	1.147	1.157

OXYGEN
THERMAL CONDUCTIVITY-BTU/SEC-FT-F
(MULTIPLY TABLE ENTRY BY .0000010)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3.918	4.009	4.214	4.353	4.483	4.611
30.00	3.93	4.09	4.22	4.36	4.49	4.62
50.00	3.94	4.09	4.23	4.37	4.50	4.63
100.00	3.97	4.12	4.26	4.40	4.53	4.65
200.00	4.03	4.17	4.32	4.45	4.58	4.70
300.00	4.09	4.23	4.37	4.50	4.63	4.75
400.00	4.15	4.29	4.43	4.56	4.68	4.80
500.00	4.21	4.35	4.48	4.61	4.73	4.85
1000.00	4.54	4.66	4.79	4.89	4.99	5.10
2000.00	5.39	5.42	5.48	5.55	5.60	5.67
3000.00	6.35	6.29	6.26	6.25	6.26	6.27
4000.00	7.37	7.23	7.14	7.08	7.03	6.98
5000.00	8.37	8.13	8.00	7.87	7.74	7.70

OXYGEN
CP-BTU/LB-F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.2189	.2192	.2195	.2199	.2204	.2208
30.00	.2194	.2197	.2199	.2203	.2207	.2212
50.00	.2201	.2203	.2205	.2208	.2212	.2216
100.00	.2217	.2219	.2221	.2224	.2227	.2230
200.00	.2251	.2248	.2246	.2246	.2246	.2247
300.00	.2287	.2280	.2274	.2271	.2269	.2267
400.00	.2322	.2311	.2303	.2297	.2293	.2290
500.00	.2357	.2342	.2331	.2323	.2316	.2312
1000.00	.2541	.2546	.2476	.2454	.2434	.2419
2000.00	.29	.29	.29	.27	.27	.26
3000.00	.33	.31	.30	.29	.29	.28
4000.00	.33	.31	.31	.30	.29	.29
5000.00						

OXYGEN
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.716	.717	.718	.718	.717	.717
30.00	.72	.71	.71	.71	.71	.71
50.00	.72	.71	.71	.71	.71	.71
100.00	.72	.72	.71	.71	.71	.71
200.00	.73	.72	.72	.72	.72	.71
300.00	.74	.73	.72	.72	.72	.72
400.00	.74	.74	.73	.73	.72	.72
500.00	.75	.74	.74	.73	.73	.73
1000.00	.78	.77	.76	.75	.75	.75
2000.00	.84	.85	.83	.81	.80	.79
3000.00	.97	.92	.89	.86	.86	.86
4000.00	.98	.93	.91	.88	.88	.85
5000.00						

OXYGEN
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.399	1.398	1.397	1.395	1.394	1.393
30.00	1.402	1.400	1.399	1.397	1.396	1.395
50.00	1.405	1.403	1.402	1.400	1.398	1.397
100.00	1.414	1.411	1.409	1.407	1.406	1.402
200.00	1.434	1.429	1.426	1.424	1.417	1.414
300.00	1.451	1.445	1.439	1.436	1.429	1.425
400.00	1.470	1.462	1.454	1.450	1.442	1.438
500.00	1.489	1.478	1.470	1.466	1.456	1.450
1000.00	1.509	1.506	1.500	1.501	1.510	1.500
2000.00	1.7	1.7	1.6	1.6	1.6	1.6
3000.00	1.6	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

OXYGEN
VISCOSITY-LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.201	1.327	1.364	1.404	1.440	1.482
30.00	1.25	1.32	1.37	1.41	1.44	1.48
50.00	1.28	1.33	1.37	1.41	1.44	1.48
100.00	1.29	1.33	1.37	1.41	1.45	1.49
200.00	1.36	1.34	1.38	1.42	1.46	1.50
300.00	1.31	1.35	1.39	1.43	1.46	1.51
400.00	1.32	1.36	1.40	1.44	1.47	1.51
500.00	1.34	1.37	1.41	1.45	1.48	1.52
1000.00	1.40	1.43	1.46	1.50	1.54	1.57
2000.00	1.63	1.63	1.65	1.66	1.66	1.72
3000.00	1.80	1.86	1.85	1.85	1.80	1.80
4000.00	2.17	2.12	2.09	2.07	2.00	2.07
5000.00	2.47	2.39	2.33	2.30	2.27	2.26

OXYGEN
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1031	1051	1071	1092	1111	1130
30.00	1031	1051	1071	1092	1111	1130
50.00	1031	1051	1072	1092	1112	1131
100.00	1031	1051	1072	1093	1113	1132
200.00	1031	1052	1073	1094	1114	1134
300.00	1031	1053	1075	1096	1116	1136
400.00	1031	1053	1075	1097	1116	1136
500.00	1031	1054	1077	1099	1118	1138
1000.00	1040	1064	1088	1111	1133	1155
2000.00	1062	1086	1110	1131	1152	1205
3000.00						
4000.00						
5000.00						

T-15

DENSITY

HELIUM

HELIUM DENSITY LBS/CU FT

DEPTH FT	PRESSURE		TEMPERATURES OF										
	PSIA	ATM	30	40	50	60	70	80	90	100	110	120	130
000	378.25	25.19	.27433	.27204	.26754	.26207	.25750	.25298	.24834	.24399	.23973	.23565	.23170
010	376.09	25.50	.27013	.27007	.27473	.26559	.26065	.25588	.25129	.24680	.24254	.23846	.23455
020	374.16	25.80	.26602	.27030	.27340	.26874	.26370	.25894	.25423	.24975	.24547	.24126	.23720
030	363.50	26.10	.26222	.26253	.27107	.27101	.26674	.26199	.25716	.25260	.24827	.24403	.23995
040	360.63	26.40	.25851	.26576	.26926	.27444	.26980	.26497	.26007	.25540	.25087	.24643	.24210
050	352.67	26.71	.25480	.26099	.26360	.27003	.27205	.26707	.26200	.25703	.25233	.24783	.24345
060	346.91	27.01	.25110	.25922	.26187	.26851	.27080	.26550	.26020	.25530	.25070	.24620	.24180
070	341.35	27.31	.24740	.25554	.25820	.26490	.26730	.26180	.25630	.25150	.24690	.24240	.23800
080	335.99	27.61	.24370	.25187	.25450	.26120	.26370	.25800	.25250	.24770	.24310	.23860	.23420
090	330.82	27.92	.24000	.24817	.25070	.25740	.26000	.25420	.24870	.24390	.23930	.23480	.23040
100	325.84	28.22	.23630	.24447	.24700	.25370	.25630	.25050	.24500	.24020	.23560	.23110	.22670
110	321.05	28.52	.23260	.24077	.24330	.25000	.25260	.24680	.24130	.23650	.23190	.22740	.22300
120	316.45	28.82	.22890	.23707	.23960	.24630	.24890	.24310	.23760	.23280	.22820	.22370	.21930
130	312.04	29.12	.22520	.23337	.23590	.24260	.24520	.23940	.23390	.22910	.22450	.22000	.21560
140	307.81	29.43	.22150	.22967	.23220	.23890	.24150	.23570	.23020	.22540	.22080	.21630	.21190
150	303.75	29.73	.21780	.22597	.22850	.23520	.23780	.23200	.22650	.22170	.21710	.21260	.20820
160	299.86	30.03	.21410	.22227	.22480	.23150	.23410	.22830	.22280	.21800	.21340	.20890	.20450
170	296.13	30.33	.21040	.21857	.22110	.22780	.23040	.22460	.21910	.21430	.20970	.20520	.20080
180	292.56	30.63	.20670	.21487	.21740	.22410	.22670	.22090	.21540	.21060	.20600	.20150	.19710
190	289.14	30.93	.20300	.21117	.21370	.22040	.22300	.21720	.21170	.20690	.20230	.19780	.19340
200	285.87	31.23	.19930	.20747	.21000	.21670	.21930	.21350	.20800	.20320	.19860	.19410	.18970
210	282.74	31.53	.19560	.20377	.20630	.21300	.21560	.20980	.20430	.19950	.19490	.19040	.18600
220	279.75	31.83	.19190	.20007	.20260	.20930	.21190	.20610	.20060	.19580	.19120	.18670	.18230
230	276.90	32.13	.18820	.19637	.19890	.20560	.20820	.20240	.19690	.19210	.18750	.18300	.17860
240	274.19	32.43	.18450	.19267	.19520	.20190	.20450	.19870	.19320	.18840	.18380	.17930	.17490
250	271.62	32.73	.18080	.18897	.19150	.19820	.20080	.19500	.18950	.18470	.18010	.17560	.17120
260	269.19	33.03	.17710	.18527	.18780	.19450	.19710	.19130	.18580	.18100	.17640	.17190	.16750
270	266.90	33.33	.17340	.18157	.18410	.19080	.19340	.18760	.18210	.17730	.17270	.16820	.16380
280	264.74	33.63	.16970	.17787	.18040	.18710	.18970	.18390	.17840	.17360	.16900	.16450	.16010
290	262.71	33.93	.16600	.17417	.17670	.18340	.18600	.18020	.17470	.16990	.16530	.16080	.15640
300	260.81	34.23	.16230	.17047	.17300	.17970	.18230	.17650	.17100	.16620	.16160	.15710	.15270
310	259.04	34.53	.15860	.16677	.16930	.17600	.17860	.17280	.16730	.16250	.15790	.15340	.14900
320	257.40	34.83	.15490	.16307	.16560	.17230	.17490	.16910	.16360	.15880	.15420	.14970	.14530
330	255.89	35.13	.15120	.15937	.16190	.16860	.17120	.16540	.16000	.15520	.15060	.14610	.14170
340	254.50	35.43	.14750	.15567	.15820	.16490	.16750	.16170	.15620	.15140	.14680	.14230	.13790
350	253.23	35.73	.14380	.15197	.15450	.16120	.16380	.15800	.15250	.14770	.14310	.13860	.13420
360	252.08	36.03	.14010	.14827	.15080	.15750	.16010	.15430	.14880	.14400	.13940	.13490	.13050
370	251.05	36.33	.13640	.14457	.14710	.15380	.15640	.15060	.14510	.14030	.13570	.13120	.12680
380	250.14	36.63	.13270	.14087	.14340	.15010	.15270	.14690	.14140	.13660	.13200	.12750	.12310
390	249.34	36.93	.12900	.13717	.13970	.14640	.14900	.14320	.13770	.13290	.12830	.12380	.11940
400	248.64	37.23	.12530	.13347	.13600	.14270	.14530	.13950	.13400	.12920	.12460	.12010	.11570
410	248.04	37.53	.12160	.12977	.13230	.13900	.14160	.13580	.13030	.12550	.12090	.11640	.11200
420	247.54	37.83	.11790	.12607	.12860	.13530	.13790	.13210	.12660	.12180	.11720	.11270	.10830
430	247.14	38.13	.11420	.12237	.12490	.13160	.13420	.12840	.12290	.11810	.11350	.10900	.10460
440	246.84	38.43	.11050	.11867	.12120	.12790	.13050	.12470	.11920	.11440	.10980	.10530	.10090
450	246.64	38.73	.10680	.11497	.11750	.12420	.12680	.12100	.11550	.11070	.10610	.10160	.9720
460	246.54	39.03	.10310	.11127	.11380	.12050	.12310	.11730	.11180	.10700	.10240	.9790	.9350
470	246.54	39.33	.9940	.10757	.11010	.11680	.11940	.11360	.10810	.10330	.9880	.9430	.8990
480	246.64	39.63	.9570	.10387	.10640	.11310	.11570	.10990	.10440	.9960	.9510	.9060	.8620
490	246.84	39.93	.9200	.10017	.10270	.10940	.11200	.10620	.10070	.9590	.9140	.8690	.8250
500	247.14	40.23	.8830	.96487	.98740	.10510	.10770	.10190	.9640	.9160	.8710	.8260	.7820

HELIUM

HELIUM SPECIFIC VOLUME, CU/IL FT/LB

Table with columns for DEPTH FT, PRESSURE PSIA and ATM, TEMPERATURES °F (30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130), and specific volume values. The table is organized in a grid with depth on the y-axis and temperature on the x-axis.

HELIUM

T-18

HELIUM
ENTHALPY, BTU/LB A

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	619.39	644.21	669.03	693.86	718.68	743.50
30.00	619.53	644.35	669.17	694.00	718.82	743.64
50.00	619.71	644.53	669.36	694.18	719.00	743.82
100.00	620.16	644.94	669.81	694.64	719.45	744.26
200.00	621.04	645.48	670.71	695.56	720.37	745.20
300.00	621.92	646.06	671.60	696.43	721.27	746.10
400.00	622.80	646.64	672.49	697.33	722.16	747.00
500.00	623.67	647.23	673.37	698.26	723.06	747.90
1000.00	628.02	652.49	677.71	702.63	727.50	752.36
2000.00	636.49	661.66	686.32	711.22	736.12	761.03
3000.00	644.80	669.77	694.72	719.66	744.59	769.53
4000.00	652.85	677.85	702.84	727.86	752.81	777.76
5000.00	660.70	685.73	710.75	735.77	760.79	785.80

HELIUM
CV, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.740	.740	.740	.740	.740	.740
30.00	.740	.740	.740	.740	.740	.740
50.00	.740	.740	.740	.740	.740	.740
100.00	.740	.740	.740	.740	.740	.740
200.00	.740	.740	.740	.740	.740	.740
300.00	.740	.740	.740	.740	.740	.740
400.00	.740	.740	.740	.740	.740	.740
500.00	.740	.740	.740	.740	.740	.740
1000.00	.740	.740	.740	.740	.740	.740
2000.00	.740	.740	.740	.740	.740	.740
3000.00	.740	.740	.740	.740	.740	.740
4000.00	.740	.740	.740	.740	.740	.740
5000.00	.740	.740	.740	.740	.740	.740

HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	6.6002	6.7300	6.7777	6.8230	6.8681	6.9109
30.00	6.5500	6.5907	6.6475	6.6930	6.7379	6.7807
50.00	6.3798	6.4295	6.4773	6.5233	6.5677	6.6105
100.00	5.9543	6.0040	6.0518	6.0976	6.1422	6.1850
200.00	5.4610	5.5113	5.5592	5.6052	5.6496	5.6925
300.00	5.3274	5.3771	5.4250	5.4710	5.5154	5.5583
400.00	5.1932	5.2429	5.2908	5.3368	5.3812	5.4241
500.00	5.0590	5.1087	5.1566	5.2026	5.2470	5.2899
1000.00	4.6166	4.6664	4.7143	4.7604	4.8049	4.8478
2000.00	4.2701	4.3200	4.3679	4.4141	4.4586	4.5016
3000.00	4.0394	4.0893	4.1374	4.1830	4.2282	4.2712
4000.00	3.8950	3.9450	3.9932	4.0395	4.0841	4.1272
5000.00	3.7826	3.8327	3.8809	3.9272	3.9719	4.0151

HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.26	2.32	2.38	2.43	2.49	2.54
30.00	2.27	2.32	2.38	2.43	2.49	2.54
50.00	2.27	2.32	2.38	2.44	2.49	2.54
100.00	2.28	2.33	2.39	2.44	2.49	2.55
200.00	2.29	2.34	2.40	2.45	2.51	2.56
300.00	2.30	2.36	2.41	2.47	2.52	2.57
400.00	2.31	2.37	2.43	2.48	2.53	2.59
500.00	2.33	2.38	2.44	2.49	2.55	2.60
1000.00	2.39	2.45	2.50	2.55	2.61	2.66
2000.00	2.43	2.49	2.55	2.60	2.66	2.71

HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.2420	1.2420	1.2420	1.2420	1.2420	1.2420
30.00	1.2421	1.2421	1.2421	1.2421	1.2421	1.2421
50.00	1.2421	1.2421	1.2421	1.2421	1.2421	1.2421
100.00	1.2423	1.2423	1.2422	1.2422	1.2422	1.2422
200.00	1.2426	1.2425	1.2425	1.2425	1.2424	1.2424
300.00	1.2426	1.2424	1.2423	1.2423	1.2423	1.2423
400.00	1.2431	1.2424	1.2423	1.2423	1.2424	1.2426
500.00	1.2436	1.2433	1.2432	1.2431	1.2431	1.2430
1000.00	1.2466	1.2466	1.2466	1.2463	1.2461	1.2460
2000.00	1.2467	1.2466	1.2464	1.2461	1.2459	1.2458
3000.00	1.2491	1.2491	1.2485	1.2482	1.2480	1.2477
4000.00	1.2509	1.2506	1.2503	1.2500	1.2497	1.2494
5000.00	1.2527	1.2524	1.2521	1.2518	1.2515	1.2512

HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.69	.69	.69	.69	.69	.69
30.00	.69	.69	.69	.69	.69	.69
50.00	.69	.69	.69	.69	.69	.69
100.00	.69	.69	.69	.69	.69	.69
200.00	.69	.69	.69	.69	.69	.69
300.00	.69	.69	.69	.69	.69	.69
400.00	.67	.67	.66	.66	.66	.66
500.00	.67	.67	.67	.67	.66	.66
1000.00	.65	.65	.65	.65	.65	.65
2000.00	.64	.64	.64	.64	.64	.64

HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.665	1.665	1.665	1.665	1.665	1.665
30.00	1.665	1.665	1.665	1.665	1.665	1.665
50.00	1.665	1.665	1.665	1.665	1.665	1.665
100.00	1.665	1.665	1.665	1.665	1.665	1.665
200.00	1.666	1.666	1.666	1.666	1.666	1.666
300.00	1.666	1.666	1.666	1.666	1.666	1.666
400.00	1.667	1.667	1.667	1.667	1.667	1.667
500.00	1.667	1.667	1.667	1.667	1.667	1.667
1000.00	1.669	1.669	1.669	1.669	1.669	1.669
2000.00	1.671	1.671	1.671	1.670	1.670	1.670
3000.00	1.674	1.674	1.674	1.673	1.673	1.672
4000.00	1.676	1.676	1.676	1.675	1.675	1.674
5000.00	1.678	1.678	1.678	1.677	1.677	1.676

HELIUM
VISCOSITY • LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.257	1.290	1.323	1.357	1.389	1.422
30.00	1.257	1.290	1.323	1.357	1.389	1.422
50.00	1.257	1.290	1.323	1.357	1.389	1.422
100.00	1.257	1.290	1.323	1.357	1.389	1.422
200.00	1.257	1.290	1.323	1.357	1.389	1.422
300.00	1.257	1.290	1.323	1.357	1.389	1.422
400.00	1.257	1.290	1.323	1.357	1.389	1.422
500.00	1.257	1.290	1.323	1.357	1.389	1.422
1000.00	1.257	1.290	1.323	1.357	1.389	1.422
2000.00	1.257	1.290	1.323	1.357	1.389	1.422
3000.00	1.257	1.290	1.323	1.357	1.389	1.422
4000.00	1.257	1.290	1.323	1.357	1.389	1.422
5000.00	1.257	1.290	1.323	1.357	1.389	1.422

HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3162	3207	3242	3302	3431	3491
30.00	3164	3204	3244	3303	3433	3493
50.00	3166	3207	3246	3305	3435	3495
100.00	3173	3214	3243	3371	3440	3500
200.00	3186	3230	3306	3382	3450	3510
300.00	3199	3245	3314	3393	3460	3520
400.00	3212	3260	3332	3408	3470	3530
500.00	3225	3274	3345	3415	3480	3540
1000.00	3240	3350	3410	3470	3530	3590
2000.00	3340	3450	3510	3570	3630	3690
3000.00	3425	3574	3645	3695	3755	3815
4000.00	3460	3700	3760	3820	3880	3940
5000.00	3760	3800	3860	3920	3977	4032

DENSITY

T-20

NITROGEN

NITROGEN DENSITY, LBS/CUBIC FT

Table with columns: DEPTH FT, PRESSURE PSIA, ATM, TEMPERATURE 30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130. Rows represent depth intervals from 0 to 790 feet.

NITROGEN

T-24

NITROGEN
ENTHALPY (BTU/LB)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	226.74	231.73	236.67	241.60	246.55	251.44
30.00	226.68	231.63	236.57	241.51	246.46	251.31
50.00	226.53	231.49	236.44	241.37	246.33	251.21
100.00	226.15	231.14	236.12	241.10	246.07	251.05
200.00	225.40	230.44	235.48	240.51	245.53	250.54
300.00	224.84	229.75	234.84	239.94	244.99	250.05
400.00	223.84	229.06	234.21	239.33	244.45	249.55
500.00	223.13	228.47	233.57	238.75	243.92	249.06
1000.00	219.47	225.06	230.55	236.00	241.42	246.77
2000.00	213.14	219.74	225.38	231.31	237.15	242.84
3000.00	209.15	215.81	221.93	227.11	234.21	240.14
4000.00	206.84	213.71	219.80	225.15	232.41	238.55
5000.00	205.46	212.14	218.71	224.14	231.50	237.73

NITROGEN
ENTHALPY (KJ/KG)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.176	1.174	1.176	1.176	1.176	1.176
30.00	1.174	1.174	1.176	1.176	1.176	1.176
50.00	1.174	1.174	1.176	1.176	1.176	1.176
100.00	1.174	1.174	1.176	1.176	1.176	1.176
200.00	1.174	1.174	1.176	1.176	1.176	1.176
300.00	1.177	1.174	1.176	1.176	1.176	1.176
400.00	1.177	1.177	1.177	1.177	1.177	1.177
500.00	1.177	1.177	1.177	1.177	1.177	1.177
1000.00	1.174	1.174	1.174	1.174	1.174	1.174
2000.00	1.143	1.142	1.142	1.142	1.141	1.141
3000.00	1.126	1.124	1.124	1.124	1.123	1.123
4000.00	1.118	1.117	1.118	1.118	1.118	1.118
5000.00	1.141	1.140	1.141	1.141	1.141	1.141

NITROGEN
ENTHALPY (BTU/LB F)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.6075	1.6174	1.6268	1.6359	1.6447	1.6532
30.00	1.5775	1.5874	1.5969	1.6060	1.6149	1.6234
50.00	1.5384	1.5483	1.5578	1.5669	1.5758	1.5843
100.00	1.4744	1.4843	1.4939	1.5031	1.5120	1.5206
200.00	1.4261	1.4362	1.4459	1.4551	1.4641	1.4727
300.00	1.3953	1.4055	1.4153	1.4246	1.4337	1.4424
400.00	1.3691	1.3795	1.3894	1.3988	1.4080	1.4168
500.00	1.3527	1.3631	1.3731	1.3827	1.3919	1.4008
1000.00	1.2956	1.3064	1.3174	1.3276	1.3371	1.3463
2000.00	1.2341	1.2444	1.2548	1.2649	1.2744	1.2843
3000.00	1.1967	1.2066	1.2210	1.2332	1.2441	1.2543
4000.00	1.1701	1.1833	1.1958	1.2075	1.2187	1.2292
5000.00	1.1444	1.1624	1.1754	1.1872	1.1986	1.2093

NITROGEN
THERMAL CONDUCTIVITY (BTU/SEC FT F)
(MULTIPLY TABLE ENTRY BY .0000010)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3.834	3.947	4.100	4.227	4.344	4.470
30.00	3.84	3.94	4.11	4.24	4.36	4.48
50.00	3.86	3.99	4.12	4.25	4.37	4.49
100.00	3.89	4.02	4.15	4.27	4.39	4.51
200.00	3.95	4.08	4.21	4.33	4.45	4.56
300.00	4.01	4.14	4.26	4.38	4.50	4.62
400.00	4.08	4.20	4.32	4.44	4.55	4.67
500.00	4.14	4.26	4.38	4.49	4.60	4.72
1000.00	4.48	4.54	4.68	4.78	4.88	4.98
2000.00	5.28	5.32	5.37	5.44	5.52	5.58
3000.00	6.17	6.18	6.14	6.17	6.23	6.23
4000.00	7.07	6.99	6.94	6.91	6.90	6.88
5000.00	7.93	7.81	7.72	7.64	7.58	7.54

NITROGEN
CUMULATIVE

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.464	2.464	2.470	2.471	2.472	2.473
30.00	2.473	2.473	2.474	2.475	2.475	2.476
50.00	2.479	2.479	2.479	2.479	2.479	2.480
100.00	2.493	2.492	2.491	2.491	2.490	2.490
200.00	2.524	2.521	2.517	2.514	2.512	2.511
300.00	2.557	2.550	2.544	2.539	2.534	2.530
400.00	2.590	2.580	2.571	2.563	2.554	2.550
500.00	2.624	2.610	2.598	2.587	2.577	2.569
1000.00	2.805	2.771	2.741	2.713	2.681	2.671
2000.00	3.114	3.051	2.995	2.945	2.895	2.854
3000.00	3.266	3.194	3.130	3.072	3.017	2.971
4000.00	3.350	3.279	3.214	3.155	3.100	3.052
5000.00	3.377	3.312	3.251	3.190	3.143	3.094

NITROGEN
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.72	1.72	1.71	1.71	1.71	1.71
30.00	1.72	1.72	1.71	1.71	1.71	1.71
50.00	1.72	1.72	1.71	1.71	1.71	1.71
100.00	1.72	1.72	1.71	1.71	1.71	1.71
200.00	1.73	1.72	1.72	1.72	1.72	1.72
300.00	1.73	1.73	1.72	1.72	1.72	1.72
400.00	1.73	1.73	1.72	1.72	1.72	1.72
500.00	1.74	1.73	1.73	1.72	1.72	1.72
1000.00	1.77	1.76	1.75	1.74	1.74	1.74
2000.00	1.82	1.80	1.74	1.74	1.74	1.74
3000.00	1.85	1.83	1.82	1.81	1.79	1.74
4000.00	1.86	1.84	1.81	1.81	1.80	1.79
5000.00	1.87	1.85	1.84	1.82	1.81	1.80

NITROGEN
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.40	1.40	1.40	1.40	1.40	1.40
30.00	1.40	1.40	1.40	1.40	1.40	1.41
50.00	1.41	1.41	1.41	1.41	1.41	1.41
100.00	1.41	1.41	1.41	1.41	1.41	1.41
200.00	1.43	1.43	1.43	1.43	1.43	1.43
300.00	1.45	1.45	1.45	1.45	1.45	1.45
400.00	1.47	1.47	1.47	1.47	1.47	1.47
500.00	1.48	1.47	1.47	1.47	1.46	1.45
1000.00	1.50	1.50	1.50	1.52	1.51	1.50
2000.00	1.58	1.55	1.63	1.61	1.59	1.57
3000.00	1.75	1.73	1.70	1.68	1.65	1.63
4000.00	1.78	1.75	1.72	1.70	1.66	1.66
5000.00	1.77	1.75	1.73	1.71	1.69	1.67

NITROGEN
VISCOSITY*LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.114	1.140	1.164	1.219	1.255	1.289
30.00	1.12	1.15	1.19	1.22	1.26	1.29
50.00	1.12	1.15	1.19	1.22	1.26	1.29
100.00	1.12	1.16	1.19	1.23	1.26	1.30
200.00	1.13	1.16	1.20	1.23	1.27	1.30
300.00	1.14	1.17	1.21	1.24	1.28	1.31
400.00	1.15	1.18	1.22	1.25	1.28	1.32
500.00	1.16	1.19	1.22	1.26	1.29	1.33
1000.00	1.22	1.25	1.28	1.31	1.34	1.37
2000.00	1.40	1.41	1.43	1.45	1.47	1.49
3000.00	1.60	1.60	1.60	1.61	1.62	1.63
4000.00	1.82	1.80	1.79	1.79	1.80	1.80
5000.00	2.04	2.00	1.98	1.97	1.96	1.95

NITROGEN
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1103	1125	1147	1166	1189	1212
30.00	1104	1126	1148	1169	1190	1213
50.00	1105	1127	1149	1170	1191	1214
100.00	1106	1129	1151	1172	1194	1217
200.00	1110	1133	1155	1176	1199	1222
300.00	1114	1137	1159	1181	1204	1227
400.00	1117	1141	1163	1185	1209	1232
500.00	1121	1145	1168	1190	1214	1237
1000.00	1140	1171	1195	1218	1243	1266
2000.00	1212	1239	1265	1289	1311	1333
3000.00	1285	1312	1340	1363	1385	1405
4000.00	1376	1401	1425	1448	1470	1491
5000.00	1466	1490	1515	1536	1556	1576

DENSITY

CARBON DIOXIDE

CARBON DIOXIDE DENSITY, LBS/CUBIC FT

DEPTH FT	PRESSURE		TEMPERATURE, °F										
	PSIA	ATM	30	40	50	60	70	80	90	100	110	120	130
0	14.70	1.00	.12393	.12100	.11807	.11663	.11439	.11226	.11016	.10810	.10626	.10438	.10259
10	19.14	1.30	.15175	.14862	.14523	.14217	.13923	.13640	.13369	.13100	.12844	.12591	.12347
20	23.50	1.60	.19976	.19500	.19053	.18703	.18368	.18047	.17730	.17420	.17116	.16818	.16524
30	28.03	1.91	.23788	.23292	.22817	.22362	.21925	.21505	.21102	.20714	.20340	.19978	.19628
40	32.67	2.21	.27619	.27039	.26475	.25923	.25481	.25051	.24633	.24226	.23831	.23447	.23073
50	36.92	2.51	.31467	.30802	.30144	.29557	.28973	.28402	.27847	.27307	.26781	.26269	.25769
60	41.30	2.81	.35332	.34580	.33841	.33116	.32415	.31730	.31060	.30407	.29771	.29151	.28546
70	45.81	3.12	.39216	.38376	.37571	.36780	.35968	.35135	.34320	.33531	.32758	.31999	.31254
80	50.25	3.42	.43116	.42183	.41285	.40410	.39558	.38728	.37920	.37134	.36369	.35624	.34898
90	54.70	3.72	.47031	.46009	.45016	.44051	.43112	.42198	.41308	.40442	.39599	.38778	.37978
100	59.14	4.02	.50965	.49850	.48767	.47716	.46697	.45700	.44734	.43799	.42886	.41994	.41124
110	63.58	4.33	.54919	.53708	.52526	.51373	.50248	.49150	.48088	.47061	.46068	.45108	.44180
120	68.03	4.63	.58889	.57583	.56309	.55065	.53851	.52666	.51509	.50380	.49287	.48229	.47204
130	72.47	4.93	.62879	.61476	.60103	.58759	.57444	.56158	.54909	.53696	.52518	.51374	.50263
140	76.92	5.23	.66887	.65382	.63901	.62443	.61008	.59596	.58207	.56850	.55524	.54230	.52967
150	81.36	5.54	.70914	.69308	.67727	.66169	.64634	.63121	.61630	.60160	.58711	.57283	.55885
160	85.81	5.84	.74961	.73258	.71579	.69923	.68290	.66679	.65089	.63520	.61981	.60472	.58993
170	90.25	6.14	.79027	.77218	.75434	.73673	.71934	.70216	.68519	.66842	.65185	.63557	.61958
180	94.70	6.44	.83112	.81196	.79302	.77429	.75577	.73746	.71936	.70146	.68376	.66625	.64894
190	99.14	6.75	.87218	.85193	.83205	.81243	.79308	.77391	.75492	.73612	.71750	.69907	.68083
200	103.58	7.05	.91344	.89198	.87164	.85140	.83127	.81134	.79160	.77205	.75269	.73351	.71451
210	108.03	7.35	.95490	.93229	.91091	.88962	.86843	.84734	.82644	.80572	.78518	.76481	.74461
220	112.47	7.65	.99657	.97290	.94943	.92603	.90272	.87950	.85637	.83342	.81064	.78803	.76559
230	116.92	7.96	1.03845	1.01350	.98822	.96260	.93757	.91214	.88640	.86034	.83445	.80873	.78318
240	121.36	8.26	1.08056	1.05439	1.02794	1.00121	.97421	.94694	.91940	.89158	.86347	.83547	.80767
250	125.81	8.56	1.12287	1.09547	1.06787	1.04007	1.01207	.98387	.95537	.92657	.89747	.86847	.83967
260	130.25	8.86	1.16540	1.13675	1.10795	1.07890	1.04960	1.01995	.98995	.95960	.92890	.89835	.86805
270	134.69	9.17	1.20816	1.17823	1.14813	1.11777	1.08716	1.05630	1.02510	.99355	.96165	.92940	.89740
280	139.14	9.47	1.25111	1.21996	1.18863	1.15703	1.12517	1.09306	1.06070	1.02809	.99524	.96214	.92940
290	143.58	9.77	1.29424	1.26180	1.22917	1.19627	1.16311	1.12970	1.09604	1.06214	1.02800	.99372	.95940
300	148.03	10.07	1.33756	1.30393	1.27017	1.23617	1.20193	1.16745	1.13273	1.09777	1.06257	1.02713	.99165

CARBON DIOXIDE

CARBON DIOXIDE SPECIFIC VOLUME-CUBIC FT/LB

DEPTH FT	PRESSURE		TEMPERATURE °F										
	PSIA	ATM	30	40	50	60	70	80	90	100	110	120	130
0	14.70	1.00	6.0689	6.2374	6.4057	6.5730	6.7419	6.9107	7.0775	7.2452	7.4124	7.5803	7.7475
10	19.70	1.38	6.1423	6.3122	6.4809	6.6491	6.8171	6.9845	7.1510	7.3170	7.4824	7.6474	7.8119
20	24.70	1.76	5.9966	6.1685	6.3392	6.5087	6.6770	6.8441	7.0100	7.1746	7.3378	7.4997	7.6603
30	29.70	2.14	6.2837	6.4573	6.6297	6.8007	6.9701	7.1379	7.3041	7.4688	7.6320	7.7937	7.9540
40	34.70	2.51	3.1779	3.2605	3.3430	3.4254	3.5075	3.5893	3.6708	3.7519	3.8326	3.9129	3.9928
50	39.70	2.89	2.8933	2.9769	3.0602	3.1432	3.2258	3.3080	3.3898	3.4711	3.5519	3.6322	3.7121
60	44.70	3.27	2.5501	2.6347	2.7189	2.8027	2.8861	2.9690	3.0514	3.1333	3.2147	3.2956	3.3760
70	49.70	3.64	2.3194	2.4040	2.4881	2.5717	2.6548	2.7374	2.8195	2.9011	2.9822	3.0628	3.1429
80	54.70	4.02	2.1203	2.2049	2.2889	2.3724	2.4554	2.5379	2.6199	2.7014	2.7824	2.8629	2.9429
100	59.70	4.40	1.9621	2.0467	2.1307	2.2141	2.2970	2.3794	2.4613	2.5427	2.6236	2.7040	2.7839
110	64.70	4.78	1.8289	1.9135	1.9974	2.0807	2.1635	2.2458	2.3276	2.4089	2.4897	2.5700	2.6498
120	69.70	5.16	1.7101	1.7947	1.8786	1.9619	2.0446	2.1268	2.2085	2.2897	2.3704	2.4506	2.5303
130	74.70	5.54	1.6040	1.6886	1.7724	1.8556	1.9382	2.0203	2.1019	2.1830	2.2636	2.3437	2.4233
140	79.70	5.92	1.5091	1.5937	1.6774	1.7605	1.8430	1.9250	2.0065	2.0875	2.1680	2.2480	2.3275
150	84.70	6.30	1.4142	1.4988	1.5824	1.6654	1.7478	1.8297	1.9111	1.9920	2.0724	2.1523	2.2317
160	89.70	6.68	1.3193	1.4039	1.4874	1.5703	1.6526	1.7343	1.8154	1.8960	1.9761	2.0557	2.1348
170	94.70	7.06	1.2244	1.3090	1.3924	1.4752	1.5574	1.6390	1.7201	1.8007	1.8808	1.9604	2.0395
180	99.70	7.44	1.1295	1.2141	1.2974	1.3801	1.4622	1.5437	1.6246	1.7050	1.7849	1.8643	1.9432
190	104.70	7.82	1.0346	1.1192	1.2024	1.2849	1.3667	1.4478	1.5282	1.6081	1.6875	1.7664	1.8448
200	109.70	8.20	0.9397	1.0243	1.1074	1.1897	1.2712	1.3519	1.4319	1.5113	1.5901	1.6684	1.7461
210	114.70	8.58	0.8448	0.9294	1.0124	1.0945	1.1758	1.2563	1.3360	1.4151	1.4936	1.5715	1.6488
220	119.70	8.96	0.7499	0.8345	0.9174	0.9991	1.0799	1.1598	1.2389	1.3173	1.3950	1.4721	1.5486
230	124.70	9.34	0.6550	0.7396	0.8224	0.9039	0.9844	1.0640	1.1427	1.2206	1.2977	1.3741	1.4498
240	129.70	9.72	0.5601	0.6447	0.7274	0.8087	0.8890	0.9683	1.0467	1.1242	1.2008	1.2765	1.3513
250	134.70	10.10	0.4652	0.5498	0.6324	0.7135	0.7936	0.8727	0.9508	1.0280	1.1043	1.1797	1.2542
260	139.70	10.48	0.3703	0.4549	0.5374	0.6181	0.6976	0.7759	0.8531	0.9293	1.0045	1.0787	1.1520
270	144.70	10.86	0.2754	0.3599	0.4424	0.5227	0.6015	0.6787	0.7543	0.8284	0.9010	0.9721	1.0426
280	149.70	11.24	0.1805	0.2650	0.3474	0.4273	0.5054	0.5817	0.6562	0.7289	0.7999	0.8692	0.9378
290	154.70	11.62	0.0856	0.1701	0.2524	0.3321	0.4099	0.4858	0.5598	0.6319	0.7021	0.7704	0.8378
300	159.70	12.00	0.0000	0.0845	0.1668	0.2457	0.3218	0.3960	0.4683	0.5387	0.6072	0.6738	0.7394

CARBON DIOXIDE

T-30

CARBON DIOXIDE
ENTHALPY, BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	81.67	95.45	89.67	93.76	97.87	102.05
30.00	81.14	85.16	89.22	92.34	97.49	101.70
50.00	80.44	84.52	88.63	92.00	96.99	101.23
100.00	78.69	82.92	87.17	91.44	95.74	100.07
200.00				88.33	92.95	97.53
300.00				84.94	89.99	94.05
400.00				81.55	87.02	92.10
500.00				78.15	84.06	89.51
1000.00						73.00
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
CV, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.1506	.1534	.1501	.1500	.1513	.1534
30.00	.1514	.1541	.1507	.1503	.1518	.1543
50.00	.1524	.1550	.1514	.1509	.1524	.1549
100.00	.1544	.1572	.1534	.1510	.1533	.1562
200.00				.157	.166	.164
300.00				.173	.173	.172
400.00				.180	.174	.175
500.00				.187	.183	.176
1000.00						.201
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.1471	1.1487	1.1504	1.1529	1.1713	1.1705
30.00	1.1300	1.1327	1.1400	1.1482	1.1556	1.1620
50.00	1.1076	1.1119	1.1199	1.1279	1.1350	1.1423
100.00	1.0517	1.0590	1.0681	1.0760	1.0837	1.0911
200.00				1.0454	1.0530	1.0610
300.00				1.0267	1.0359	1.0442
400.00				1.0079	1.0180	1.0260
500.00				.9892	1.0001	1.0094
1000.00						.9527
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000010)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.324	2.459	2.595	2.732	2.870	3.000
30.00	2.34	2.44	2.61	2.75	2.88	3.02
50.00	2.37	2.50	2.63	2.76	2.90	3.04
100.00	2.42	2.55	2.68	2.81	2.95	3.08
200.00				2.91	3.04	3.17
300.00				3.00	3.12	3.25
400.00				3.10	3.22	3.34
500.00				3.23	3.33	3.45
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.1973	.1997	.2023	.2040	.2073	.2096
30.00	.1997	.2019	.2042	.2065	.2094	.2111
50.00	.2020	.2047	.2067	.2087	.2107	.2120
100.00	.2107	.2117	.2126	.2142	.2155	.2171
200.00				.2124	.2104	.2080
300.00				.2064	.2093	.2029
400.00				.2004	.2047	.2070
500.00				.1944	.2071	.2070
1000.00						.1720
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.76	.77	.77	.77	.77	.77
30.00	.79	.78	.78	.77	.77	.77
50.00	.79	.79	.78	.76	.76	.76
100.00	.81	.80	.80	.79	.79	.79
200.00				.84	.83	.82
300.00				.92	.84	.86
400.00				.94	.94	.90
500.00				1.06	.99	.93
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.310	1.302	1.296	1.291	1.286	1.281
30.00	1.319	1.310	1.303	1.297	1.291	1.285
50.00	1.331	1.321	1.313	1.305	1.296	1.291
100.00	1.361	1.347	1.335	1.325	1.315	1.307
200.00				1.34	1.37	1.35
300.00				1.40	1.43	1.41
400.00				1.54	1.50	1.46
500.00				1.62	1.56	1.52
1000.00						2.68
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
VISCOSITY LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.917	.954	.991	1.027	1.066	1.104
30.00	.920	.957	.993	1.029	1.068	1.106
50.00	.924	.960	.996	1.032	1.072	1.109
100.00	.934	.969	1.005	1.040	1.079	1.117
200.00				1.057	1.095	1.132
300.00				1.075	1.111	1.147
400.00				1.095	1.131	1.165
500.00				1.120	1.153	1.185
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

CARBON DIOXIDE
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	842	857	873	887	902	916
30.00	838	854	870	885	899	914
50.00	833	850	866	881	896	911
100.00	822	840	857	873	889	905
200.00				854	873	891
300.00				832	857	876
400.00				810	840	865
500.00				789	824	851
1000.00						766
2000.00						
3000.00						
4000.00						
5000.00						

0.600 OXYGEN 0.400 HELIUM
ENTHALPY, BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	145.74	151.70	157.66	163.63	169.61	175.58
30.00	145.63	151.60	157.56	163.54	169.51	175.50
50.00	145.48	151.46	157.44	163.42	169.42	175.40
100.00	145.11	151.12	157.12	163.13	169.14	175.15
200.00	144.4	150.4	156.5	162.5	168.6	174.7
300.00	143.6	149.7	155.9	162.0	168.1	174.2
400.00	142.9	149.0	155.2	161.4	167.5	173.7
500.00	142.1	148.3	154.6	160.8	167.0	173.2
1000.00	138.3	145.0	151.5	157.9	164.4	170.7
2000.00	131.8	139.1	146.2	153.3	159.9	166.4
3000.00	125.7	133.6	141.6	149.1	156.0	167.1
4000.00	122.6	130.8	138.9	146.6	153.9	161.0
5000.00	120.3	128.1	136.0	144.6	151.9	159.5

0.600 OXYGEN 0.400 HELIUM
CV, BTU/LH F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.2019	.2022	.2025	.2029	.2034	.2037
30.00	.2020	.2022	.2025	.2029	.2034	.2037
50.00	.2020	.2023	.2026	.2030	.2035	.2038
100.00	.2022	.2025	.2027	.2031	.2036	.2039
200.00	.2025	.2028	.2030	.2034	.2038	.2041
300.00	.2029	.2031	.2032	.2036	.2040	.2044
400.00	.2032	.2034	.2035	.2039	.2042	.2046
500.00	.2035	.2037	.2038	.2041	.2044	.2048
1000.00	.2051	.2051	.2051	.2053	.2054	.2057
2000.00	.22	.22	.22	.21	.21	.21
3000.00	.23	.23	.23	.22	.22	.22
4000.00						
5000.00						

0.600 OXYGEN 0.400 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.9728	1.9846	1.9961	2.0071	2.0179	2.0281
30.00	1.9350	1.9468	1.9583	1.9693	1.9801	1.9904
50.00	1.8956	1.8975	1.9090	1.9200	1.9307	1.9411
100.00	1.8118	1.8238	1.8354	1.8464	1.8572	1.8675
200.00	1.7704	1.7825	1.7942	1.8053	1.8162	1.8265
300.00	1.7289	1.7412	1.7530	1.7642	1.7752	1.7856
400.00	1.6875	1.6999	1.7118	1.7232	1.7342	1.7446
500.00	1.6461	1.6586	1.6706	1.6821	1.6932	1.7037
1000.00	1.5598	1.5730	1.5855	1.5974	1.6089	1.6199
2000.00	1.484	1.499	1.512	1.525	1.537	1.549
3000.00	1.408	1.445	1.460	1.475	1.486	1.499
4000.00	1.397	1.412	1.428	1.441	1.453	1.465
5000.00	1.369	1.385	1.400	1.415	1.428	1.441

0.600 OXYGEN 0.400 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.76	.79	.81	.83	.85	.87
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.600 OXYGEN 0.400 HELIUM
CP, BTU/LH F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.2977	.2980	.2982	.2986	.2991	.2994
30.00	.2982	.2984	.2986	.2990	.2994	.2997
50.00	.2988	.2990	.2991	.2994	.2999	.3001
100.00	.3004	.3004	.3004	.3006	.3009	.3011
200.00	.3036	.3033	.3031	.3030	.3031	.3031
300.00	.3068	.3061	.3057	.3053	.3052	.3051
400.00	.3100	.3099	.3083	.3077	.3074	.3071
500.00	.3134	.3120	.3110	.3101	.3096	.3091
1000.00	.3306	.3272	.3244	.3223	.3204	.3190
2000.00	.37	.36	.36	.35	.35	.34
3000.00	.40	.38	.38	.37	.37	.36
4000.00	.40	.38	.38	.37	.37	.36
5000.00						

0.600 OXYGEN 0.400 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.53	.53	.53	.53	.53	.53
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.600 OXYGEN 0.400 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.474	1.474	1.473	1.472	1.471	1.470
30.00	1.476	1.476	1.475	1.473	1.472	1.471
50.00	1.479	1.478	1.477	1.475	1.474	1.473
100.00	1.486	1.484	1.482	1.480	1.479	1.477
200.00	1.499	1.496	1.493	1.490	1.487	1.485
300.00	1.513	1.507	1.504	1.499	1.496	1.493
400.00	1.526	1.519	1.515	1.509	1.505	1.501
500.00	1.540	1.532	1.526	1.519	1.514	1.509
1000.00	1.612	1.595	1.582	1.570	1.560	1.551
2000.00	1.7	1.6	1.6	1.6	1.6	1.6
3000.00	1.7	1.7	1.7	1.7	1.7	1.6
4000.00						
5000.00						

0.600 OXYGEN 0.400 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.343	1.384	1.427	1.468	1.505	1.548
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.600 OXYGEN 0.400 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1370	1374	1377	1381	1416	1440
30.00	1370	1374	1378	1382	1417	1441
50.00	1371	1375	1379	1384	1418	1442
100.00	1371	1377	1382	1387	1421	1446
200.00	1373	1381	1388	1394	1428	1453
300.00	1375	1384	1394	1411	1435	1459
400.00	1376	1388	1400	1418	1442	1466
500.00	1378	1392	1406	1425	1449	1473
1000.00	1386	1410	1435	1459	1483	1507
2000.00	1438	1448	1458	1481	1517	1553
3000.00	1490	1486	1482	1504	1522	1600
4000.00						
5000.00						

T-41

SP. VOL.
40 % O₂, 60 % He

SPECIFIC VOLUME-CUBIC FT/LB

Table with columns: DEPTH FT, PRESSURE PSIA, ATM, TEMPERATURE F (30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130), and columns for 400 OXYGEN, 400 HELIUM, AVERAGE MOLECULAR WEIGHT 15.202. The table contains a grid of numerical values for specific volume at various depths and temperatures.

40 % O₂, 60 % He

T-42

0.400 OXYGEN 0.600 HELIUM
ENTHALPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	187.32	194.92	202.54	210.17	217.80	225.43
30.00	187.23	194.84	202.46	210.10	217.74	225.37
50.00	187.11	194.73	202.37	210.01	217.65	225.30
100.00	186.81	194.46	202.12	209.78	217.44	225.11
200.00	186.2	193.9	201.6	207.7	217.0	224.8
300.00	185.6	193.3	201.1	205.6	216.6	224.4
400.00	185.0	192.8	200.6	203.5	216.2	224.1
500.00	184.4	192.2	200.1	201.4	215.8	223.7
1000.00	181.3	189.5	197.7	205.7	213.8	221.8
2000.00	176.1	184.9	193.6	202.2	210.4	218.6
3000.00	171.3	180.7	190.1	199.2	207.6	215.8
4000.00	169.2	178.9	188.3	197.6	206.4	215.2
5000.00	167.7	177.0	186.4	196.5	205.3	214.5

0.400 OXYGEN 0.600 HELIUM
CV, RTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.2456	.2499	.2502	.2505	.2510	.2513
30.00	.2496	.2499	.2502	.2505	.2510	.2513
50.00	.2497	.2500	.2503	.2506	.2511	.2514
100.00	.2499	.2501	.2504	.2507	.2512	.2515
200.00	.2502	.2504	.2507	.2510	.2514	.2517
300.00	.2505	.2507	.2509	.2512	.2515	.2519
400.00	.2508	.2510	.2512	.2515	.2517	.2521
500.00	.2511	.2513	.2514	.2517	.2519	.2523
1000.00	.2526	.2528	.2528	.2529	.2528	.2531
2000.00	.26	.26	.26	.25	.25	.25
3000.00	.27	.27	.27	.26	.26	.26
4000.00						
5000.00						

0.400 OXYGEN 0.600 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.4151	2.4304	2.4450	2.4591	2.4728	2.4859
30.00	2.3692	2.3845	2.3991	2.4132	2.4269	2.4401
50.00	2.3092	2.3245	2.3391	2.3533	2.3669	2.3802
100.00	2.2045	2.2199	2.2347	2.2488	2.2625	2.2757
200.00	2.1471	2.1626	2.1775	2.1917	2.2054	2.2187
300.00	2.0897	2.1052	2.1202	2.1345	2.1484	2.1617
400.00	2.0323	2.0479	2.0630	2.0774	2.0913	2.1047
500.00	1.9749	1.9906	2.0058	2.0203	2.0343	2.0477
1000.00	1.8573	1.8737	1.8894	1.9043	1.9186	1.9324
2000.00	1.758	1.776	1.792	1.808	1.823	1.837
3000.00	1.688	1.706	1.725	1.742	1.757	1.771
4000.00	1.646	1.665	1.683	1.698	1.713	1.729
5000.00	1.610	1.629	1.648	1.665	1.681	1.696

0.400 OXYGEN 0.600 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.05	1.04	1.11	1.14	1.17	1.20
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.400 OXYGEN 0.600 HELIUM
CP, RTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.3805	.3808	.3811	.3814	.3818	.3821
30.00	.3810	.3812	.3815	.3817	.3821	.3824
50.00	.3815	.3817	.3819	.3822	.3825	.3828
100.00	.3830	.3830	.3831	.3833	.3835	.3837
200.00	.3860	.3857	.3855	.3854	.3855	.3855
300.00	.3889	.3883	.3874	.3876	.3874	.3874
400.00	.3919	.3909	.3903	.3898	.3894	.3892
500.00	.3950	.3937	.3928	.3920	.3914	.3910
1000.00	.4108	.4074	.4051	.4032	.4015	.4002
2000.00	.44	.43	.43	.43	.43	.42
3000.00	.47	.46	.45	.44	.44	.43
4000.00	.47	.46	.46	.45	.44	.44
5000.00						

0.400 OXYGEN 0.600 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.50	.50	.50	.50	.50	.50
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.400 OXYGEN 0.600 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.524	1.524	1.523	1.523	1.521	1.520
30.00	1.526	1.525	1.524	1.524	1.522	1.521
50.00	1.528	1.527	1.526	1.525	1.523	1.522
100.00	1.533	1.531	1.530	1.529	1.527	1.525
200.00	1.543	1.540	1.538	1.536	1.534	1.532
300.00	1.553	1.548	1.546	1.543	1.540	1.538
400.00	1.563	1.557	1.554	1.550	1.547	1.544
500.00	1.573	1.566	1.562	1.557	1.554	1.550
1000.00	1.626	1.614	1.604	1.595	1.588	1.581
2000.00	1.7	1.7	1.7	1.7	1.7	1.7
3000.00	1.7	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

0.400 OXYGEN 0.600 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.369	1.410	1.453	1.491	1.531	1.573
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.400 OXYGEN 0.600 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1564	1594	1625	1655	1684	1714
30.00	1565	1596	1626	1656	1685	1714
50.00	1567	1598	1628	1658	1687	1716
100.00	1572	1602	1633	1662	1691	1720
200.00	1582	1612	1642	1671	1700	1729
300.00	1591	1621	1651	1680	1708	1737
400.00	1601	1630	1660	1689	1717	1745
500.00	1610	1640	1669	1697	1726	1754
1000.00	1658	1686	1714	1742	1769	1796
2000.00	1729	1761	1793	1825	1856	1888
3000.00	1800	1836	1872	1908	1944	1980
4000.00						
5000.00						

0.320 OXYGEN 0.680 HELIUM
ENTHALPY, BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	214.00	222.67	231.35	240.05	248.73	257.43
30.00	213.92	222.60	231.29	239.99	248.68	257.38
50.00	213.82	222.51	231.21	239.91	248.61	257.32
100.00	213.57	222.29	231.00	239.72	248.45	257.17
200.00	211.1	221.8	230.6	239.4	248.1	256.9
300.00	212.5	221.3	230.2	239.0	247.8	256.6
400.00	212.0	220.9	229.8	238.7	247.4	256.3
500.00	211.5	220.4	229.4	238.3	247.1	256.0
1000.00	208.9	218.1	227.3	236.4	245.5	254.5
2000.00	204.6	214.1	224.1	233.7	242.9	252.1
3000.00	200.5	210.9	221.3	231.3	240.8	250.0
4000.00	199.0	209.6	220.1	230.3	240.2	249.9
5000.00	194.2	208.5	218.8	229.8	239.5	249.8

0.320 OXYGEN 0.680 HELIUM
CV, RTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.2403	.2406	.2409	.2412	.2416	.2419
30.00	.2403	.2406	.2409	.2412	.2415	.2419
50.00	.2404	.2407	.2410	.2413	.2417	.2420
100.00	.2405	.2408	.2411	.2414	.2418	.2421
200.00	.2408	.2411	.2413	.2416	.2419	.2422
300.00	.2411	.2413	.2415	.2418	.2421	.2424
400.00	.2414	.2416	.2417	.2421	.2424	.2426
500.00	.2417	.2418	.2419	.2423	.2425	.2428
1000.00	.2430	.2430	.2430	.2432	.2433	.2435
2000.00	.24	.24	.24	.24	.24	.24
3000.00	.30	.30	.30	.29	.29	.29
4000.00						
5000.00						

0.320 OXYGEN 0.680 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.6922	2.7095	2.7261	2.7423	2.7578	2.7727
30.00	2.6411	2.6584	2.6750	2.6912	2.7067	2.7218
50.00	2.5742	2.5916	2.6083	2.6243	2.6400	2.6553
100.00	2.4497	2.4672	2.4840	2.5000	2.5157	2.5307
200.00	2.3820	2.3994	2.4165	2.4326	2.4483	2.4634
300.00	2.3143	2.3319	2.3489	2.3652	2.3810	2.3961
400.00	2.2466	2.2643	2.2814	2.2977	2.3136	2.3288
500.00	2.1789	2.1967	2.2139	2.2305	2.2463	2.2615
1000.00	2.0412	2.0597	2.0773	2.0942	2.1104	2.1260
2000.00	1.926	1.947	1.965	1.983	1.999	2.015
3000.00	1.866	1.887	1.887	1.906	1.923	1.940
4000.00	1.798	1.819	1.839	1.857	1.877	1.890
5000.00	1.757	1.779	1.799	1.818	1.837	1.853

0.320 OXYGEN 0.680 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.21	1.24	1.28	1.31	1.34	1.38
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.320 OXYGEN 0.680 HELIUM
CP, RTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.4338	.4340	.4343	.4346	.4350	.4353
30.00	.4342	.4344	.4346	.4349	.4351	.4356
50.00	.4348	.4349	.4351	.4353	.4357	.4359
100.00	.4362	.4361	.4362	.4363	.4366	.4368
200.00	.4389	.4386	.4384	.4384	.4384	.4385
300.00	.4417	.4410	.4407	.4404	.4403	.4402
400.00	.4445	.4435	.4429	.4424	.4421	.4419
500.00	.4475	.4451	.4452	.4445	.4440	.4438
1000.00	.4623	.4593	.4569	.4552	.4535	.4523
2000.00	.49	.49	.49	.49	.48	.47
3000.00	.52	.51	.50	.49	.49	.49
4000.00	.52	.51	.51	.50	.49	.49
5000.00						

0.320 OXYGEN 0.680 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.49	.50	.50	.50	.50	.50
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.320 OXYGEN 0.680 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.540	1.547	1.546	1.546	1.545	1.544
30.00	1.549	1.548	1.547	1.547	1.546	1.545
50.00	1.551	1.550	1.548	1.548	1.547	1.546
100.00	1.555	1.553	1.552	1.551	1.550	1.548
200.00	1.563	1.560	1.559	1.557	1.555	1.554
300.00	1.572	1.568	1.565	1.562	1.561	1.559
400.00	1.580	1.575	1.572	1.568	1.566	1.564
500.00	1.589	1.583	1.579	1.574	1.572	1.569
1000.00	1.633	1.623	1.614	1.607	1.601	1.595
2000.00	1.7	1.7	1.7	1.7	1.7	1.7
3000.00	1.7	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

0.320 OXYGEN 0.680 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.376	1.416	1.457	1.496	1.533	1.574
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.320 OXYGEN 0.680 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1707	1740	1773	1806	1838	1870
30.00	1709	1742	1775	1807	1839	1871
50.00	1711	1744	1777	1809	1841	1873
100.00	1716	1749	1782	1814	1846	1878
200.00	1726	1759	1791	1823	1855	1887
300.00	1736	1769	1801	1833	1864	1895
400.00	1746	1778	1810	1842	1873	1904
500.00	1757	1788	1820	1851	1882	1913
1000.00	1808	1838	1868	1898	1928	1958
2000.00	1889	1922	1955	1988	2021	2054
3000.00	1970	2006	2042	2078	2114	2150
4000.00						
5000.00						

20 % O₂, 80 % He

T-54

0.200 OXYGEN 0.800 HELIUM
ENTHALPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	277.37	288.58	299.78	310.99	322.20	333.41
30.00	277.33	288.54	299.75	310.96	322.18	333.39
50.00	277.27	288.49	299.71	310.93	322.15	333.37
100.00	277.13	288.37	299.60	310.84	322.08	333.32
200.00	276.8	288.1	299.4	310.7	321.9	333.2
300.00	276.6	287.8	299.1	310.5	321.8	333.1
400.00	276.3	287.6	298.9	310.3	321.6	333.0
500.00	276.0	287.3	298.7	310.1	321.5	332.9
1000.00	274.4	286.2	297.8	309.3	320.8	332.4
2000.00	272.0	284.2	296.4	308.3	320.1	331.7
3000.00	269.9	282.6	295.3	307.7	319.6	331.2
4000.00	270.0	282.8	295.6	308.1	320.4	332.4
5000.00	270.4	283.5	295.7	308.9	321.1	333.6

0.200 OXYGEN 0.800 HELIUM
CV, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.3532	.3534	.3536	.3539	.3542	.3544
30.00	.3532	.3534	.3536	.3539	.3541	.3543
50.00	.3533	.3535	.3537	.3540	.3543	.3545
100.00	.3534	.3536	.3538	.3541	.3543	.3545
200.00	.3536	.3538	.3539	.3542	.3545	.3547
300.00	.3539	.3540	.3541	.3544	.3546	.3548
400.00	.3541	.3542	.3543	.3546	.3548	.3550
500.00	.3541	.3543	.3545	.3548	.3549	.3551
1000.00	.3554	.3554	.3554	.3555	.3557	.3557
2000.00	.36	.36	.36	.36	.36	.36
3000.00	.36	.37	.37	.37	.36	.36
4000.00						
5000.00						

0.200 OXYGEN 0.800 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3.3301	3.3605	3.3820	3.4028	3.4229	3.4422
30.00	3.2746	3.2970	3.3185	3.3393	3.3598	3.3788
50.00	3.1917	3.2141	3.2356	3.2564	3.2765	3.2959
100.00	3.0201	3.0426	3.0643	3.0850	3.1052	3.1245
200.00	2.4280	2.4506	2.4723	2.4931	2.5134	2.5327
300.00	2.8359	2.8585	2.8804	2.9012	2.9215	2.9410
400.00	2.7438	2.7665	2.7884	2.8094	2.8297	2.8492
500.00	2.6517	2.6745	2.6965	2.7175	2.7379	2.7575
1000.00	2.4664	2.4897	2.5121	2.5335	2.5542	2.5740
2000.00	2.316	2.341	2.364	2.386	2.407	2.427
3000.00	2.212	2.238	2.262	2.285	2.306	2.327
4000.00	2.149	2.174	2.199	2.221	2.242	2.263
5000.00	2.097	2.123	2.147	2.171	2.193	2.402

0.200 OXYGEN 0.800 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.93	1.97	1.98	1.96	1.70	1.75
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.200 OXYGEN 0.800 HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.5601	.5603	.5605	.5607	.5611	.5613
30.00	.5605	.5606	.5608	.5610	.5614	.5615
50.00	.5610	.5611	.5612	.5613	.5617	.5618
100.00	.5622	.5621	.5622	.5622	.5625	.5626
200.00	.5646	.5643	.5642	.5640	.5641	.5641
300.00	.5670	.5664	.5661	.5658	.5654	.5657
400.00	.5694	.5688	.5681	.5676	.5674	.5672
500.00	.5719	.5709	.5701	.5694	.5689	.5687
1000.00	.5847	.5822	.5801	.5786	.5762	.5761
2000.00	.61	.60	.60	.59	.59	.59
3000.00	.63	.62	.61	.61	.61	.60
4000.00	.63	.62	.62	.61	.61	.61
5000.00						

0.200 OXYGEN 0.800 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.50	.50	.50	.50	.50	.50
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

T-55

20 % O₂, 80 % He

0.200 OXYGEN 0.800 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.586	1.585	1.585	1.584	1.584	1.584
30.00	1.587	1.586	1.586	1.585	1.585	1.585
50.00	1.588	1.587	1.587	1.586	1.585	1.585
100.00	1.591	1.589	1.589	1.588	1.587	1.587
200.00	1.597	1.595	1.594	1.592	1.591	1.591
300.00	1.602	1.600	1.598	1.597	1.595	1.594
400.00	1.608	1.605	1.603	1.601	1.599	1.598
500.00	1.614	1.610	1.608	1.605	1.602	1.601
1000.00	1.645	1.638	1.632	1.627	1.620	1.619
2000.00	1.7	1.7	1.7	1.7	1.7	1.7
3000.00	1.7	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

0.200 OXYGEN 0.800 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.370	1.409	1.449	1.489	1.526	1.566
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.200 OXYGEN 0.800 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2088	2047	2086	2125	2163	2201
30.00	2010	2049	2088	2127	2164	2202
50.00	2012	2051	2090	2129	2166	2204
100.00	2017	2056	2095	2134	2171	2209
200.00	2028	2067	2106	2144	2181	2218
300.00	2040	2078	2116	2154	2191	2228
400.00	2051	2088	2126	2164	2200	2237
500.00	2062	2099	2137	2174	2210	2247
1000.00	2117	2153	2188	2224	2259	2294
2000.00	2208	2246	2284	2322	2359	2397
3000.00	2300	2340	2380	2420	2460	2500
4000.00						
5000.00						

T-59

SP. VOL.

SPECIFIC VOLUME-CUBIC FT/LB

15 % O₂, 85 % He

.150 OXYGEN .450 HELIUM AVERAGE MOLECULAR WEIGHT R-203

Table with columns: DEPTH FT, PRESSURE PSIA, ATM, TEMPERATURE °F (30-130), and specific volume values for various depths (800-5000 ft).

15 % O₂, 85 % He

T-60

0.150 OXYGEN 0.850 HELIUM
ENTHALPY, BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	319.11	331.97	344.82	357.71	370.58	383.46
30.00	319.04	331.95	344.81	357.70	370.58	383.46
50.00	319.06	331.93	344.80	357.69	370.58	383.46
100.00	318.99	331.88	344.77	357.66	370.57	383.46
200.00	318.8	331.8	344.7	357.6	370.6	383.5
300.00	318.7	331.6	344.6	357.6	370.5	383.5
400.00	318.5	331.5	344.5	357.5	370.5	383.5
500.00	318.4	331.4	344.4	357.5	370.5	383.5
1000.00	317.6	330.8	344.1	357.3	370.5	383.7
2000.00	316.6	330.2	343.9	357.4	370.8	384.1
3000.00	315.8	329.9	344.0	357.9	371.4	384.7
4000.00	316.7	331.2	345.3	359.3	373.1	386.8
5000.00	318.1	332.2	346.3	360.0	374.7	388.7

0.150 OXYGEN 0.850 HELIUM
CV, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.401	.401	.401	.402	.402	.402
30.00	.401	.401	.401	.402	.402	.402
50.00	.401	.401	.401	.402	.402	.402
100.00	.401	.401	.401	.402	.402	.402
200.00	.401	.401	.401	.402	.402	.402
300.00	.402	.402	.402	.402	.402	.403
400.00	.402	.402	.402	.402	.402	.403
500.00	.402	.402	.402	.402	.402	.403
1000.00	.403	.403	.403	.403	.403	.403
2000.00	.41	.41	.41	.41	.41	.41
3000.00	.41	.41	.41	.41	.41	.41
4000.00						
5000.00						

0.150 OXYGEN 0.850 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3.7573	3.7830	3.8077	3.8317	3.8547	3.8768
30.00	3.6857	3.7114	3.7361	3.7601	3.7831	3.8052
50.00	3.5921	3.6178	3.6426	3.6666	3.6896	3.7117
100.00	3.3897	3.4153	3.4402	3.4640	3.4871	3.5093
200.00	3.2815	3.3072	3.3322	3.3560	3.3792	3.4014
300.00	3.1732	3.1991	3.2241	3.2481	3.2712	3.2936
400.00	3.0650	3.0910	3.1161	3.1401	3.1633	3.1857
500.00	2.9568	2.9829	3.0081	3.0322	3.0554	3.0779
1000.00	2.7402	2.7667	2.7922	2.8167	2.8403	2.8629
2000.00	2.565	2.593	2.619	2.644	2.668	2.691
3000.00	2.446	2.474	2.501	2.527	2.551	2.575
4000.00	2.372	2.401	2.428	2.454	2.477	2.501
5000.00	2.314	2.342	2.370	2.396	2.421	2.445

0.150 OXYGEN 0.850 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.68	1.73	1.78	1.82	1.87	1.92
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.150 OXYGEN 0.850 HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.6433	.6434	.6437	.6439	.6442	.6444
30.00	.6436	.6437	.6440	.6441	.6444	.6446
50.00	.6441	.6441	.6443	.6445	.6447	.6449
100.00	.6451	.6450	.6452	.6453	.6454	.6455
200.00	.6473	.6469	.6469	.6468	.6469	.6469
300.00	.6494	.6488	.6486	.6484	.6483	.6482
400.00	.6516	.6507	.6503	.6500	.6498	.6495
500.00	.6539	.6527	.6521	.6516	.6512	.6508
1000.00	.6652	.6629	.6610	.6598	.6584	.6576
2000.00	.69	.68	.68	.68	.68	.66
3000.00	.72	.69	.68	.68	.68	.67
4000.00	.72	.70	.70	.69	.68	.68
5000.00						

0.150 OXYGEN 0.850 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.52	.52	.52	.52	.52	.52
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

T-61

15 % O₂, 85 % He

0.150 OXYGEN 0.850 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.604	1.604	1.604	1.603	1.603	1.602
30.00	1.605	1.605	1.605	1.604	1.603	1.602
50.00	1.606	1.605	1.605	1.604	1.604	1.603
100.00	1.608	1.607	1.607	1.606	1.606	1.604
200.00	1.613	1.611	1.611	1.609	1.609	1.607
300.00	1.617	1.615	1.614	1.613	1.612	1.610
400.00	1.622	1.619	1.618	1.616	1.615	1.613
500.00	1.627	1.623	1.622	1.619	1.618	1.616
1000.00	1.651	1.645	1.640	1.637	1.633	1.630
2000.00	1.7	1.7	1.7	1.7	1.7	1.7
3000.00	1.7	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

0.150 OXYGEN 0.850 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.359	1.398	1.437	1.476	1.512	1.551
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.150 OXYGEN 0.850 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2185	2227	2270	2312	2353	2395
30.00	2187	2229	2271	2313	2355	2396
50.00	2189	2231	2274	2315	2357	2398
100.00	2195	2237	2279	2321	2362	2403
200.00	2206	2248	2290	2331	2372	2413
300.00	2217	2259	2300	2341	2382	2423
400.00	2229	2270	2311	2352	2392	2433
500.00	2240	2281	2321	2362	2402	2442
1000.00	2297	2336	2375	2413	2452	2491
2000.00	2433	2459	2484	2517	2556	2595
3000.00	2570	2582	2594	2620	2660	2700
4000.00						
5000.00						

SP. VOL. 10% O2, 90% He

T-64

SPECIFIC VOLUME-CUBIC FT/LB

Table with columns for Depth (FT), Pressure (PSIA, ATM), Specific Volume (30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130), and Average Molecular Weight (6.803). The table provides specific volume and molecular weight data for a mixture of 10% oxygen and 90% helium at various depths and pressures.

0.010 OXYGEN 0.990 HELIUM
ENTHALPY, BTU/LB F

0.010 OXYGEN 0.990 HELIUM
CV, RTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	581.05	604.67	627.04	650.94	676.24	697.53
30.00	581.17	604.73	627.76	651.06	676.36	697.65
50.00	581.32	604.81	627.92	651.22	676.52	697.81
100.00	581.71	605.01	628.31	651.62	676.91	698.21
200.00	582.45	605.77	629.08	652.39	677.69	699.00
300.00	583.20	606.52	629.84	653.17	678.48	699.79
400.00	583.94	607.28	630.61	653.94	679.26	700.59
500.00	584.69	608.04	631.38	654.72	680.05	701.38
1000.00	588.38	611.77	635.17	658.54	683.92	705.28
2000.00	595.63	619.17	642.60	666.03	689.47	712.90
3000.00	602.78	626.37	649.95	673.47	696.95	720.40
4000.00	609.93	633.57	657.19	680.77	704.33	727.86
5000.00	616.95	640.59	664.22	687.92	711.50	734.88

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.702	.702	.702	.702	.702	.702
30.00	.702	.702	.702	.702	.702	.702
50.00	.702	.702	.702	.702	.702	.702
100.00	.702	.702	.702	.702	.702	.702
200.00	.702	.702	.702	.702	.702	.702
300.00	.702	.702	.702	.702	.702	.702
400.00	.702	.702	.702	.702	.702	.702
500.00	.702	.702	.702	.702	.702	.702
1000.00	.702	.702	.702	.702	.702	.702
2000.00	.703	.703	.703	.702	.702	.702
3000.00	.704	.704	.704	.703	.703	.703
4000.00						
5000.00						

0.010 OXYGEN 0.990 HELIUM
ENTROPY, BTU/LB F

0.010 OXYGEN 0.990 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	6.3200	6.3667	6.4114	6.4547	6.4963	6.5365
30.00	6.1973	6.2440	6.2887	6.3320	6.3736	6.4138
50.00	6.0304	6.0835	6.1283	6.1715	6.2132	6.2533
100.00	5.6398	5.6864	5.7313	5.7745	5.8161	5.8563
200.00	5.4307	5.4773	5.5223	5.5655	5.6071	5.6473
300.00	5.2216	5.2683	5.3132	5.3564	5.3981	5.4383
400.00	5.0126	5.0592	5.1042	5.1474	5.1891	5.2294
500.00	4.8035	4.8502	4.8952	4.9384	4.9801	5.0204
1000.00	4.3980	4.4368	4.4718	4.5252	4.5670	4.6073
2000.00	4.0654	4.1125	4.1576	4.2011	4.2424	4.2834
3000.00	3.8489	3.8961	3.9415	3.9831	4.0271	4.0676
4000.00	3.7136	3.7609	3.8064	3.8500	3.8914	3.9326
5000.00	3.6080	3.6553	3.7009	3.7447	3.7868	3.8275

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.20	2.26	2.32	2.39	2.45	2.51
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.010 OXYGEN 0.990 HELIUM
CP, RTU/LB F

0.010 OXYGEN 0.990 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.1655	1.1656	1.1656	1.1650	1.1657	1.1657
30.00	1.1656	1.1657	1.1657	1.1657	1.1658	1.1658
50.00	1.1657	1.1658	1.1658	1.1657	1.1658	1.1658
100.00	1.1660	1.1660	1.1660	1.1660	1.1660	1.1660
200.00	1.1665	1.1665	1.1664	1.1664	1.1664	1.1664
300.00	1.1671	1.1669	1.1669	1.1668	1.1668	1.1667
400.00	1.1676	1.1674	1.1673	1.1672	1.1672	1.1671
500.00	1.1681	1.1679	1.1677	1.1676	1.1675	1.1674
1000.00	1.1708	1.1703	1.1699	1.1697	1.1694	1.1691
2000.00	1.176	1.174	1.174	1.173	1.173	1.172
3000.00	1.180	1.179	1.178	1.177	1.177	1.176
4000.00	1.172	1.180	1.180	1.179	1.178	1.178
5000.00						

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.67	.67	.67	.67	.67	.67
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

T-67

10 % O₂, 90 % He

0.010 OXYGEN 0.990 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.660	1.660	1.660	1.660	1.661	1.661
30.00	1.660	1.660	1.660	1.660	1.661	1.661
50.00	1.660	1.660	1.660	1.660	1.661	1.661
100.00	1.661	1.661	1.661	1.661	1.661	1.661
200.00	1.661	1.661	1.661	1.661	1.662	1.661
300.00	1.662	1.662	1.662	1.662	1.662	1.662
400.00	1.663	1.663	1.663	1.663	1.663	1.662
500.00	1.664	1.664	1.663	1.663	1.663	1.662
1000.00	1.667	1.667	1.666	1.666	1.666	1.665
2000.00	1.67	1.67	1.67	1.67	1.67	1.67
3000.00	1.67	1.67	1.67	1.67	1.67	1.67
4000.00						
5000.00						

0.010 OXYGEN 0.990 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.269	1.303	1.336	1.371	1.403	1.437
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.010 OXYGEN 0.990 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3076	3137	3190	3257	3314	3372
30.00	3078	3139	3199	3258	3316	3374
50.00	3080	3141	3202	3261	3318	3376
100.00	3086	3147	3207	3266	3324	3381
200.00	3098	3158	3218	3277	3334	3392
300.00	3109	3169	3229	3288	3345	3402
400.00	3121	3181	3240	3299	3356	3413
500.00	3133	3192	3251	3310	3366	3423
1000.00	3191	3249	3307	3364	3420	3476
2000.00	3305	3358	3411	3465	3519	3573
3000.00	3420	3468	3516	3566	3618	3670
4000.00						
5000.00						

6 % O₂, 94 % He

T-72

0.040 OXYGEN 0.960 HELIUM
ENTHALPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	446.02	463.93	481.85	499.77	517.70	535.62
30.00	446.06	463.98	481.90	499.80	517.76	535.66
50.00	446.12	464.05	481.97	499.91	517.83	535.76
100.00	446.27	464.21	482.14	500.09	518.07	535.96
200.00	446.55	464.57	482.47	500.44	518.39	536.36
300.00	446.84	464.87	482.80	500.80	518.77	536.75
400.00	447.12	465.13	483.14	501.15	519.14	537.15
500.00	447.41	465.44	483.47	501.51	519.52	537.55
1000.00	448.78	466.97	485.13	503.25	521.37	539.45
2000.00	451.74	470.21	488.63	506.98	525.19	543.38
3000.00	454.78	473.52	492.27	510.81	529.13	547.33
4000.00	458.76	477.62	496.40	515.07	533.57	552.02
5000.00	462.88	481.62	500.35	519.36	537.90	556.56

0.040 OXYGEN 0.960 HELIUM
CV, RTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.547	.547	.547	.547	.547	.548
30.00	.547	.547	.547	.547	.547	.548
50.00	.547	.547	.547	.547	.547	.548
100.00	.547	.547	.547	.547	.547	.548
200.00	.547	.547	.547	.547	.547	.548
300.00	.547	.547	.547	.547	.547	.548
400.00	.547	.547	.547	.547	.547	.548
500.00	.547	.547	.547	.547	.547	.548
1000.00	.548	.548	.548	.548	.548	.548
2000.00	.551	.551	.551	.551	.551	.551
3000.00	.556	.556	.556	.556	.556	.556
4000.00						
5000.00						

0.040 OXYGEN 0.960 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	5.0129	5.0488	5.0832	5.1165	5.1485	5.1794
30.00	4.9165	4.9524	4.9868	5.0201	5.0521	5.0830
50.00	4.7905	4.8264	4.8609	4.8941	4.9262	4.9571
100.00	4.4937	4.5297	4.5626	4.5944	4.6262	4.6574
200.00	4.3366	4.3727	4.4035	4.4345	4.4726	4.4955
300.00	4.1796	4.2156	4.2345	4.2535	4.3157	4.4347
400.00	4.0225	4.0586	4.0854	4.1266	4.1588	4.2338
500.00	3.8655	3.9016	3.9366	3.9597	4.0019	4.0330
1000.00	3.5535	3.5896	3.6248	3.6584	3.6908	3.7220
2000.00	3.3240	3.3430	3.3587	3.3732	3.4451	3.4766
3000.00	3.1394	3.1772	3.2131	3.2478	3.2804	3.3123
4000.00	3.0363	3.0738	3.1058	3.1443	3.1769	3.2087
5000.00	2.9390	2.9925	3.0287	3.0634	3.0968	3.1287

0.040 OXYGEN 0.960 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.97	2.03	2.08	2.14	2.20	2.25
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.040 OXYGEN 0.960 HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.8962	.8964	.8964	.8965	.8967	.8969
30.00	.8964	.8966	.8966	.8967	.8968	.8970
50.00	.8967	.8968	.8968	.8969	.8970	.8972
100.00	.8974	.8974	.8974	.8974	.8975	.8976
200.00	.8988	.8987	.8985	.8985	.8985	.8985
300.00	.9002	.8999	.8997	.8995	.8994	.8993
400.00	.9016	.9011	.9008	.9006	.9004	.9002
500.00	.9030	.9024	.9019	.9016	.9013	.9010
1000.00	.9101	.9087	.9074	.9067	.9059	.9053
2000.00	.92	.92	.92	.92	.92	.91
3000.00	.94	.93	.93	.93	.93	.92
4000.00	.94	.93	.93	.93	.93	.93
5000.00						

0.040 OXYGEN 0.960 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.60	.60	.60	.60	.60	.60
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

T-73

6 % O₂, 94 % He

0.040 OXYGEN 0.960 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.639	1.639	1.638	1.638	1.638	1.638
30.00	1.639	1.639	1.638	1.638	1.638	1.638
50.00	1.640	1.640	1.639	1.639	1.639	1.638
100.00	1.641	1.641	1.640	1.640	1.639	1.639
200.00	1.643	1.642	1.642	1.641	1.641	1.640
300.00	1.645	1.644	1.644	1.643	1.642	1.642
400.00	1.647	1.646	1.646	1.645	1.644	1.643
500.00	1.649	1.648	1.648	1.646	1.645	1.644
1000.00	1.661	1.658	1.656	1.654	1.653	1.651
2000.00	1.7	1.7	1.7	1.7	1.7	1.7
3000.00	1.7	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

0.040 OXYGEN 0.960 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.316	1.351	1.387	1.424	1.458	1.495
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.040 OXYGEN 0.960 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2653	2705	2756	2807	2858	2909
30.00	2655	2706	2758	2809	2860	2911
50.00	2657	2709	2760	2811	2862	2913
100.00	2663	2714	2766	2817	2867	2917
200.00	2675	2726	2777	2827	2877	2927
300.00	2686	2737	2788	2838	2888	2937
400.00	2698	2748	2799	2849	2898	2947
500.00	2710	2760	2810	2859	2908	2957
1000.00	2764	2816	2865	2913	2960	3007
2000.00	2879	2927	2975	3023	3071	3118
3000.00	2990	3038	3086	3134	3182	3230
4000.00						
5000.00						

0.030 OXYGEN 0.970 HELIUM
ENTHALPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	519.60	540.45	561.29	582.15	603.00	623.85
30.00	519.64	540.54	561.38	582.24	603.09	623.94
50.00	519.79	540.65	561.50	582.30	603.22	624.07
100.00	520.07	540.93	561.79	582.66	603.52	624.37
200.00	520.60	541.48	562.35	583.24	604.11	624.97
300.00	521.13	542.02	562.91	583.82	604.70	625.54
400.00	521.67	542.57	563.48	584.41	605.30	626.18
500.00	522.20	543.12	564.04	584.99	605.89	626.79
1000.00	524.81	545.84	566.85	587.83	608.82	629.76
2000.00	530.06	551.28	572.47	593.44	614.64	635.68
3000.00	535.30	556.69	578.09	599.34	620.44	641.55
4000.00	541.00	562.46	583.88	605.23	626.50	647.73
5000.00	546.67	568.04	589.50	611.08	632.34	653.73

0.030 OXYGEN 0.970 HELIUM
CV, FT/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.632	.632	.632	.632	.632	.632
30.00	.632	.632	.632	.632	.632	.632
50.00	.632	.632	.632	.632	.632	.632
100.00	.632	.632	.632	.632	.632	.632
200.00	.632	.632	.632	.632	.632	.632
300.00	.632	.632	.632	.632	.632	.632
400.00	.632	.632	.632	.632	.632	.632
500.00	.632	.632	.632	.632	.632	.632
1000.00	.632	.632	.632	.632	.632	.632
2000.00	.634	.634	.634	.634	.634	.634
3000.00	.637	.637	.637	.637	.637	.637
4000.00						
5000.00						

0.030 OXYGEN 0.970 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	5.7320	5.7738	5.8139	5.8526	5.8898	5.9258
30.00	5.6213	5.6630	5.7031	5.7418	5.7791	5.8150
50.00	5.4765	5.5182	5.5583	5.5970	5.6343	5.6702
100.00	5.1252	5.1669	5.2071	5.2457	5.2831	5.3190
200.00	4.9398	4.9816	5.0218	5.0605	5.0974	5.1338
300.00	4.7545	4.7963	4.8365	4.8752	4.9124	4.9486
400.00	4.5691	4.6110	4.6513	4.6900	4.7273	4.7634
500.00	4.3838	4.4257	4.4660	4.5048	4.5421	4.5782
1000.00	4.0165	4.0586	4.0991	4.1380	4.1755	4.2116
2000.00	3.7272	3.8081	3.8104	3.8446	3.8814	3.9230
3000.00	3.5334	3.5762	3.6172	3.6570	3.6945	3.7311
4000.00	3.4126	3.4556	3.4968	3.5362	3.5737	3.6105
5000.00	3.3181	3.3611	3.4023	3.4421	3.4807	3.5169

0.030 OXYGEN 0.970 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.10	2.16	2.22	2.28	2.34	2.40
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.030 OXYGEN 0.970 HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.0431	1.0431	1.0432	1.0433	1.0434	1.0435
30.00	1.0432	1.0432	1.0433	1.0434	1.0435	1.0436
50.00	1.0434	1.0434	1.0435	1.0435	1.0436	1.0437
100.00	1.0439	1.0438	1.0438	1.0439	1.0440	1.0440
200.00	1.0448	1.0446	1.0446	1.0446	1.0446	1.0446
300.00	1.0457	1.0455	1.0453	1.0452	1.0453	1.0451
400.00	1.0466	1.0463	1.0461	1.0459	1.0459	1.0457
500.00	1.0476	1.0472	1.0469	1.0466	1.0465	1.0463
1000.00	1.0524	1.0515	1.0507	1.0502	1.0494	1.0492
2000.00	1.061	1.061	1.061	1.064	1.069	1.067
3000.00	1.072	1.067	1.065	1.063	1.063	1.061
4000.00	1.074	1.069	1.065	1.063	1.063	1.063
5000.00						

0.030 OXYGEN 0.970 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.64	.64	.64	.64	.64	.64
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.010 OXYGEN 0.970 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.652	1.652	1.652	1.652	1.652	1.651
30.00	1.652	1.652	1.652	1.652	1.652	1.651
50.00	1.652	1.652	1.652	1.652	1.652	1.651
100.00	1.653	1.653	1.653	1.653	1.653	1.652
200.00	1.654	1.654	1.654	1.654	1.653	1.653
300.00	1.656	1.655	1.655	1.654	1.654	1.654
400.00	1.657	1.656	1.656	1.655	1.655	1.655
500.00	1.658	1.657	1.657	1.656	1.656	1.656
1000.00	1.665	1.663	1.662	1.661	1.660	1.654
2000.00	1.67	1.67	1.67	1.67	1.67	1.67
3000.00	1.68	1.68	1.67	1.67	1.67	1.67
4000.00						
5000.00						

0.010 OXYGEN 0.970 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.290	1.325	1.359	1.395	1.428	1.463
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.030 OXYGEN 0.970 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2886	2947	2999	3054	3109	3164
30.00	2888	2944	3000	3056	3111	3166
50.00	2890	2946	3003	3058	3113	3168
100.00	2896	2952	3008	3064	3118	3173
200.00	2907	2963	3019	3074	3129	3183
300.00	2919	2975	3030	3085	3140	3193
400.00	2931	2986	3041	3096	3149	3203
500.00	2942	2997	3052	3106	3160	3213
1000.00	3000	3054	3107	3160	3212	3264
2000.00	3110	3161	3212	3262	3312	3362
3000.00	3220	3268	3316	3364	3412	3460
4000.00						
5000.00						

1.0 % O₂, 99.0 % He

T-88

0.100 OXYGEN 0.900 HELIUM
ENTHALPY, BTU/LB

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	378.00	393.21	408.42	423.64	438.85	454.07
30.00	378.01	393.22	408.44	423.66	438.88	454.10
50.00	378.02	393.24	408.46	423.68	438.91	454.14
100.00	378.06	393.29	408.52	423.76	439.00	454.24
200.00	378.11	393.37	408.63	423.90	439.16	454.44
300.00	378.17	393.44	408.75	424.05	439.33	454.63
400.00	378.22	393.52	408.86	424.19	439.49	454.83
500.00	378.28	393.60	408.98	424.34	439.66	455.03
1000.00	378.46	394.03	409.56	425.03	440.51	455.93
2000.00	379.27	395.20	411.07	426.85	442.44	457.99
3000.00	380.25	396.53	412.85	428.88	444.62	460.18
4000.00	382.63	399.06	415.41	431.60	447.56	463.46
5000.00	385.27	401.54	417.81	434.49	450.47	466.64

0.100 OXYGEN 0.900 HELIUM
CV, BTU/LH F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.469	.469	.469	.469	.469	.470
30.00	.469	.469	.469	.469	.469	.470
50.00	.469	.469	.469	.469	.469	.470
100.00	.469	.469	.469	.469	.469	.470
200.00	.469	.469	.469	.469	.469	.470
300.00	.469	.469	.470	.470	.470	.470
400.00	.469	.469	.470	.470	.470	.470
500.00	.469	.469	.470	.470	.470	.470
1000.00	.470	.470	.470	.470	.471	.471
2000.00	.475	.475	.475	.475	.476	.476
3000.00	.481	.481	.481	.475	.475	.475
4000.00						
5000.00						

0.100 OXYGEN 0.900 HELIUM
ENTROPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	4.3433	4.3737	4.4029	4.4312	4.4584	4.4846
30.00	4.2602	4.2906	4.3199	4.3481	4.3751	4.4015
50.00	4.1515	4.1819	4.2113	4.2394	4.2667	4.2929
100.00	3.9053	4.0029	3.9651	3.9932	4.0206	4.1938
200.00	3.7744	3.9152	3.8343	3.8625	3.8899	4.0264
300.00	3.6435	3.7476	3.7036	3.7319	3.7593	3.8591
400.00	3.5127	3.5801	3.5720	3.6012	3.6286	3.6917
500.00	3.3818	3.4125	3.4421	3.4708	3.4980	3.5244
1000.00	3.1209	3.1520	3.1819	3.2108	3.2382	3.2649
2000.00	2.9123	2.9445	2.9749	3.0042	3.0320	3.0593
3000.00	2.7713	2.8040	2.8350	2.8657	2.8931	2.9205
4000.00	2.6842	2.7166	2.7483	2.7775	2.8053	2.8329
5000.00	2.6147	2.6473	2.6790	2.7092	2.7380	2.7654

0.100 OXYGEN 0.900 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY 0.000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.83	1.89	1.94	1.99	2.04	2.10
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.100 OXYGEN 0.900 HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.7607	.7610	.7611	.7614	.7614	.7618
30.00	.7610	.7612	.7613	.7616	.7616	.7620
50.00	.7613	.7615	.7616	.7619	.7618	.7622
100.00	.7622	.7623	.7623	.7625	.7624	.7627
200.00	.7640	.7639	.7638	.7638	.7637	.7637
300.00	.7658	.7654	.7652	.7651	.7649	.7648
400.00	.7676	.7670	.7667	.7664	.7661	.7658
500.00	.7695	.7687	.7682	.7677	.7673	.7669
1000.00	.7789	.7771	.7755	.7744	.7733	.7726
2000.00	.80	.79	.79	.79	.79	.78
3000.00	.82	.81	.81	.80	.80	.79
4000.00	.82	.81	.81	.81	.80	.79
5000.00						

0.100 OXYGEN 0.900 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.56	.56	.56	.56	.56	.55
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.100 OXYGEN 0.900 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.622	1.623	1.623	1.623	1.623	1.623
30.00	1.623	1.623	1.623	1.623	1.623	1.623
50.00	1.623	1.624	1.624	1.624	1.624	1.624
100.00	1.625	1.625	1.625	1.625	1.625	1.624
200.00	1.628	1.628	1.628	1.627	1.626	1.626
300.00	1.632	1.631	1.630	1.630	1.628	1.627
400.00	1.635	1.634	1.633	1.632	1.630	1.629
500.00	1.638	1.637	1.636	1.634	1.632	1.631
1000.00	1.656	1.652	1.649	1.646	1.644	1.642
2000.00	1.7	1.7	1.7	1.7	1.7	1.7
3000.00	1.7	1.7	1.7	1.7	1.7	1.7
4000.00						
5000.00						

0.100 OXYGEN 0.900 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.344	1.377	1.415	1.452	1.488	1.525
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.100 OXYGEN 0.900 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2412	2460	2507	2554	2600	2646
30.00	2414	2461	2509	2556	2602	2648
50.00	2416	2464	2511	2558	2604	2649
100.00	2422	2469	2516	2563	2609	2654
200.00	2433	2480	2527	2573	2619	2664
300.00	2445	2492	2538	2584	2629	2674
400.00	2457	2503	2549	2594	2639	2684
500.00	2468	2514	2560	2605	2649	2694
1000.00	2526	2570	2614	2657	2700	2743
2000.00	2633	2677	2721	2765	2808	2851
3000.00	2740	2784	2828	2872	2916	2960
4000.00						
5000.00						

T-93

SP. VOL.

0.7 % O₂, 99.3 % He

SPECIFIC VOLUME-CUBIC FT/LB

.007 OXYGEN .993 HELIUM AVERAGE MOLECULAR WEIGHT 4.199

Table with columns: DEPTH FT, PRESSURE PSIA, ATM, TEMPERATURE F (30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130), and specific volume values for various depths and temperatures.

T-97

SP. VOL. 0.5 % O₂, 99.5 % He

SPECIFIC VOLUME CUBIC FT/LB

.005 OXYGEN .995 HELIUM AVERAGE MOLECULAR WEIGHT 4.143

Table with columns: DEPTH FT, PRESSURE PSIA, ATM, TEMPERATURE °F (30, 40, 50, 60, 70, 80, 90, 100, 110, 120, 130), and specific volume values.

0.005 OXYGEN 0.995 HELIUM
ENTHALPY, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	599.57	623.60	647.63	671.66	695.71	719.74
30.00	599.70	623.73	647.76	671.81	695.84	719.87
50.00	599.86	623.90	647.93	671.98	696.01	720.04
100.00	600.28	624.32	648.36	672.40	696.43	720.47
200.00	601.09	625.14	649.18	673.23	697.27	721.32
300.00	601.90	625.96	650.01	674.06	698.11	722.16
400.00	602.71	626.78	650.83	674.90	698.95	723.01
500.00	603.52	627.60	651.66	675.73	699.79	723.86
1000.00	607.53	631.64	655.75	674.84	703.94	728.02
2000.00	615.36	639.55	663.72	677.87	712.01	736.15
3000.00	623.08	647.34	671.58	685.79	719.96	744.13
4000.00	630.66	654.96	679.24	703.50	726.79	751.98
5000.00	638.09	662.40	686.70	711.04	735.31	759.60

0.005 OXYGEN 0.995 HELIUM
CV, BTU/LH F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.723	.723	.723	.723	.723	.723
30.00	.723	.723	.723	.723	.723	.723
50.00	.723	.723	.723	.723	.723	.723
100.00	.723	.723	.723	.723	.723	.723
200.00	.723	.723	.723	.723	.723	.723
300.00	.723	.723	.723	.723	.723	.723
400.00	.723	.723	.723	.723	.723	.723
500.00	.723	.723	.723	.723	.723	.723
1000.00	.723	.723	.723	.723	.723	.723
2000.00	.724	.724	.724	.724	.724	.724
3000.00	.724	.724	.724	.724	.724	.724
4000.00	.724	.724	.724	.724	.724	.724
5000.00	.724	.724	.724	.724	.724	.724

0.005 OXYGEN 0.995 HELIUM
ENTROPY, BTU/LH F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	6.4955	6.5439	6.5900	6.6347	6.6774	6.7190
30.00	6.3692	6.4175	6.4637	6.5083	6.5513	6.5927
50.00	6.2042	6.2523	6.2986	6.3431	6.3861	6.4275
100.00	5.7934	5.8415	5.8878	5.9323	5.9752	6.0168
200.00	5.5772	5.6253	5.6716	5.7161	5.7591	5.8007
300.00	5.3610	5.4091	5.4554	5.5000	5.5429	5.5845
400.00	5.1448	5.1930	5.2393	5.2838	5.3268	5.3684
500.00	4.9286	4.9768	5.0231	5.0677	5.1107	5.1523
1000.00	4.5011	4.5494	4.5958	4.6405	4.6836	4.7252
2000.00	4.1659	4.2149	4.2608	4.3057	4.3488	4.3905
3000.00	3.9426	3.9911	4.0378	4.0827	4.1259	4.1676
4000.00	3.8029	3.8515	3.8983	3.9432	3.9864	4.0282
5000.00	3.6940	3.7427	3.7895	3.8345	3.8779	3.9198

0.005 OXYGEN 0.995 HELIUM
THERMAL CONDUCTIVITY, BTU/SEC FT F
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	2.23	2.24	2.26	2.42	2.49	2.55
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.005 OXYGEN 0.995 HELIUM
CP, BTU/LB F

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.2025	1.2025	1.2025	1.2025	1.2026	1.2026
30.00	1.2026	1.2026	1.2026	1.2026	1.2026	1.2026
50.00	1.2026	1.2026	1.2026	1.2026	1.2027	1.2027
100.00	1.2029	1.2029	1.2028	1.2028	1.2029	1.2028
200.00	1.2033	1.2032	1.2032	1.2031	1.2032	1.2031
300.00	1.2037	1.2035	1.2035	1.2035	1.2035	1.2033
400.00	1.2041	1.2039	1.2039	1.2038	1.2038	1.2036
500.00	1.2045	1.2043	1.2042	1.2041	1.2041	1.2039
1000.00	1.2065	1.2062	1.2059	1.2057	1.2055	1.2053
2000.00	1.210	1.209	1.209	1.208	1.208	1.208
3000.00	1.214	1.213	1.212	1.211	1.211	1.210
4000.00	1.215	1.214	1.214	1.213	1.213	1.212
5000.00						

0.005 OXYGEN 0.995 HELIUM
PRANDTL NUMBER

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	.68	.68	.68	.68	.67	.67
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.005 OXYGEN 0.995 HELIUM
CP/CV

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.663	1.663	1.663	1.663	1.663	1.663
30.00	1.663	1.663	1.663	1.663	1.663	1.663
50.00	1.663	1.663	1.663	1.663	1.663	1.663
100.00	1.663	1.663	1.663	1.663	1.663	1.663
200.00	1.664	1.663	1.663	1.663	1.663	1.663
300.00	1.664	1.664	1.664	1.664	1.664	1.664
400.00	1.665	1.664	1.664	1.664	1.664	1.664
500.00	1.665	1.664	1.664	1.664	1.664	1.664
1000.00	1.668	1.667	1.667	1.667	1.666	1.666
2000.00	1.67	1.67	1.67	1.67	1.67	1.67
3000.00	1.68	1.67	1.67	1.67	1.67	1.67
4000.00						
5000.00						

0.005 OXYGEN 0.995 HELIUM
VISCOSITY, LB/FT SEC
(MULTIPLY TABLE ENTRY BY .0000100)

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	1.263	1.296	1.329	1.364	1.396	1.429
30.00						
50.00						
100.00						
200.00						
300.00						
400.00						
500.00						
1000.00						
2000.00						
3000.00						
4000.00						
5000.00						

0.005 OXYGEN 0.995 HELIUM
SONIC VELOCITY FT/SEC

PRESSURE PSIA	TEMPERATURE					
	30	50	70	90	110	130
14.70	3130	3192	3253	3314	3373	3433
30.00	3132	3193	3255	3315	3375	3435
50.00	3134	3196	3257	3318	3377	3437
100.00	3140	3201	3263	3323	3382	3442
200.00	3152	3213	3274	3334	3393	3452
300.00	3163	3224	3285	3345	3404	3463
400.00	3175	3235	3296	3356	3414	3473
500.00	3187	3247	3307	3366	3425	3483
1000.00	3245	3304	3363	3421	3478	3535
2000.00	3362	3416	3469	3523	3578	3632
3000.00	3480	3528	3576	3626	3678	3730
4000.00						
5000.00						

T-101

DENSITY

DENSITY, LBS./CUBIC FT.

0.3 % O₂, 99.7 % He

Table with columns for DEPTH (FT), PRESSURE (PSIA, ATM), TEMPERATURE (F), and DENSITY (0.003 OXYGEN, .007 HELIUM, AVERAGE MOLECULAR WEIGHT, 4.007). The table contains multiple rows of data corresponding to different depths and pressures.

